

(19) World Intellectual Property  
Organization  
International Bureau



(43) International Publication Date  
19 August 2004 (19.08.2004)

PCT

(10) International Publication Number  
**WO 2004/069293 A1**

(51) International Patent Classification<sup>7</sup>: **A61L 15/60**

(21) International Application Number:  
PCT/JP2004/001294

(22) International Filing Date: 6 February 2004 (06.02.2004)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:  
2003-032698 10 February 2003 (10.02.2003) JP  
2003-063548 10 March 2003 (10.03.2003) JP  
2003-310386 2 September 2003 (02.09.2003) JP  
2003-322031 12 September 2003 (12.09.2003) JP

(71) Applicant (for all designated States except US): NIPPON SHOKUBAI CO., LTD. [JP/JP]; 1-1, Koraibashi 4-chome, Chuo-ku, Osaka-shi, Osaka 5410043 (JP).

(72) Inventors; and

(75) Inventors/Applicants (for US only): TORII, Kazushi [JP/JP]; F-202, 931-11, Hamada, Aboshi-ku, Himeji-shi, Hyogo 6711242 (JP). MITSUKAMI, Yoshiro [JP/JP]; 931-11, Hamada, Aboshi-ku, Himeji-shi, Hyogo 6711242 (JP). IMURA, Motohiro [JP/JP]; 1-8-1-604, Ooakashi-cho, Akashi-shi, Hyogo 6730891 (JP). IWA-MURA, Taku [JP/JP]; 931-11, Hamada, Aboshi-ku, Himeji-shi, Hyogo 6711242 (JP). KITAYAMA, Toshimasa [JP/JP]; 2-60, Kiyomi-cho, Hirohata-ku, Himeji-shi, Hyogo 6711106 (JP). KADONAGA, Kenji [JP/JP]; 915-9, Nakatsu, Kakogawa-cho, Kakogawa-shi, Hyogo 6750068 (JP). INOUE, Hiroki [JP/JP]; 50, Uchida-cho, Arashiyama, Nishikyoku-ku, Kyoto-shi, Kyoto 6160027 (JP). MIYAKE, Koji [JP/JP]; 318-5, Mantomi, Seto-cho, Akaiwa-gun, Okayama 7090841 (JP). NOGI, Kozo [JP/JP]; 485-3, Minori, Kakogawa-cho, Kakogawa-shi,

Hyogo 6750062 (JP). NAKAMURA, Masatoshi [JP/JP]; 721-4, Tsuichiba, Aboshi-ku, Himeji-shi, Hyogo 6711224 (JP). SAKAMOTO, Shigeru [JP/JP]; 391-10, Shimoyobe, Yobe-ku, Himeji-shi, Hyogo 6711261 (JP). MACHIDA, Sayaka [JP/JP]; 1-1-202, Honmachi, Hirohata-ku, Himeji-shi, Hyogo 6711114 (JP).

(74) Agent: MATSUMOTO, Takehiko; 25-6, Hannan-cho 1-chome, Abeno-ku, Osaka-shi, Osaka 5450021 (JP).

(81) Designated States (unless otherwise indicated, for every kind of national protection available): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BW, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NA, NL, NO, NZ, OM, PG, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, SY, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.

(84) Designated States (unless otherwise indicated, for every kind of regional protection available): ARIPO (BW, GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IT, LU, MC, NL, PT, RO, SE, SI, SK, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

**Published:**

- with international search report
- before the expiration of the time limit for amending the claims and to be republished in the event of receipt of amendments

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: WATER-ABSORBENT RESIN COMPOSITION AND ITS PRODUCTION PROCESS

(57) Abstract: There are disclosed a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition causes little gel-blocking and is excellent in the liquid permeability and liquid diffusibility and is high also in the absorption performances and further is strong also against the physical damage; and there are further disclosed a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition has the following further advantages, in addition to the above, of involving little segregation of the metal compound and further having a dust prevention effect. One of water-absorbent resin compositions according to the present invention is a water-absorbent resin composition comprising water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt, with the composition having a mass-average particle diameter of 100 to 600 µm and comprising water-soluble polyvalent metal salt particles and the water-absorbent resin particles that have been surface-crosslinked.

WO 2004/069293 A1

## DESCRIPTION

## Water-Absorbent Resin Composition and Its Production Process

5

## TECHNICAL FIELD

The present invention relates to a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition is used for sanitary materials such as disposable diapers, sanitary napkins, and so-called incontinent pads.

10

## BACKGROUND ART

For sanitary materials such as disposable diapers, sanitary napkins, and incontinent pads, there are widely utilized absorbent structures comprising hydrophilic fibers (e.g. pulp) and water-absorbent resins as constituent materials for the purpose of absorption of body fluids.

15

In recent years, as to these sanitary materials, their high functionalization and thinning are making progress, so there is a tendency toward increases in the amount of the water-absorbent resin as used per piece of sanitary material and in the ratio of the water-absorbent resin relative to a whole absorbent structure comprising the water-absorbent resin and the hydrophilic fibers. Specifically, the ratio of the water-absorbent resin in the absorbent structure is raised by decreasing the amount of the hydrophilic fibers (which have a small bulk density) and increasing the amount of the water-absorbent resin (which has excellent water absorbency and a large bulk density) as used. Thereby the thinning of the sanitary materials is aimed at without lowering the water absorption quantity.

20

25

However, the sanitary materials, in which the ratio of the hydrophilic fibers has been decreased and that of the water-absorbent resin has been increased in the above way, are favorable from the viewpoint of simple storage of liquids, but rather involve problems in the case of consideration of distribution and diffusion of the liquids

under circumstances of actual use as such as diapers. For example, the large amount of water-absorbent resin becomes a soft gel due to water absorption to cause a phenomenon "gel-blocking", thus dramatically deteriorating the ability to diffuse the liquids in the sanitary materials. In order to avoid such problems to maintain  
5 absorption properties of the absorbent structure, the ratios of the hydrophilic fibers and the water-absorbent resin have axiomatically been limited, so a limit has occurred also to the thinning of the sanitary materials.

As means for preventing the gel-blocking to thus obtain the water-absorbent resin excellent in the liquid permeability and liquid diffusibility, there are known the  
10 following arts in which metal compounds (e.g. metal salts, metal cations) are added to water-absorbent resins.

There is known a water-insoluble water-absorbent resin composition obtained by adding water, containing a salt and/or hydroxide of a polyvalent metal, to a water-absorbent resin (patent document 1).

15 There is known a process for production of a water-absorbent resin in which a water-absorbent resin is treated with an aluminum compound in the presence of a polyhydric alcohol and water, wherein the aluminum compound is reactable with the water-absorbent resin (patent document 2).

There is known a process for production of a water-absorbent resin in which a  
20 water-absorbent resin is treated with an aluminum compound and a crosslinking agent in the presence of a polyhydric alcohol and water, wherein the aluminum compound is reactable with the water-absorbent resin, and the crosslinking agent has not fewer than two functional groups reactable with the water-absorbent resin (patent document 3).

25 There is known a process for production of water-absorbent resin particles having the modified particulate brittleness, in which process, to water-absorbent resin particles obtained by heat-crosslinking of surfaces and their neighborhood of particles of a water-absorbent resin, there are added, after this heat-crosslinking,

water in which an inorganic salt is dissolved in a concentration of 5 to 50 weight % relative to water and/or water in which an inorganic hydroxide is dissolved in a concentration of 5 to 50 weight % relative to water, thereby adjusting the water content to 3-9 % (patent document 4).

5        There is known a polymer produced by a process in which a water-absorbent resin is treated with a polyol and a cation which is in a state of an aqueous solution and then surface-crosslinked at 150-300 °C (patent document 5).

          There is known a polymer produced by a process in which a water-absorbent resin is treated with an organic surface-secondary-crosslinking agent (except polyols)  
10        and a cation which is in a state of an aqueous solution and then surface-crosslinked (patent document 6).

          There is known a composition comprising aqueous-fluid-absorbent polymer particles having been heat-treated at a temperature higher than 170 °C for more than 10 minutes, wherein the composition is remoisturized with an aqueous additive  
15        solution in the absence of an organic solvent or a water-insoluble and non-water-swelling powder after the heat-treatment and has a water content of 1-10 weight % based on the total weight of the composition and displays an absorption capacity of more than 20 g/g under 0.3 psi in 60 minutes (patent document 7).

          These (patent documents 1 to 7) are arts in which the metal compounds (e.g.  
20        metal salts, metal cations) are added in aqueous solution states. As to these arts, because the metal compounds (e.g. metal salts, metal cations) are added in aqueous solution states, the metal components unfavorably permeate the inside of the water-absorbent resins, thus resulting in insufficiency of the effect of enhancing the liquid permeability and liquid diffusibility to a degree corresponding to the addition  
25        amount. In addition, because the metal components permeate the inside of the water-absorbent resins, there have unfavorably occurred deteriorations of such as absorption capacity without load and absorption capacity under load.

          There is known a modified water-insoluble water-absorbent resin composition



obtained by adding water to a mixture of a water-absorbent resin and a salt and/or hydroxide of a polyvalent metal (patent document 8).

There is known a method in which: a water-absorbent resin and a polyvalent metal salt are mixed together, and then the resultant mixture is brought into close  
5 contact with a binder in the absence of a volatile alcohol (patent document 9).

As to these arts (patent documents 8 to 9), there have been problems such that: the dissolved metal salt causes binding between particles to thus easily form a strong agglomerate and, in the case where this agglomerate is crushed by physical damage such as during the actual production or practical use, the absorption capacity under  
10 load is deteriorated. In addition, there have also been problems such that: the dissolved metal salt unfavorably goes so far as permeating into particles of the water-absorbent resin. The case where particles of the polyvalent metal salt having small particle diameters are used has been remarkable for the aforementioned permeation. Because of this permeation, there have been the same problems as the  
15 aforementioned. Specifically, the effect of enhancing the liquid permeability and liquid diffusibility to a degree corresponding to the addition amount has been insufficient or, because the metal components permeate the inside of the water-absorbent resins, there have unfavorably occurred deteriorations of such as absorption capacity without load and absorption capacity under load. In addition,  
20 as to particles of the polyvalent metal salt having comparatively large particle diameters, no sufficient binding force between particles can be obtained with the binder, and therefore such as release or elimination unfavorably occurs, so that problems of such as segregation of the metal compounds (e.g. metal salts) have also been caused.

25 As to other than these methods, for example, as to a method in which a water-absorbent resin and a metal compound (e.g. metal salt) are dry-blended together, particles are mixed with each other. Therefore, there is a possibility of occurrence of problems such that the segregation occurs to thus result in unstable

performances of the water-absorbent resin.

As means for preventing the gel-blocking to thus obtain the water-absorbent resin excellent in the liquid permeability and liquid diffusibility, there are known some other arts besides the above arts as follows.

5 For example,, there are proposed such as: a method in which two kinds of water-absorbent resins different as to water absorption performance are used (patent document 10); a method in which a composition containing a cationic ion-exchange hydrogel-forming polymer and an anionic ion-exchange hydrogel-forming polymer is used (patent document 11); and a method in which a water-absorbent resin having  
10 a high surface-crosslinking density is used (patent document 12). However, they have problems such that the absorption properties are unsatisfactory as the absorbent structure having a high water-absorbent resin concentration or that the cost is high.

In addition, a water-absorbent resin which contains a large amount of fine powder due to such as abrasion in processes for production of the water-absorbent  
15 resin has a tendency to cause the gel-blocking. Therefore, there is proposed a method in which the water-absorbent resin is made to contain water in an amount of not smaller than 3 %, thereby improving the brittleness (patent document 13). However, there are problems such that the absorption capacity is deteriorated, and that, when water is added to the water-absorbent resin, this resin swells to thus form  
20 particles having too large particle diameters. In addition, it is also proposed that a special stirring apparatus is used to reduce the formation of the fine powder in processes for production of the water-absorbent resin (patent document 14).

In addition, there are known such as: a method in which a water-absorbent resin is mixed with a powder of an organic or inorganic water-soluble salt (specific  
25 salt such as thiourea, saccharide, or carboxylate salt), thereby enhancing the absorption of blood (patent document 15); a method in which a water-absorbent resin and a permeability-retaining agent (e.g. silica, alumina, titania, clay, emulsion-polymerized material, precipitation-polymerized material) are mixed

together by a Vortex Mixer and then subjected to mechanical stress by such as Osterizer blender (patent document 16); a method in which a water-insoluble and water-swellaable hydrogel is coated with steric or electrostatic spacers (patent document 17); a method in which a water-absorbent resin having been crosslinked  
5 with a specific metal ion is used (patent documents 18 and 19); and a super-water-absorbent resin composition comprising a super-water-absorbent resin and a fine powder of an aggregate of a hydro-oxide which contains two kinds of metals M1 and M2 at least partially having an -M1-O-M2- bond (patent document 20).

10 As to these publicly known methods (patent documents 15 to 20), the gel-blocking can be prevented, but there have occurred problems such that the durability of the performance to diffuse liquids in diapers, particularly, the Saline Flow Conductivity (hereinafter abbreviated to SFC), is low. Or, even if the performance to diffuse the liquids is enough, there is not taken into consideration  
15 various performance deteriorations due to such as mechanical impact or friction which the water-absorbent resin undergoes when it is produced or used to produce absorbent articles, and therefore no sufficient performance can be maintained in the actual production. For example, as the case may be, even if improvement effects are seen in laboratories, those effects are not seen or are deteriorated when the  
20 production is carried out with a production machine involving the step in which physical energy works against the powder such as stirring or pneumatic transportation.

There is known a water-absorbing agent comprising 100 weight parts of water-absorbent resin particles and 1 to 30 weight parts of a heat-fusible resin  
25 powder having a melting point in the range of 50 to 160 °C (patent document 21).

In this art (patent document 21), there is disclosed a method in which the water-absorbent resin particles and the heat-fusible resin powder having a melting point in the range of 50 to 160 °C are heat-treated after or during their mixing.

whereby the heat-fusible resin powder is fixed to the water-absorbent resin particles. Such a heat-fusible resin powder is used for the purpose of enhancing the fixability to fibers such as pulp, in other words, as a binder for the fibers and the water-absorbent resin particles. However, such a heat-fusible resin powder enhances the fixability of the water-absorbent resin particles to the fibers, but has no interactions with a carboxyl group. Therefore, it has been impossible to obtain the effect of enhancing the liquid permeability and liquid diffusibility of the water-absorbent resin. In addition, in the case where the heat-fusible resin powder has strong hydrophobicity, it may cause such as deterioration of the capillary suction force of the resultant water-absorbent resin composition. Therefore, the resultant water-absorbent resin composition has not necessarily been a water-absorbent resin composition having sufficient performances.

- [Patent document 1] JP-A-007745/1987 (Kokai)
- [Patent document 2] JP-A-270741/1988 (Kokai)
- 15 [Patent document 3] JP-A-056707/1989 (Kokai)
- [Patent document 4] JP-A-124879/1997 (Kokai)
- [Patent document 5] JP-A-539281/2002 (Kohyo)
- [Patent document 6] JP-A-538275/2002 (Kohyo)
- [Patent document 7] JP-A-523287/2001 (Kohyo)
- 20 [Patent document 8] JP-A-257235/1986 (Kokai)
- [Patent document 9] JP-A-523289/2001 (Kohyo)
- [Patent document 10] JP-A-252307/2001 (Kokai)
- [Patent document 11] pamphlet of WO 98/037149
- [Patent document 12] JP-A-057010/1994 (Kokai)
- 25 [Patent document 13] pamphlet of WO 01/25290
- [Patent document 14] pamphlet of WO 97/24394
- [Patent document 15] USP 4,693,713
- [Patent document 16] pamphlet of WO 01/66056

[Patent document 17] US 2002/0128618A1

[Patent document 18] JP-A-513043/2002 (Kohyo)

[Patent document 19] JP-A-513059/2002 (Kohyo)

[Patent document 20] JP-A-147724/1998 (Kokai)

5 [Patent document 21] JP-A-248187/1994 (Kokai)

## DISCLOSURE OF THE INVENTION

### OBJECTS OF THE INVENTION

An object of the present invention is to provide a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition causes little gel-blocking and is excellent in the liquid permeability and liquid diffusibility and is high also in the absorption performances (e.g. absorption capacity without load, absorption capacity under load, capillary absorption capacity) and further is strong also against the physical damage such as during the actual production or practical use.

15 Also, another object of the present invention is to provide a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition has the following further advantages, in addition to the above, of involving little segregation of the metal compound (e.g. metal salt), and being excellent also in the handling property during the moisture absorption, and further  
20 having a dust prevention effect.

### SUMMARY OF THE INVENTION

The present inventors diligently studied to solve the aforementioned problems. As a result, they have found out that: if there is constructed a water-absorbent resin composition which comprises water-absorbent resin particles and water-soluble polyvalent metal salt particles, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt, and wherein the water-absorbent resin particles are surface-crosslinked ones, and wherein the water-absorbent resin composition has a mass-average particle diameter of 100 to

600  $\mu\text{m}$ , then the above problems can successfully be solved, because the resultant water-absorbent resin composition causes little gel-blocking and is excellent in the liquid permeability and liquid diffusibility and is high also in the absorption performances (e.g. absorption capacity without load, absorption capacity under load, capillary absorption capacity) and further is strong also against the physical damage such as during the actual production or practical use. At the same time, the present inventors have found out also that: such a water-absorbent resin composition is a water-absorbent resin composition which displays a high saline flow conductivity (SFC) and is excellent in the retention ratio of the saline flow conductivity (SFC) after a paint shaker test and in the retention ratio of the saline flow conductivity (SFC) after a long-term liquid absorption.

In addition, the present inventors have further found out that: if, in a process for production of a water-absorbent resin composition in which a polyvalent metal is fixed to surfaces of water-absorbent resin particles, there are involved a step in which a binder (e.g. water) is beforehand added to the water-absorbent resin particles to thus put them in a state where the binder is permeated across surfaces of the water-absorbent resin particles and a step in which they are thereafter mixed with water-soluble polyvalent metal salt particles, then the above problems can successfully be solved, because: the permeation of the polyvalent metal into the water-absorbent resin particles can effectively be prevented, and further, the polyvalent metal is fixed all over the surfaces of the water-absorbent resin particles uniformly and moderately (i.e. in a state where the fixation is incomplete, but is not so weak as to enable free migration), and consequently, the gel-blocking can sufficiently be prevented, and therefore excellent liquid permeability and liquid diffusibility can be displayed and also excellent absorption performances can be displayed, and further, the resultant water-absorbent resin composition comes into a state which is strong also against the physical damage such as during the actual production or practical use.

Moreover, the present inventors have further found out that: if at least a part of a metal compound is fused to surfaces of water-absorbent resin particles in a water-absorbent resin composition comprising the water-absorbent resin particles and the metal compound wherein the water-absorbent resin particles are obtained by  
5 polymerizing a monomer including acrylic acid and/or its salt, then surprisingly the gel-blocking can sufficiently be prevented, and therefore excellent liquid permeability and liquid diffusibility can be displayed and also excellent absorption performances can be displayed, and further, the resultant water-absorbent resin composition comes into a state which is strong also against the physical damage such  
10 as during the actual production or practical use, and besides, this composition involves little segregation of the metal compound (e.g. metal salt), and is excellent also in the handling property during the moisture absorption, and further has a dust prevention effect.

That is to say, a first water-absorbent resin composition according to the  
15 present invention (which may hereinafter be referred to as water-absorbent resin composition (1)) is a water-absorbent resin composition comprising water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt,

with the composition having a mass-average particle diameter of 100 to 600  
20  $\mu\text{m}$  and comprising water-soluble polyvalent metal salt particles and the water-absorbent resin particles that have been surface-crosslinked.

As to the first water-absorbent resin composition according to the present invention, it is favorable that at least a part of the water-absorbent resin particles are agglomerates.

25 As to the first water-absorbent resin composition according to the present invention, it is favorable that the water-soluble polyvalent metal salt particles are particles of an aluminum salt having water of crystallization.

As to the first water-absorbent resin composition according to the present

invention, it is favorable that the water-absorbent resin particles are those which have been surface-crosslinked with a polyhydric alcohol.

Also, the first water-absorbent resin composition according to the present invention is a water-absorbent resin composition comprising water-absorbent resin particles and water-soluble polyvalent metal salt particles, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt,

with the composition of which the saline flow conductivity is at least  $50 (\times 10^{-7} \text{ cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$  and of which the retention ratio of the saline flow conductivity is not less than 40 %.

As to the first water-absorbent resin composition according to the present invention, favorably, its retention ratio of the saline flow conductivity after a paint shaker test is not less than 70 %.

A first process according to the present invention for production of a water-absorbent resin composition (which may hereinafter be referred to as production process (1)) is characterized by comprising the steps of:

adding a binder to water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt; and then

mixing the binder and the water-absorbent resin particles with water-soluble polyvalent metal salt particles.

As to the first process according to the present invention for production of a water-absorbent resin composition, it is favorable that the water-absorbent resin particles are surface-crosslinked ones.

As to the first process according to the present invention for production of a water-absorbent resin composition, it is favorable that the binder contains a surface-crosslinking agent.

As to the first process according to the present invention for production of a water-absorbent resin composition, it is favorable that the binder includes water



and/or a polyhydric alcohol.

As to the first process according to the present invention for production of a water-absorbent resin composition, it is favorable that, when the binder is added to the water-absorbent resin particles, the temperature of the water-absorbent resin particles is in the range of 40 to 100 °C.

A second water-absorbent resin composition according to the present invention (which may hereinafter be referred to as water-absorbent resin composition (2)) comprises water-absorbent resin particles and a metal compound, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt, and wherein:

the metal compound is one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts (except polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule); and

at least a part of the metal compound is fused to surfaces of the water-absorbent resin particles.

As to the second water-absorbent resin composition according to the present invention, it is favorable that the water-absorbent resin particles are materials having been surface-crosslinked with a compound having at least two functional groups which make a dehydration reaction or transesterification reaction with a carboxyl group.

As to the second water-absorbent resin composition according to the present invention, it is favorable that at least a part of the metal compound is fused in the form of coating at least a part of surfaces of the water-absorbent resin particles in a layered state

As to the second water-absorbent resin composition according to the present invention, it is favorable that the metal compound has a melting point of not higher than 250 °C.

As to the second water-absorbent resin composition according to the present

invention, it is favorable that the metal compound is a water-soluble polyvalent metal salt.

As to the second water-absorbent resin composition according to the present invention, it is favorable that the metal compound is a water-soluble polyvalent metal salt having water of hydration and containing aluminum.

As to the second water-absorbent resin composition according to the present invention, favorably, it displays an absorption capacity of not less than 20 g/g under load.

As to the second water-absorbent resin composition according to the present invention, favorably, it displays a saline flow conductivity of not less than  $30 (\times 10^{-7} \text{ cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$  for a 0.69 mass % physiological saline solution.

A second process according to the present invention for production of a water-absorbent resin composition (which may hereinafter be referred to as production process (2)) is a process for production of a water-absorbent resin composition which includes water-absorbent resin particles and a metal compound, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt, and wherein:

the metal compound is one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts (except polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule);

with the process comprising the steps of:

heating the water-absorbent resin particles and/or the metal compound to a temperature of not lower than the melting point of the metal compound; and

thereby fusing at least a part of the metal compound to surfaces of the water-absorbent resin particles.

As to the second process according to the present invention for production of a water-absorbent resin composition, it is favorable that the fusion is carried out under stirring of the water-absorbent resin particles and/or the metal compound.

As to the second process according to the present invention for production of a water-absorbent resin composition, it is favorable that the fusion is carried out after a surface-crosslinking treatment of the water-absorbent resin particles.

As to the second process according to the present invention for production of a water-absorbent resin composition, it is favorable that the metal compound has a melting point of not higher than 250 °C.

As to the second process according to the present invention for production of a water-absorbent resin composition, it is favorable that the metal compound is a water-soluble polyvalent metal salt having water of hydration and containing aluminum.

#### EFFECTS OF THE INVENTION

The present invention can provide a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition causes little gel-blocking and is excellent in the liquid permeability and liquid diffusibility and is high also in the absorption performances (e.g. absorption capacity without load, absorption capacity under load, capillary absorption capacity) and further is strong also against the physical damage such as during the actual production or practical use. Also, the present invention can further provide a water-absorbent resin composition and its production process, wherein the water-absorbent resin composition has the following further effects, in addition to the above, of involving little segregation of the metal compound (e.g. metal salt), and being excellent also in the handling property during the moisture absorption, and further having a dust prevention effect.

#### BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1 is a schematic sectional view of a measurement apparatus as used for measuring the AAP (method A).

Fig. 2 is a schematic sectional view of a measurement apparatus as used for measuring the AAP (method B).

Fig. 3 is a schematic sectional view of a measurement apparatus as used for

measuring the SFC.

Fig. 4 is a schematic sectional view of a measurement apparatus as used for measuring the CSF.

Fig. 5 is a view (Fig. 5-(a)) obtained by taking an electron photomicrograph of the water-absorbent resin composition (22) and, as to this photomicrograph, a view (Fig. 5-(b)) obtained by taking an X-ray image photomicrograph of the sulfur element by an SEM-EDS.

Fig. 6 is a view (Fig. 6-(a)) obtained by taking an electron photomicrograph of the water-absorbent resin composition (22) and, as to this photomicrograph, a view (Fig. 6-(b)) obtained by taking an X-ray image photomicrograph of the sulfur element by an SEM-EDS.

[Explanation of the Symbols]

- 1: Porous glass plate
- 2: Glass filter
- 15 3: Conduit
- 4: Liquid storage container
- 5: Supporting ring
- 6: 0.90 mass % physiological saline solution
- 7: Balance
- 20 8: Stand
- 9: Specimen to be measured (e.g. water-absorbent resin particles or water-absorbent resin composition)
- 10: Load (0.41 kPa (0.06 psi))
- 11: Air-intake pipe
- 25 12: Conduit
- 13: Glass filter
- 14: 0.90 mass % physiological saline solution
- 15: Liquid storage container

- 16: Balance
- 17: Filter paper
- 18: Metal gauze
- 19: Plastic cylinder
- 5 20: Load (2.07 kPa (0.3 psi))
- 21: Load (4.83 kPa (0.7 psi))
- 31: Tank
- 32: Glass tube
- 33: 0.69 mass % aqueous sodium chloride solution
- 10 34: L-tube having cock
- 35: Cock
- 40: Receptacle
- 41: Cell
- 42: Stainless metal gauze
- 15 43: Stainless metal gauze
- 44: Swollen gel
- 45: Glass filter
- 46: Piston
- 47: Holes in piston
- 20 48: Collecting receptacle
- 49: Balance
- 100: Plastic supporting cylinder
- 101: Stainless metal gauze of 400 meshes
- 102: Swollen gel
- 25 103: Piston
- 104: Load (weight)
- 105: Petri dish
- 106: Glass filter plate

107: Filter paper

108: 0.90 mass % physiological saline solution

#### DETAILED DESCRIPTION OF THE INVENTION

Hereinafter, the present invention is described in detail. Incidentally, the  
5 water-absorbent resin composition according to the present invention is a  
composition comprising a water-absorbent resin (water-absorbent resin particles) as  
the main component, and is a particulate composition comprising the  
water-absorbent resin in an amount of favorably 80 to 100 mass % (or weight %: in  
the present invention, the weight and the mass have the same meaning, and their uses  
10 herein are unified into the mass), more favorably 90 to 100 mass %, and is used  
favorably for sanitary materials (e.g. disposable diapers, sanitary napkins, incontinent  
pads, and medical pads).

[Water-absorbent resin particles]:

The water-absorbent resin particles, as used in the present invention, are  
15 particles of a water-insoluble, water-swellaable, and hydrogel-formable polymer  
(which may hereinafter be referred to as water-absorbent resin) obtainable by a  
process including the step of polymerizing a hydrophilic monomer, and has an  
absorption capacity of at least not less than 10 times for a physiological saline  
solution, and is the shape of spherical or irregular particles. Incidentally, in the  
20 present invention, the water-absorbent resin particles may be referred to simply as  
water-absorbent resin.

Specific examples of the water-insoluble, water-swellaable, and  
hydrogel-formable polymer include: partially-neutralized and crosslinked polymers  
of poly(acrylic acids) (e.g. USP 4,625,001, USP 4,654,039, USP 5,250,640, USP  
25 5,275,773, EP 0456136); crosslinked and partially-neutralized graft polymers of  
starch-acrylic acid (USP 4,076,663); copolymers of isobutylene-maleic acid (USP  
4,389,513); saponified copolymers of vinyl acetate-acrylic acid (USP 4,124,748);  
hydrolyzed (co)polymers of acrylamide (USP 3,959,569); and hydrolyzed polymers

- of acrylonitrile (USP 3,935,099). The water-absorbent resin, as used in the present invention, is favorably a water-absorbent resin including a crosslinked poly(acrylic acid) (salt) polymer obtained by a process including the step of polymerizing a monomer including acrylic acid and/or its salt. The crosslinked poly(acrylic acid) (salt) polymer in the present invention is a crosslinked polymer obtained by a process including the step of polymerizing a monomer including acrylic acid and/or its salt in an amount of not smaller than 50 mol %, favorably not smaller than 70 mol %, more favorably not smaller than 90 mol %. In addition, it is favorable that 50 to 90 mol %, preferably 60 to 80 mol %, of acid groups in the polymer are neutralized.
- As examples of the salt, there can be cited such as: alkaline metal (e.g. sodium, potassium, lithium) salts, ammonium salts, and amine salts. The neutralization of the water-absorbent resin for forming the salt may be carried out in a monomer state before the polymerization, or may be carried out in a polymer state on the way of or after the polymerization, or may be carried out both in these states.
- The crosslinked poly(acrylic acid) (salt) polymer, which is a water-absorbent resin as favorably used in the present invention, may be a copolymer obtained by copolymerizing another monomer jointly with the monomer used as the main component (acrylic acid and/or its salt), if necessary. Specific examples of the above other monomer include: anionic unsaturated monomers (e.g. methacrylic acid, maleic acid, vinylsulfonic acid, styrenesulfonic acid, 2-(meth)acrylamido-2-methylpropanesulfonic acid, 2-(meth)acryloylethanesulfonic acid, 2-(meth)acryloylpropanesulfonic acid) and their salts; nonionic-hydrophilic-group-containing unsaturated monomers (e.g. acrylamide, methacrylamide, N-ethyl(meth)acrylamide, N-n-propyl(meth)acrylamide, N-isopropyl(meth)acrylamide, N,N-dimethyl(meth)acrylamide, 2-hydroxyethyl (meth)acrylate, 2-hydroxypropyl (meth)acrylate, methoxypolyethylene glycol (meth)acrylate, polyethylene glycol mono(meth)acrylate, vinylpyridine, N-vinylpyrrolidone, N-acryloylpiperidine, N-acryloylpyrrolidine, N-vinylacetamide);

and cationic unsaturated monomers (e.g. N,N-dimethylaminoethyl (meth)acrylate, N,N-diethylaminoethyl (meth)acrylate, N,N-dimethylaminopropyl (meth)acrylate, N,N-dimethylaminopropyl(meth)acrylamide, and their quaternary salts). The amount of these monomers used as monomers other than acrylic acid and/or its salt is  
5 favorably in the range of 0 to 30 mol %, more favorably 0 to 10 mol %, of the entire monomers.

The water-absorbent resin, as used in the present invention, is a crosslinked polymer having a internal crosslinked structure.

As to methods for introducing the internal crosslinked structure into the  
10 water-absorbent resin as used in the present invention, examples thereof include: a method in which the introduction is carried out by self-crosslinking without any crosslinking agent; and a method in which the introduction is carried out by copolymerization or reaction with an internal-crosslinking agent having at least two polymerizable unsaturated groups and/or at least two reactive groups per molecule.  
15 A favorable example is the method in which the introduction is carried out by copolymerization or reaction with the internal-crosslinking agent.

Specific examples of these internal-crosslinking agents include:  
N,N'-methylenebis(meth)acrylamide, (poly)ethylene glycol di(meth)acrylate, (poly)propylene glycol di(meth)acrylate, trimethylolpropane tri(meth)acrylate,  
20 trimethylolpropane di(meth)acrylate, glycerol tri(meth)acrylate, glycerol acrylate methacrylate, ethylene-oxide-modified trimethylolpropane tri(meth)acrylate, pentaerythritol tetra(meth)acrylate, dipentaerythritol hexa(meth)acrylate, triallyl cyanurate, triallyl isocyanurate, triallyl phosphate, triallylamine, poly(meth)allyloxyalkanes, (poly)ethylene glycol diglycidyl ether, glycerol  
25 diglycidyl ether, ethylene glycol, polyethylene glycol, propylene glycol, glycerol, pentaerythritol, ethylenediamine, polyethylenimine, and glycidyl (meth)acrylate. These internal-crosslinking agents may be used either alone respectively or in combinations with each other. Above all, from the viewpoint of such as water



absorption properties of the obtained water-absorbent resin, it is favorable that a compound having at least two polymerizable unsaturated groups is essentially used as the internal-crosslinking agent. The amount of the above internal-crosslinking agent as used is favorably in the range of 0.005 to 3 mol %, more favorably 0.01 to 1.5 mol %, relative to the entire monomers.

When the polymerization is carried out, there can be added such as: hydrophilic polymers (e.g. starch, cellulose, starch derivatives, cellulose derivatives, polyvinyl alcohol, poly(acrylic acid) (salts), and crosslinked poly(acrylic acid) (salts)); and chain transfer agents such as hypophosphorous acid (salts).

When the above monomer including acrylic acid and/or its salt as the major component is polymerized to obtain the water-absorbent resin used in the present invention, then bulk polymerization, reversed-phase suspension polymerization, or precipitation polymerization may be carried out, but, from the viewpoint of the performance or the easiness in controlling the polymerization, it is favorable to carry out aqueous solution polymerization in which the monomer is used in the form of an aqueous solution. Such polymerization methods are disclosed in such as USP 4,625,001, USP 4,769,427, USP 4,873,299, USP 4,093,776, USP 4,367,323, USP 4,446,261, USP 4,683,274, USP 4,690,996, USP 4,721,647, USP 4,738,867, USP 4,748,076, and EP 1178059.

When the polymerization is carried out, there may, for example, be used the following: radical polymerization initiators such as potassium persulfate, ammonium persulfate, sodium persulfate, t-butyl hydroperoxide, hydrogen peroxide, and 2,2'-azobis(2-amidinopropane) dihydrochloride; and active energy rays such as ultraviolet rays and electron beams. In addition, in the case where the radical polymerization initiators are used, they may be used jointly with reducing agents such as sodium sulfite, sodium hydrogensulfite, ferrous sulfate, and L-ascorbic acid to carry out redox polymerization. The amount of these polymerization initiators as used is favorably in the range of 0.001 to 2 mol %, more favorably 0.01 to 0.5 mol %,

relative to the entire monomers.

The shape of the water-absorbent resin, obtained by the above polymerization, is generally such as irregularly pulverized shape, spherical shape, fibrous shape, bar shape, approximately spherical shape, or flat shape. However, the water-absorbent resin as used in the present invention is, desirably, particulate. If a water-absorbent resin of the irregularly pulverized shape as obtained by pulverization after drying is used, there are advantages in that the effects of the present invention are more enhanced.

The water-absorbent resin, as used in the present invention, is favorably that of which the surfaces and their neighborhood have further been crosslinked with a surface-crosslinking agent.

Examples of the surface-crosslinking agent usable for the surface-crosslinking treatment include: organic surface-crosslinking agents which have at least two functional groups reactable with a functional group (particularly, a carboxyl group) of the water-absorbent resin; and polyvalent metal compounds. Examples thereof include: polyhydric alcohol compounds (e.g. ethylene glycol, diethylene glycol, propylene glycol, triethylene glycol, tetraethylene glycol, polyethylene glycol, 1,3-propanediol, dipropylene glycol, 2,2,4-trimethyl-1,3-pentanediol, polypropylene glycol, glycerol, polyglycerol, 2-butene-1,4-diol, 1,3-butanediol, 1,4-butanediol, 1,5-pentanediol, 1,6-hexanediol, 1,2-cyclohexanedimethanol, 1,2-cyclohexanol, trimethylolpropane, diethanolamine, triethanolamine, polyoxypropylene, oxyethylene-oxypropylene block copolymers, pentaerythritol, and sorbitol); epoxy compounds (e.g. ethylene glycol diglycidyl ether, polyethylene glycol diglycidyl ether, glycerol polyglycidyl ether, diglycerol polyglycidyl ether, polyglycerol polyglycidyl ether, propylene glycol diglycidyl ether, polypropylene glycol diglycidyl ether, and glycidol); polyamine compounds (e.g. ethylenediamine, diethylenetriamine, triethylenetetramine, tetraethylenepentamine, pentaethylenhexamine, and polyethylenimine) and their inorganic or organic salts

(e.g. azetidinium salts); polyisocyanate compounds (e.g. 2,4-tolylene diisocyanate and hexamethylene diisocyanate); polyoxazoline compounds (e.g. 1,2-ethylenebisoxazoline); carbonic acid derivatives (e.g. urea, thiourea, guanidine, dicyandiamide, 2-oxazolidinone); alkylene carbonate compounds (e.g. 1,3-dioxolan-2-one, 4-methyl-1,3-dioxolan-2-one, 4,5-dimethyl-1,3-dioxolan-2-one, 4,4-dimethyl-1,3-dioxolan-2-one, 4-ethyl-1,3-dioxolan-2-one, 4-hydroxymethyl-1,3-dioxolan-2-one, 1,3-dioxan-2-one, 4-methyl-1,3-dioxan-2-one, 4,6-dimethyl-1,3-dioxan-2-one, and 1,3-dioxopan-2-one); haloepoxy compounds (e.g. epichlorohydrin, epibromohydrin, and  $\alpha$ -methylepichlorohydrin) and their polyamine-added products (e.g. Kymene (registered trademark) produced by Hercules); silane coupling agents (e.g.  $\gamma$ -glycidoxypyrpyltrimethoxysilane and  $\gamma$ -aminopropyltriethoxysilane); oxetane compounds (e.g. 3-methyl-3-oxetanemethanol, 3-ethyl-3-oxetanemethanol, 3-butyl-3-oxetanemethanol, 3-methyl-3-oxetaneethanol, 3-ethyl-3-oxetaneethanol, 3-butyl-3-oxetaneethanol, 3-chloromethyl-3-methyloxetane, 3-chloromethyl-3-ethyloxetane, and polyoxetane compounds); and polyvalent metallic compounds (e.g. hydroxides and chlorides of such as zinc, calcium, magnesium, aluminum, iron and zirconium). These surface-crosslinking agents may be used either alone respectively or in combinations with each other. Above all, the polyhydric alcohols are favorable, because they are high in safety and can enhance the hydrophilicity of water-absorbent resin particle surfaces. In addition, the use of the polyhydric alcohols enhance the affinity of water-absorbent resin particle surfaces to the polyvalent metal particles, so that interactions between the polyhydric alcohol residue and the polyvalent metal surface enables more uniform existence of the polyvalent metal particles on surfaces of the water-absorbent resin particles.

The amount of the surface-crosslinking agent, as used, is favorably in the range of 0.001 to 5 mass parts, per 100 mass parts of the solid content of the

water-absorbent resin.

When the surface-crosslinking agent and the water-absorbent resin are mixed together, water may be used. The amount of water, as used, is favorably larger than 0.5 but not larger than 10 mass parts, more favorably in the range of 1 to 5 mass parts, per 100 mass parts of the solid content of the water-absorbent resin.

When the surface-crosslinking agent and/or its aqueous solution is mixed, a hydrophilic organic solvent and/or a third substance may be used as a mixing assistant.

In the case where the hydrophilic organic solvent is used, its examples include:

10 lower alcohols (e.g. methyl alcohol, ethyl alcohol, n-propyl alcohol, isopropyl alcohol, n-butyl alcohol, isobutyl alcohol, and t-butyl alcohol); ketones (e.g. acetone); ethers (e.g. dioxane, tetrahydrofuran, and methoxy(poly)ethylene glycol); amides (e.g.  $\epsilon$ -caprolactam and N,N-dimethylformamide); sulfoxides (e.g. dimethyl sulfoxide); and polyhydric alcohols (e.g. ethylene glycol, diethylene glycol, propylene glycol, triethylene glycol, tetraethylene glycol, polyethylene glycol, 1,3-propanediol, dipropylene glycol, 2,2,4-trimethyl-1,3-pentanediol, polypropylene glycol, glycerol, polyglycerol, 2-butene-1,4-diol, 1,3-butanediol, 1,4-butanediol, 1,5-pentanediol, 1,6-hexanediol, 1,2-cyclohexanedimethanol, 1,2-cyclohexanol, trimethylolpropane, diethanolamine, triethanolamine, polyoxypropylene, oxyethylene-oxypropylene block copolymers, pentaerythritol, and sorbitol).

15 20 Though depending on such as kind, particle diameters, and water content of the water-absorbent resin, the amount of the hydrophilic organic solvent as used is favorably not larger than 10 mass parts, more favorably in the range of 0.1 to 5 mass parts, per 100 mass parts of the solid content of the water-absorbent resin. In addition, as the third substance, there may be caused to coexist those which are disclosed in EP 0668080, such as inorganic acids, organic acids, and polyamino acids. These mixing assistants may act as surface-crosslinking agents, but are favorably those which do not give a surface-crosslinked water-absorbent resin having low

water absorption performance. Particularly, volatile alcohols having boiling points of lower than 150 °C are desirable in that they volatilize during the surface-crosslinking treatment and thus their residues do not remain.

When the water-absorbent resin and the surface-crosslinking agent are mixed together, there may be caused to coexist a noncrosslinkable water-soluble inorganic base (favorably, alkaline metal salts, ammonium salts, alkaline metal hydroxides, and ammonia or its hydroxide) and/or an irreducible alkaline-metal-salt pH buffer (favorably such as hydrogencarbonates, dihydrogenphosphates, and hydrogenphosphates) for the purpose of more uniformly mixing the water-absorbent resin and the surface-crosslinking agent together. The amount of these materials, as used, depends upon such as type or particle diameters of the water-absorbent resin, but is favorably in the range of 0.005 to 10 mass parts, more favorably 0.05 to 5 mass parts, per 100 mass parts of the solid content of the water-absorbent resin.

Although not especially limited, the method for mixing the water-absorbent resin and the surface-crosslinking agent together can be exemplified by such as: a method in which the water-absorbent resin is immersed into the hydrophilic organic solvent and then mixed with the surface-crosslinking agent that is, if necessary, dissolved in water and/or a hydrophilic organic solvent; and a method in which the surface-crosslinking agent that is dissolved in water and/or the hydrophilic organic solvent is spraywise or dropwise added directly to the water-absorbent resin to mix them together.

After the mixing of the water-absorbent resin and the surface-crosslinking agent, usually, a heating treatment is carried out to conduct the crosslinking reaction. Though depending on the surface-crosslinking agent as used, the temperature of the above heating treatment is favorably in the range of 40 to 250 °C, more favorably 150 to 250 °C. In the case where the treatment temperature is lower than 40 °C, the absorption properties such as absorption capacity under a load are sometimes not sufficiently improved. In the case where the treatment temperature is higher than

250 °C, the deterioration of the water-absorbent resin is sometimes caused, so that the performance is lowered, therefore caution is needed. The duration of the heating treatment is favorably in the range of 1 minute to 2 hours, more favorably 5 minutes to 1 hour.

5        There is no especial limitation on the particle diameters and particle diameter distribution of the water-absorbent resin as used in the present invention. However, if there is used a water-absorbent resin having comparatively small particle diameters and a particle diameter distribution such that the content of components having small particle diameters is high, then there are advantages in that the water absorption  
10       performances such as water absorption rate and capillary absorption capacity are greatly enhanced.

As to the water-absorbent resin as used in the present invention, it is favorable for enhancing the performances such as water absorption rate and capillary absorption capacity that the mass-average particle diameter is not larger than 500  $\mu\text{m}$ ,  
15       more favorably not larger than 400  $\mu\text{m}$ . In addition, the ratio of particles having particle diameters of smaller than 300  $\mu\text{m}$  in the water-absorbent resin is favorably not less than 10 mass %, more favorably not less than 30 mass %, still more favorably not less than 50 mass %, relative to the water-absorbent resin. A water-absorbent resin having such particle diameters either is that which is obtained  
20       by pulverizing a water-absorbent resin obtained by the aqueous solution polymerization, or can favorably be obtained by sieving the pulverized water-absorbent resin to thus regulate its particle diameters. In addition, it is also possible to use a water-absorbent resin obtained by a process including the steps of: agglomerating a fine powder of a water-absorbent resin having particle diameters of  
25       not larger than 300  $\mu\text{m}$ ; and then regulating the particle diameters of the agglomerated water-absorbent resin. Furthermore, it is also possible to use a water-absorbent resin obtained by a process in which irregularly pulverized particles which are primary particles obtained by pulverization are partially mixed with the

agglomerated material of the fine powder. In the case of having been partially mixed with the agglomerated material of the water-absorbent resin, there can be obtained a water-absorbent resin composition according to the present invention which is still more excellent in the absorption properties such as water absorption rate and capillary absorption capacity. The amount of the agglomerated material of the fine powder, which is mixed, is favorably not smaller than 5 mass %, more favorably not smaller than 10 mass %, still more favorably not smaller than 15 mass %.

As methods for preparing the agglomerated material of the fine powder, it is possible to use publicly known arts to regenerate a fine powder. Examples of such usable arts include methods in which: warm water and a fine powder of a water-absorbent resin are mixed together and then dried (USP 6,228,930); a fine powder of a water-absorbent resin is mixed with an aqueous monomer solution, and then the resultant mixture is polymerized (USP 5,264,495); water is added to a fine powder of a water-absorbent resin, and then the resultant mixture is agglomerated under not less than a specific face pressure (EP 0844270); a fine powder of a water-absorbent resin is sufficiently wetted to thus form an amorphous gel, and then this gel is dried and pulverized (USP 4,950,692); and a fine powder of a water-absorbent resin and a polymer gel are mixed together (USP 5,478,879). However, there is favorably used the aforementioned method in which warm water and a fine powder of a water-absorbent resin are mixed together and then dried. Incidentally, the particle diameter is indicated by the sieve mesh opening size of the classification.

[Water-soluble polyvalent metal salt particles]:

It has already been commonly known that a certain kind of inorganic compound particles prevent the gel-blocking to thus provide a water-absorbent resin with high distributing and diffusing abilities. However, the present inventors studied about a water-absorbent resin composition which displays high performance

even in the case where, as is aforementioned, a water-absorbent resin has been damaged by such as pneumatic transportation which is usually included in processes for production of water-absorbent resins. As a result, the present inventors have found out that: unexpectedly, in the case where the water-soluble polyvalent metal salt particles are used as the inorganic compound particles, the effect of enhancing the liquid permeation rate under load is great and further, only in the case where the water-soluble polyvalent metal salt is not added to a water-absorbent resin in a state of an aqueous solution, but the water-soluble polyvalent metal salt is added to water-absorbent resin particles in a state of particles, there is obtained a water-absorbent resin composition which is excellent both in a performance of retaining the liquid permeation rate under load for a long time and in the physical damage resistance. In addition, the present inventors have found out also that: in the case where the water-soluble polyvalent metal salt particles are hydrous salt crystals, particularly the effects are great. As to the reason why the water-absorbent resin composition obtained by the dry mixing displays such effects, it is inferred (from great deterioration of the effects of the present invention as to a composition being in a state where the water-soluble polyvalent metal salt particles are dissolved to thus be in a non-particulate state by adding the water-soluble polyvalent metal salt to a water-absorbent resin in the form of an aqueous solution or adding water to a dry wise mixture of the water-absorbent resin particles and the water-soluble polyvalent metal salt particles) that: if the water-soluble polyvalent metal salt is intermingled with the water-absorbent resin particles in the form of particles, then, when the physical damage such as impact is done to the water-absorbent resin composition, the water-soluble polyvalent metal salt particles absorb the impact energy to thus reduce the damage to the water-absorbent resin. Hereupon, the impact energy can be considered as being consumed by pulverization of the water-soluble polyvalent metal salt particles and by uniformization due to rearrangement of the water-soluble polyvalent metal salt particles. Therefore, it can be considered as desirable that the



water-soluble polyvalent metal salt particles are in a state of the dry wise mixture which can move with some degree of freedom rather than are entirely fixed on surfaces of the water-absorbent resin.

In addition to the above, the water-soluble polyvalent metal salt has an action  
5 of hydrophilizing the surfaces of the water-absorbent resin particles and, when the water-absorbent resin composition absorbs an aqueous liquid, the water-soluble polyvalent metal salt particles dissolve to thus make actions of ion-crosslinking the surfaces of the water-absorbent resin and keeping the spaces between water-absorbent resins wide. These actions have the effect of enhancing the liquid  
10 permeation rate under load. Hereupon, this effect is greater when the polyvalent metal exists in the periphery of and/or near the surfaces of the water-absorbent resin particles than when the polyvalent metal exists inside the water-absorbent resin particles. As to the water-absorbent resin composition as produced by adding the water-soluble polyvalent metal salt to a water-absorbent resin in the form of an  
15 aqueous solution or adding water to a dry wise mixture of the water-absorbent resin particles and the water-soluble polyvalent metal salt particles, much of the polyvalent metal salt has already permeated the inside of the water-absorbent resin, and therefore the effects of the polyvalent metal salt upon water-absorbent resin surfaces are low during the absorption of an aqueous liquid such as urine. As a result, the  
20 above water-absorbent resin composition is low in performance, particularly, in the liquid permeation rate under load, and its durability is also bad. In addition, because the permeated polyvalent metal salt reacts with the carboxyl group to thus form a crosslinked structure, the liquid permeability is deteriorated due to the damage done by the process. In comparison, because the water-soluble polyvalent  
25 metal salt usable in the present invention is mixed with the water-absorbent resin particles in the form of particles, the water-soluble polyvalent metal salt does not dissolve to act on the water-absorbent resin surfaces until the water-absorbent resin composition absorbs urine or an aqueous liquid. This action can more efficiently

make the effects of the polyvalent metal salt upon the water-absorbent resin surfaces endure for a long time. In addition, the water-absorbent resin composition according to the present invention, in which the water-soluble polyvalent metal salt particles exist on surfaces of the water-absorbent resin particles, is excellent in the liquid permeability still after having been damaged by the process, because the water-soluble polyvalent metal salt particles exist on surfaces of the water-absorbent resin particles still after the composition has been damaged by the process. That it to say, the water-absorbent resin composition according to the present invention can be said to have a structure which effectively exercises the effects of the polyvalent metal without deteriorating the liquid absorption performance of the water-absorbent resin itself.

The water-soluble polyvalent metal salt particles, usable in the present invention, are particles of a salt of a metal having a valence of not less than 2 and are powdery. Taking it into consideration that the water-absorbent resin composition according to the present invention is utilized for absorbent structures for sanitary materials such as diapers, then it is favorable to select a water-soluble polyvalent metal salt which does not color the water-absorbent resin composition and which is low poisonous to human bodies.

For the purpose of more efficiently making the effects of the polyvalent metal salt endure for a long time during the liquid absorption, there is favorably selected and used the polyvalent metal salt which can dissolve in a concentration of not less than 5 mass %, more favorably not less than 10 mass %, still more favorably not less than 20 mass %, in pure water of normal temperature.

Examples of the water-soluble polyvalent metal salt particles, usable in the present invention, include aluminum chloride, poly(aluminum chloride), aluminum sulfate, aluminum nitrate, potassium aluminum bis(sulfate), sodium aluminum bis(sulfate), calcium chloride, calcium nitrate, magnesium chloride, magnesium sulfate, magnesium nitrate, zinc chloride, zinc sulfate, zinc nitrate, zirconium

chloride, zirconium sulfate, and zirconium nitrate. In addition, also from the viewpoint of solubility into a liquid being absorbed such as urine, it is favorable to use these salts having water of crystallization. Particularly favorable are the aluminum compounds, above all, aluminum sulfate. Powders of hydrous crystals  
5 such as aluminum sulfate octadecahydrate and aluminum sulfate tetradeca- to octadecahydrates can most favorably be used.

From the viewpoint of mixability, it is favorable that the particle diameters of the water-soluble polyvalent metal salt particles, usable in the present invention, are smaller than those of the water-absorbent resin. The mass-average particle diameter  
10 is favorably not larger than 500  $\mu\text{m}$ , more favorably not larger than 400  $\mu\text{m}$ . From the viewpoint of performance, the water-soluble polyvalent metal salt particles include particles having particle diameters of not larger than 150  $\mu\text{m}$  in an amount of more favorably not smaller than 20 mass %, most favorably not smaller than 30 mass %, relative to the water-soluble polyvalent metal salt particles.

15 As the behavior and state of the water-soluble polyvalent metal salt particles usable in the present invention, it is favorable from the viewpoint of damage mitigation that the particles are such as agglomerates. The bulk density is favorably not less than 0.5  $\text{g}/\text{cm}^3$ , more favorably not less than 0.7  $\text{g}/\text{cm}^3$ .

[Water-absorbent resin composition (1)]:

20 The water-absorbent resin composition (1) according to the present invention is a water-absorbent resin composition comprising water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt,

with the composition having a mass-average particle diameter of 100 to 600  $\mu\text{m}$  and comprising water-soluble polyvalent metal salt particles and the  
25 water-absorbent resin particles that have been surface-crosslinked.

Also, the water-absorbent resin composition (1) according to the present invention is a water-absorbent resin composition comprising water-absorbent resin particles and water-soluble polyvalent metal salt particles, wherein the

water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt,

with the composition of which the saline flow conductivity (SFC) is at least 50 ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ ) and of which the retention ratio of the saline flow conductivity (SFC) is not less than 40 %.

The water-absorbent resin composition (1) according to the present invention comprises the water-absorbent resin particles as the main component and further comprises the water-soluble polyvalent metal salt particles, and is usually particulate and can be used favorably as an absorbent material for sanitary materials for absorption of urine, menstrual blood, sweat, and other body fluids.

Because the water-soluble polyvalent metal salt particles in the water-absorbent resin composition (1) according to the present invention are merely mixed with the water-absorbent resin particles in the form of particles, the water-soluble polyvalent metal salt particles exist on surfaces of the water-absorbent resin particles while still keeping their particulate shape, or exist in the periphery of the water-absorbent resin particles, for example, between particles of the water-absorbent resin, while still keeping their particulate shape. A favorable mode is a state where the water-soluble polyvalent metal salt particles substantially coexist uniformly with the water-absorbent resin particles. If the water-soluble polyvalent metal salt particles can maintain the particle shape in order to form such a state, then the mode may be that the water-soluble polyvalent metal salt particles are made to adhere weakly to the water-absorbent resin particles with such as a binder. It may be possible to observe these states from photographs taken with such as electron microscopes. However, they can be confirmed by dispersing and stirring the water-absorbent resin composition into an appropriate organic solvent or an appropriate gas and then separating the water-absorbent resin particles and the water-soluble polyvalent metal salt particles from each other by utilizing the difference between their specific gravities.

The water-absorbent resin composition (1) according to the present invention, favorably, includes the water-absorbent resin particles and the water-soluble polyvalent metal salt particles. At least a part of the aforementioned water-soluble polyvalent metal salt particles adhere weakly to water-absorbent resin particle surfaces by the binder. Therefore, the permeation of the polyvalent metal into the water-absorbent resin particles is effectively prevented, and further, the polyvalent metal is fixed all over the surfaces of the water-absorbent resin particles uniformly and moderately (i.e. in a state where the fixation is incomplete, but is not so weak as to enable free migration). Consequently, the gel-blocking can sufficiently be prevented, and therefore excellent liquid permeability and liquid diffusibility can be displayed and also excellent absorption performances can be displayed. Furthermore, the water-absorbent resin composition comes into a state which is strong also against the physical damage such as during the actual production or practical use. These states can also be observed from photographs taken with such as electron microscopes.

The water-absorbent resin particles, which are included in the water-absorbent resin composition (1) according to the present invention, are, favorably, surface-crosslink-treated ones.

The water-absorbent resin composition (1), according to the present invention, is in the form of particles having a mass-average particle diameter in the range of favorably 100 to 600  $\mu\text{m}$ , more favorably 200 to 500  $\mu\text{m}$ . In the case where the mass-average particle diameter is smaller than 100  $\mu\text{m}$ , then, even if the water-soluble polyvalent metal salt particles are added, there is a possibility that it may be difficult to obtain the effects of the present invention. In the case where the mass-average particle diameter is larger than 600  $\mu\text{m}$ , then there is a possibility that the water-soluble polyvalent metal salt particles may segregate to deteriorate the uniform mixability. In addition, the ratio of particles having particle diameters of smaller than 300  $\mu\text{m}$  in the water-absorbent resin composition (1) is favorably not

less than 10 mass %, more favorably not less than 30 mass %, still more favorably not less than 50 mass %, relative to the water-absorbent resin composition (1).

The water-absorbent resin composition (1), according to the present invention, favorably displays an absorption capacity without load (CRC) of not less than 20 (g/g), more favorably not less than 22 (g/g), still more favorably not less than 24 (g/g), yet still more favorably not less than 25 (g/g), particularly favorably not less than 27 (g/g). In the case where the absorption capacity without load (CRC) is less than 20 (g/g), the absorption efficiency is bad on an occasion of the use for sanitary materials such as diapers.

10 The water-absorbent resin composition (1), according to the present invention, favorably displays an absorption capacity under load (AAP) of not less than 16 (g/g), more favorably not less than 18 (g/g), still more favorably not less than 20 (g/g), yet still more favorably not less than 22 (g/g), particularly favorably not less than 24 (g/g), under a load of 0.7 psi. In the case where the absorption capacity under load 15 (AAP) is less than 16 (g/g), the absorption efficiency is bad on an occasion of the use for sanitary materials such as diapers.

The water-absorbent resin composition (1), according to the present invention, favorably displays a saline flow conductivity (SFC) of not less than  $50 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , more favorably not less than  $100 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , still more favorably not less than  $120 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , particularly favorably not less than  $150 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ . The saline flow conductivity (SFC) depends on the content of the water-absorbent resin composition in the sanitary material. The higher content needs the higher saline flow conductivity (SFC).

As to the water-absorbent resin composition (1) according to the present invention, it is desirable that the deterioration of the absorption capacity under load 25 (AAP) of this water-absorbent resin composition, as compared with an absorption capacity under load (AAP) (under the same load) of the water-absorbent resin particles to which the water-soluble polyvalent metal salt particles have not yet been

added, is small. The water-absorbent resin composition favorably maintains an absorption capacity under load of not less than 0.85 time, more favorably not less than 0.90 time, still more favorably not less than 0.95 time, in comparison with the absorption capacity under load (AAP) of the water-absorbent resin particles.

5        The water-absorbent resin composition (1), according to the present invention, displays an effect such that the deterioration of absorption performances is small even if it undergoes various physical energy (damage) during the production or practical use. That is to say, the water-absorbent resin composition (1), according to the present invention, is a water-absorbent resin composition which has the  
10        following absorption performances after physical energy has worked against the composition.

      The water-absorbent resin composition (1), according to the present invention, favorably displays an absorption capacity without load after the paint shaker test (shaking at 800 cycles/min for 30 minutes) (CRC after PS) of not less than 20 (g/g),  
15        more favorably not less than 22 (g/g), still more favorably not less than 24 (g/g), yet still more favorably not less than 25 (g/g), particularly favorably not less than 27 (g/g). In the case where the absorption capacity without load after the paint shaker test (CRC after PS) is less than 20 (g/g), the absorption efficiency is bad on an occasion of the use for sanitary materials such as diapers.

20        The water-absorbent resin composition (1), according to the present invention, favorably displays a saline flow conductivity after the paint shaker test (SFC after PS) of not less than  $50 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , more favorably not less than  $100 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , still more favorably not less than  $120 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , particularly favorably not less than  $150 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ . The saline flow  
25        conductivity after the paint shaker test (SFC after PS) depends on the content of the water-absorbent resin composition in the sanitary material. The higher content needs the higher saline flow conductivity (SFC).

      As to the water-absorbent resin composition (1) according to the present

invention, the ratio of a saline flow conductivity (SFC) after the paint shaker test to an SFC before the PS test, namely, the retention ratio of the saline flow conductivity after the paint shaker test (retention ratio of SFC after PS), is favorably not less than 70 %, more favorably not less than 80 %, still more favorably not less than 90 %, particularly favorably not less than 100 %.

The water-absorbent resin composition (1), according to the present invention, can maintain a high saline flow conductivity (SFC) even when used in sanitary materials for a long time.

As to the water-absorbent resin composition (1) according to the present invention, the ratio of a saline flow conductivity (SFC) after a swelling time of 120 minutes to a saline flow conductivity (SFC) after a swelling time of 60 minutes, namely, the retention ratio of the saline flow conductivity (retention ratio of SFC), is favorably not less than 40 %, more favorably not less than 50 %, still more favorably not less than 60 %. As to conventional water-absorbent resins (or water-absorbent resin compositions) to which metal particles have been added, if they are measured for a swelling duration of more than 60 minutes in the test for the saline flow conductivity (SFC), then a rapid fall of the liquid permeation rate is seen.

The water-absorbent resin composition (1), according to the present invention, further has a feature of generating little dust. As to the water-absorbent resin composition (1) according to the present invention, the dust generation degree is favorably not more than 0.25 (mg/m<sup>3</sup>), more favorably not more than 0.23 (mg/m<sup>3</sup>), still more favorably not more than 0.20 (mg/m<sup>3</sup>), yet still more favorably not more than 0.17 (mg/m<sup>3</sup>), particularly favorably not more than 0.15 (mg/m<sup>3</sup>).

The water-absorbent resin composition (1) according to the present invention is excellent in the wettability to aqueous liquids. Particularly above all, a water-absorbent resin composition including water-absorbent resin particles which have been surface-crosslinked with the polyhydric alcohol is excellent in the wettability and contributes to the enhancement of the absorption performances. The



wettability of the water-absorbent resin composition to aqueous liquids can be evaluated by measuring the contact angle. It is not easy to precisely measure the contact angle of a liquid with a liquid-absorbent powder like the water-absorbent resin composition. However, the apparent contact angle can be measured by the below-mentioned method. The water-absorbent resin composition according to the present invention displays a contact angle of favorably not more than 45 degrees, more favorably not more than 30 degrees, particularly favorably not more than 20 degrees.

The water-absorbent resin composition (1) according to the present invention may possess such functions as given or enhanced by causing this composition to, besides the water-absorbent resin particles and the water-soluble polyvalent metal salt particles, further contain additives such as: water-insoluble finely-particulate inorganic powders (e.g. silicon dioxide, titanium dioxide, aluminum oxide, magnesium oxide, zinc oxide, talc, calcium phosphate, barium phosphate, silicic acid or its salts, clay, diatomite, zeolite, bentonite, kaolin, hydrotalcite, and salts (e.g. activated clay)); deodorants, perfumes, antibacterial agents, cationic polymer compounds (e.g. polyamines), foaming agents, pigments, dyes, manures, oxidizing agents, and reducing agents. The ratio of the additives as used is less than 10 mass %, favorably less than 5 mass %, more favorably less than 1 mass %, relative to the total of the water-absorbent resin particles and the water-soluble polyvalent metal salt particles.

[Process (1) for production of water-absorbent resin composition]:

A favorable process for production of the composition (1) according to the present invention is a dry mixing process. The dry mixing process is a process in which the water-absorbent resin particles (favorably, surface-crosslinked ones) and the water-soluble polyvalent metal salt particles are mixed together in a state where the water-soluble polyvalent metal salt particles substantially keep their dry state. Hereupon, the water-soluble polyvalent metal salt particles are mixed under

conditions where they can exist as independent particles.

When the water-absorbent resin composition (1) according to the present invention is produced or preserved, it must be avoided adding or mixing water in such an amount that the water-soluble polyvalent metal salt particles may dissolve, or  
5 putting the composition under high humidity. If the water-soluble polyvalent metal salt particles contact with the water-absorbent resin in a state dissolved in such as water, then the water-soluble polyvalent metal salt falls into a state coated to water-absorbent resin particle surfaces or permeated inside the resin to thus exist as particles no longer, so that the effects of the present invention are not sufficiently  
10 displayed. For example, in JP-A-523289/2001 (Kohyo) (WO 98/48857), there is disclosed a process for preparation of a super-water-absorbent polymer which process is characterized by including the steps of mixing a super-water-absorbent polymer with a polyvalent metal salt and then bringing the resultant mixture into close contact with a binder. Water or a water-soluble liquid is stated therein as the  
15 aforementioned binder. However, the present invention entirely differs from conventional processes in that, when the water-absorbent resin composition (1) according to the present invention is produced, an aqueous liquid (e.g. water, water-soluble liquid) does not need to be added after the mixing of the water-absorbent resin particles and the water-soluble polyvalent metal salt particles.  
20 Therefore, the water-soluble polyvalent metal salt particles which are contained in the water-absorbent resin composition (1) according to the present invention are not dissolved in a liquid, but exist substantially as dry particles along with the water-absorbent resin. It is an important method according to the present invention that the water-soluble polyvalent metal salt particles are caused to exist in a state of  
25 dry particles. Thereby there can be obtained the water-absorbent resin composition (1) which is excellent in water absorption performances such as after-damage liquid permeability.

The specific mixing method is free of especial limitation if it is a dry mixing

method. For example, publicly known methods for addition and mixing of powders are used to carry out the addition in a lump or divisionally or continuously. The addition and mixing of the water-soluble polyvalent metal salt particles may be carried out while the water-absorbent resin particles are stirred. Or the stirring operation may be carried out after the addition of the water-soluble polyvalent metal salt particles. Usable as stirring apparatuses or mixing apparatuses are such as paddle blenders, ribbon mixers, rotary blenders, jar tumblers, plowshare mixers, cylinder type mixers, V-character-shaped mixers, ribbon type mixers, screw type mixers, twin-arm mixers, pulverizing type kneaders, channel type mixers, and plow type mixers.

The ratio between the water-soluble polyvalent metal salt particles and the water-absorbent resin particles as used is favorably in the range of 0.01 to 5 mass parts, more favorably 0.1 to 2 mass parts, per 100 mass parts of the solid content of the water-absorbent resin particles. In the case where the water-soluble polyvalent metal salt particles are added too much, the performance deterioration of the water-absorbent resin is brought about. In the case where the amount of the water-soluble polyvalent metal salt particles is too small, the water-soluble polyvalent metal salt particles do not take effect.

A more favorable mode of the dry mixing processes, which are favorable processes for production of the composition (1) according to the present invention, is a mode comprising the steps of:

adding a binder to water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt; and then

mixing the binder and the water-absorbent resin particles with water-soluble polyvalent metal salt particles.

Its feature is that: the binder is beforehand added to the water-absorbent resin particles to thus put them in a state where the binder is permeated across surfaces of the water-absorbent resin particles, and then they are mixed with the water-soluble

polyvalent metal salt particles.

As to the modes as reported in JP-A-523287/2001 (Kohyo), JP-A-124879/1997 (Kokai), JP-A-270741/1988 (Kokai), and JP-A-538275/2002 (Kohyo) aforementioned as background arts, namely, as to the modes that a metal salt is formed into its aqueous solution and then added to a water-absorbent resin, there have been problems such that: the metal unfavorably goes so far as permeating into particles of the water-absorbent resin, so that there occur the following: the deterioration of the absorption capacity without load by influence of the metal inside the water-absorbent resin; and the deterioration of the absorption capacity under load, the gel-blocking, and the deteriorations of the liquid permeability and liquid diffusibility, due to insufficient presence of the metal on surfaces of the water-absorbent resin.

In addition, as to the modes as reported in JP-A-257235/1986 (Kokai) and JP-A-523289/2001 (Kohyo) aforementioned as background arts, namely, as to the modes that a water-absorbent resin is dry-blended with a metal salt and then water is added to them, there have been problems such that: the dissolved metal salt causes binding between particles to thus easily form a strong agglomerate and, in the case where this agglomerate is crushed by physical damage such as during the actual production or practical use, the absorption capacity under load is deteriorated. In addition, there have also been problems such that: the dissolved metal salt unfavorably goes so far as permeating into particles of the water-absorbent resin.

On the other hand, as to the above more favorable mode, because the water-soluble polyvalent metal salt particles being in a state of a powder is mixed with the water-absorbent resin particles in a state where the binder is permeated across surfaces of the water-absorbent resin particles, the permeation of the polyvalent metal into the water-absorbent resin particles can effectively be prevented, and further, the polyvalent metal is fixed all over the surfaces of the water-absorbent resin particles uniformly and moderately (i.e. in a state where the fixation is

incomplete, but is not so weak as to enable free migration). Thereby consequently, the gel-blocking can sufficiently be prevented, and therefore excellent liquid permeability and liquid diffusibility can be displayed and also excellent absorption performances can be displayed. Furthermore, the resultant water-absorbent resin composition comes into a state which is strong also against the physical damage such as during the actual production or practical use.

The binder usable in the present invention has a role as a binder for fixing the polyvalent metal to surfaces of the water-absorbent resin particles, and is beforehand added to the water-absorbent resin particles before the water-absorbent resin particles is mixed with the water-soluble polyvalent metal salt particles.

The binder usable in the present invention is free of especial limitation if it includes a material which can play the above role. However, for example, there can be cited those which include such as water, polyhydric alcohols, water-soluble polymers, thermoplastic resins, pressure-sensitive adhesives, and adhesives. Favorable are those which include water and/or polyhydric alcohols.

The binder usable in the present invention may contain the aforementioned surface-crosslinking agent. Particularly, as is mentioned below, if the binder is made to contain the surface-crosslinking agent in the case where not yet surface-crosslinked water-absorbent resin particles are used as the water-absorbent resin particles, then it becomes possible to carry out the surface-crosslinking treatment in the process for production of the water-absorbent resin composition (1).

The amount of the binder usable in the present invention is favorably in the range of 0.1 to 10 mass %, more favorably 0.1 to 5 mass %, still more favorably 0.2 to 3 mass %, relative to the solid content of the water-absorbent resin particles. In the case where the amount of the binder is smaller than 0.1 mass %, there is a possibility that the moderate fixation of the polyvalent metal cannot be realized. In the case where the amount of the binder is larger than 10 mass %, there is a possibility that the properties of the resultant water-absorbent resin composition (1)

may be deteriorated.

The method for adding the binder to the water-absorbent resin particles is not especially limited. However, there is preferred a method which enables uniform addition of the binder to the water-absorbent resin particles to mix them together.

5 For instance, it can be exemplified by such as a method in which the binder is spraywise or dropwise added directly to the water-absorbent resin particles to mix them together. Examples of apparatuses for the mixing include cylinder type mixers, V-character-shaped mixers, ribbon type mixers, screw type mixers, twin-arm mixers, pulverizing type kneaders, channel type mixers, and plow type mixers.

10 When the binder is added to the water-absorbent resin particles, it is favorable to beforehand adjust the temperature of the water-absorbent resin particles in the range of 40 to 100 °C, more favorably 50 to 90 °C, still more favorably 60 to 80 °C. In the case of deviating from the above temperature ranges, it is difficult that the binder is uniformly added to the water-absorbent resin particles to mix them together.

15 In the above more favorable mode, either surface-crosslinked water-absorbent resin particles or not yet surface-crosslinked water-absorbent resin particles may be used as the water-absorbent resin particles. If the binder is made to contain the surface-crosslinking agent in the case where the not yet surface-crosslinked water-absorbent resin particles are used as the water-absorbent resin particles, then it becomes possible to carry out the surface-crosslinking treatment in the process for  
20 production of the water-absorbent resin composition (1).

In the above more favorable mode, the water-absorbent resin particles to which the binder has been added is mixed with the water-soluble polyvalent metal salt particles.

25 The method for mixing the water-absorbent resin particles with the water-soluble polyvalent metal salt particles is the same as the aforementioned.

In the process (1) according to the present invention for production of a water-absorbent resin composition, the stirring operation is not necessarily needed

when the water-soluble polyvalent metal salt particles are added to the water-absorbent resin particles. After the water-soluble polyvalent metal salt particles have been added to the water-absorbent resin particles, the step in which physical energy such as impact works against water-absorbent resins (which step is usually included in processes for production of the water-absorbent resins) may be utilized still in a state of non-uniform mixing, thus carrying out the mixing.

A favorable mode for carrying out the present invention is a process which utilizes energy that works against the water-absorbent resin when, as often adopted in processes for production of water-absorbent resins, the water-absorbent resin particles (powder) are pneumatically transported. That is to say, if the water-soluble polyvalent metal salt particles are added in a step as carried out before the step of pneumatically transporting the water-absorbent resin powder and if thereafter the pneumatic transportation is carried out, then the water-absorbent resin composition which is excellent in such as absorption performances can be obtained without needing to be processed with any special mixer or pulverizer.

The pneumatic transportation of the water-absorbent resin particles (powder) in the process for production of the water-absorbent resin is usually carried out over a transportation distance of 10 to 200 m and at a transportation speed of 0.1 to 15 m/second. By passing through the pneumatic transportation step, the water-absorbent resin particles and the water-soluble polyvalent metal salt particles are pulverized and mixed together due to contact and collision between particles or due to collision of particles with walls of the transportation course, so that the water-absorbent resin composition according to the present invention can be obtained.

The energy which the water-absorbent resin particles and the water-soluble polyvalent metal salt particles undergo during the pneumatic transportation corresponds to energy which they undergo when being shaken with a paint shaker in a state where they are placed in a container together with glass beads. Therefore,

the energy to be actually applied can be reproduced with the paint shaker in a laboratory (paint shaker test, which may hereinafter be abbreviated as PS). The paint shaker test (PS) is carried out as follows: a glass container of 6 cm in diameter and 11 cm in height is charged with 10 g of glass beads of 6 mm in diameter and 30 g of the water-absorbent resin composition, and then attached to a paint shaker (product No. 488, produced by Toyo Seiki Seisakusho K.K.), and then shaken at 800 cycles/min (CPM) for 30 minutes. The details of the apparatus are disclosed in JP-A-235378/1997 (Kokai). The energy which is applied to the water-absorbent resin particles and the water-soluble polyvalent metal salt particles is energy corresponding to the range of 5 to 60 minutes, preferably 5 to 30 minutes, as the duration of the aforementioned shaking with the paint shaker, and almost conforms to the aforementioned energy as applied during the pneumatic transportation. It is also possible to make design of the pneumatic transportation step, such as pneumatic transportation distance and pneumatic transportation speed, by changing the shaking duration of the PS in the above range to thus find out the optimum point for performances of the water-absorbent resin composition. In addition, it is also possible to determine the selection or amount of the water-soluble polyvalent metal salt particles being used, by carrying out a test of mixing the water-absorbent resin particles and the water-soluble polyvalent metal salt particles together for the shaking duration of the PS corresponding to energy of the already designed pneumatic transportation step. If such force is applied, a part of the water-soluble polyvalent metal salt particles as used in the present invention become extremely fine particles to come into a state where they adhere uniformly to water-absorbent resin particle surfaces.

That is to say, in a favorable process for production of the water-absorbent resin composition (1) according to the present invention, it is favorable that, for the water-absorbent resin particles, there are used the water-soluble polyvalent metal salt particles which are in itself so brittle as to be pulverized due to the above physical



energy; such as being in the form of: a powder of crystals; or an agglomerate or agglomerated material of fine particles; so that a mixture in which fine particles of the water-soluble polyvalent metal salt particles adhere uniformly to the water-absorbent resin particles (powder) can be obtained even if no special  
5 pulverization apparatus is used.

[Water-absorbent resin composition (2)]:

The water-absorbent resin composition (2) according to the present invention comprises water-absorbent resin particles and a metal compound, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including  
10 acrylic acid and/or its salt, and wherein:

the metal compound is one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts (except polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule); and

at least a part of the metal compound is fused to surfaces of the  
15 water-absorbent resin particles.

As to the metal compound as used in the present invention, taking it into consideration that the water-absorbent resin composition (2) according to the present invention is utilized as a water-absorbing agent for sanitary materials such as disposable diapers, then it is favorable to select a metal compound which does not  
20 color the water-absorbent resin composition (2) and which is low poisonous to human bodies.

The metal compound, as used in the present invention, is favorably one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts (except polyvalent metal salts of organic acids having not fewer than 7  
25 carbon atoms per molecule). Because the polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule have such high hydrophobicity as to deteriorate the capillary suction force of the water-absorbent resin composition, such polyvalent metal salts are not used as the polyvalent metal salts in the present

invention.

Favorable as the alkaline metal salts are salts of Li, Na, and K.

The polyvalent metal salts except the polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule are favorably those which  
5 contain one or not fewer than two polyvalent metals selected from among Be, Mg, Ca, Sc, Y, Ti, Zr, V, Nb, Cr, Mo, Mn, Fe, Ru, Co, Rh, Ni, Pd, Cu, Zn, Ga, Ge, and Al, and more favorably those which contain one or not fewer than two polyvalent metals selected from among Ca, Mg, Fe, Ti, Zr, Zn, and Al, and most favorably those which contain Al. In addition, the aforementioned water-soluble polyvalent metal salt  
10 particles may be used.

The metal compound, as used in the present invention, is favorably a metal compound having a melting point of not higher than 250 °C. In addition, the melting point of the metal compound is favorably in the range of 30 to 250 °C, more favorably 40 to 200 °C, still more favorably 50 to 150 °C, most favorably 60 to  
15 100 °C. In the case where the melting point is higher than 250 °C, there is a possibility that, when the metal compound is fused to the water-absorbent resin particles, damage may be done thereto, thus resulting in failure to obtain the objective properties. In addition, in the case where the melting point is lower than 30 °C, there is a possibility that the metal compound may unfavorably permeate the  
20 inside of the water-absorbent resin particles, thus resulting in failure to obtain the objective properties.

The metal compound, as used in the present invention, is favorably hydrophilic and/or water-soluble. Therefore, the polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule are unfavorable. In addition, in the  
25 case where the polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule are used, their hydrophobicity is too high, and therefore there is a possibility that they may cause such as deterioration of the surface tension of the water-absorbent resin particles, thus resulting in failure to obtain the objective

properties. In the present invention, the water solubility refers to a compound of which not less than 1 g, favorably not less than 10 g, dissolves into 100 g of water at 25 °C.

The metal compound, as used in the present invention, favorably has water of hydration in its molecule. The metal compound having the water of hydration is usually hydrophilic and easily takes effect when the water-absorbent resin composition (2) according to the present invention absorbs water.

The metal compound, as used in the present invention, is favorably a solid, more favorably a powder, at normal temperature. In the case where the metal compound is a powder, then, the finer its particle diameters are, the more easily it fuses with the water-absorbent resin particles. Therefore, its mass-average particle diameter (D50) is favorably not larger than 1,000  $\mu\text{m}$ , more favorably not larger than 600  $\mu\text{m}$ , still more favorably not larger than 300  $\mu\text{m}$ , most favorably not larger than 150  $\mu\text{m}$ .

Specific examples of the metal compound, as used in the present invention, include one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts having not more than 6 carbon atoms which are recorded as of September 10, 2003 in a GMELIN file provided by Gmelin Institute Varrentrappstr. and/or a BEILSTEIN file provided by Beilstein Chemiedaten und Software GmbH (each of them is usable as a search file of STN INTERNATIONAL, and their agency in Japan is the Scientific Information Society of Japan). As favorable ones among them, there can be cited those which have melting points of not higher than 250 °C. As more favorable ones among them, there can be cited one or not fewer than two members selected from among inorganic acid salts or salts of organic acids having not more than 6 carbon atoms per molecule, such as sulfates, nitrates, phosphates, halides, oxalates, and acetates. As still more favorable ones among them, there can be cited those which have water of hydration. Being exemplified more specifically, there are favorably used one or not fewer than two

members selected from among  $\text{Al}_2(\text{SO}_4)_3 \cdot 14-18\text{H}_2\text{O}$ ,  $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ ,  $(\text{NH}_4)\text{Al}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ ,  $\text{NaAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ ,  $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ ,  $\text{Na}_2\text{B}_4\text{O}_7 \cdot 5-10\text{H}_2\text{O}$ ,  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ , and  $\text{K}_2\text{SO}_4 \cdot \text{Fe}_2(\text{SO}_4)_3 \cdot 24\text{H}_2\text{O}$ . The most favorable is one or not fewer than two members selected from among  $\text{Al}_2(\text{SO}_4)_3 \cdot 14-18\text{H}_2\text{O}$  and  $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ .

5       The water-absorbent resin composition (2) according to the present invention comprises the water-absorbent resin particles and the metal compound, and is usually particulate and can be used favorably as an absorbent material for sanitary materials for absorption of urine, menstrual blood, sweat, and other body fluids.

10       The water-absorbent resin composition (2), according to the present invention, is a water-absorbent resin composition which comprises the water-absorbent resin particles and the metal compound, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt, and wherein: the metal compound is one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts having not more than 6 carbon  
15       atoms; and at least a part of the metal compound is fused to surfaces of the water-absorbent resin particles.

In the present invention, the term "fusion" refers to a state where at least a part of the metal compound is made to adhere strongly to surfaces of the water-absorbent resin particles in a melted state.

20       In the present invention, favorably, the water-absorbent resin particles are surface-crosslinked ones. In addition, more favorably, the water-absorbent resin particles are materials having been surface-crosslinked with a polyhydric alcohol. Even if a small amount of surface-treating agent remains on surfaces, the metal compound is not bound to surfaces of the water-absorbent resin particles merely by  
25       mixing the remaining-surface-treating-agent-containing water-absorbent resin particles and the metal compound together. Therefore, the remaining surface-treating agent is not usable as the binder.

The water-absorbent resin composition (2), according to the present invention,

is favorably free from binder as used to bind the water-absorbent resin particles and the metal compound to each other. The binder has a possibility of causing the deteriorations of the surface tension, the capillary absorption capacity (CSF), and other properties. For obtaining the water-absorbent resin composition (2) according to the present invention, it is not necessary to bind the water-absorbent resin particles and the metal compound to each other with the binder. The fusion of the metal compound to surfaces of the water-absorbent resin particles can strongly bind them together. As the case may be, the use of the binder rather hinders the fusion of the metal compound to the water-absorbent resin particles.

10 In the water-absorbent resin composition (2) according to the present invention, favorably, at least a part of the aforementioned metal compound is fused in the form of coating a part of surfaces of the water-absorbent resin particles in a layered state. The layered state refers to a state where the aforementioned metal compound thinly coats surfaces of the water-absorbent resin particles. The fusion in the form of the coating in the layered state prevents the metal compound from permeating the inside of the water-absorbent resin particles and also increases the surface area, and therefore makes it possible to easily obtain the objective performances. In addition, a form such that the surfaces of the water-absorbent resin particles are all coated with the metal compound has a possibility of bringing about deterioration of properties.

15 20 Therefore, it is favorable that the metal compound forms a discontinuous layer on surfaces of the water-absorbent resin particles. In addition, there is also preferred a form such that inter-particle binding between the aforementioned water-absorbent resin particles and the aforementioned metal compound is formed by the fusion. It may be possible to observe these states from photographs taken with such as electron

25 microscopes.

The water-absorbent resin composition (2), according to the present invention, is in the form of particles having a mass-average particle diameter in the range of favorably 100 to 600  $\mu\text{m}$ , more favorably 200 to 500  $\mu\text{m}$ . In the case where the

mass-average particle diameter is smaller than 100  $\mu\text{m}$ , then there is a possibility that: the handling property may be bad, and also much dust may be contained, and the liquid permeability and the liquid diffusibility may be bad. In the case where the mass-average particle diameter is larger than 600  $\mu\text{m}$ , then there is a possibility  
5 that damage may be easily done, thus resulting in deterioration of properties.

As to the water-absorbent resin composition (2) according to the present invention, the logarithmic standard deviation ( $\sigma\zeta$ ) of the particle diameter distribution is favorably in the range of 0.25 to 0.45, more favorably 0.27 to 0.47, still more favorably 0.30 to 0.40. The logarithmic standard deviation ( $\sigma\zeta$ ) of the  
10 particle diameter distribution is a numerical value indicating the broadness of the particle diameter distribution. The smaller this value is, the narrower particle diameter distribution it shows. That is to say, in the case where the logarithmic standard deviation ( $\sigma\zeta$ ) is more than 0.45, there is a possibility that the width of the particle diameter distribution may be too broad, thus resulting in bad handling  
15 property or in failure to obtain the objective properties. In the case where the logarithmic standard deviation ( $\sigma\zeta$ ) is less than 0.25, there is a possibility that the productivity may greatly be deteriorated, thus resulting in failure to obtain the effects corresponding to the cost.

As to the water-absorbent resin composition (2) according to the present  
20 invention, because at least a part of the metal compound is fused to surfaces of the water-absorbent resin particles, segregation of the metal compound occurs little. As an index for knowing the tendency for this segregation to occur, there is a metal compound segregation index. Specifically, the metal compound segregation index of the water-absorbent resin composition (2) according to the present invention is  
25 favorably in the range of 0.0 to 2.0, more favorably 0.0 to 1.7, most favorably 0.0 to 1.5. The metal compound segregation index is determined by the below-mentioned method.

As to the water-absorbent resin composition (2) according to the present

invention, when at least a part of the metal compound fuses to surfaces of the water-absorbent resin particles, fine particles also fuse as the case may be. Therefore, the dust prevention effect is high.

It has been found out by the present inventors that the water-absorbent resin composition (2) according to the present invention is, surprisingly, more excellent in the handling property during the moisture absorption when compared with other methods for adding the metal compound to the water-absorbent resin particles (e.g. dry mixing, aqueous solution addition, addition of a binder after dry mixing). In the case where the metal compound is fused to surfaces of the water-absorbent resin particles, the metal compound coats a part of surfaces of the water-absorbent resin particles, whereby the binding between the water-absorbent resin particles during the moisture absorption is inhibited, so that there is provided an effect of preventing the agglomeration between particles. This effect works more effectively by the fusion. For example, in the case where the metal compound is added in the form of an aqueous solution, the metal compound unfavorably permeates the inside of the particles, and it is therefore difficult to obtain the effect of preventing the agglomeration between particles. In addition, in the case of the dry mixing, the water-absorbent resin particles and the metal compound contact with each other at points, thus still resulting in failure to obtain so much effect as that in the case of the fusion.

As to the water-absorbent resin composition (2) according to the present invention, the blocking ratio (BR) when having been put at 25 °C and a relative humidity of 70 % for 1 hour is favorably not more than 20 %, more favorably not more than 10 %, still more favorably not more than 5 %.

The water-absorbent resin composition (2), according to the present invention, favorably displays an absorption capacity without load (CRC) in the range of 15 to 45 g/g, more favorably 20 to 37 g/g, still more favorably 24 to 32 g/g. In the case where the absorption capacity without load is low, there is a possibility that the

efficiency may be bad on an occasion of the use for sanitary materials such as diapers. In the case where the absorption capacity without load is too high, there is a possibility that the performance deterioration may occur due to such as deterioration of gel strength.

5           The water-absorbent resin composition (2), according to the present invention, favorably displays an absorption capacity under load (AAP) of not less than 16 g/g, more favorably not less than 20 g/g, still more favorably not less than 24 g/g. In addition, it is desirable that the deterioration of the absorption capacity under load (AAP) of the water-absorbent resin composition (2), as compared with an absorption  
10   capacity under load (AAP) of the water-absorbent resin particles to which the metal compound has not yet been added, is small. The water-absorbent resin composition (2) favorably maintains an absorption capacity under load (AAP) of not less than 0.85 time, more favorably not less than 0.90 time, most favorably not less than 0.95 time, in comparison with the absorption capacity under load (AAP) of the  
15   water-absorbent resin particles.

          The water-absorbent resin composition (2), according to the present invention, favorably displays a saline flow conductivity (SFC) of not less than  $30 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , more favorably not less than  $50 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , still more favorably not less than  $80 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ , most favorably not less than  $100 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$ .  
20   The saline flow conductivity (SFC) is a numerical value indicating the liquid permeability and the liquid diffusibility. The higher this value is, the more excellent in the liquid permeability and the liquid diffusibility the water-absorbent resin composition is. The saline flow conductivity (SFC) depends on the content of the water-absorbent resin composition in the sanitary material. The higher content  
25   needs the higher value of the saline flow conductivity (SFC).

          The water-absorbent resin composition (2), according to the present invention, favorably displays a capillary absorption capacity (CSF) of not less than 15 g/g, more favorably not less than 18 g/g, most favorably not less than 20 g/g, at a height of 20



cm. The capillary absorption capacity (CSF) is a value indicating the strength of the capillary suction force. The higher capillary absorption capacity (CSF) can more diffuse an absorbed liquid also in a height direction and is therefore more desirable.

5 As the water-absorbent resin composition (2) according to the present invention, there are used those of which the water-extractable component content is favorably not higher than 20 mass %, more favorably not higher than 15 mass %, most favorably not higher than 10 mass %. In the case where the water-extractable component content is higher than 20 mass % in the present invention, there is a  
10 possibility not only that no effects of the present invention may be obtained, but also that the performance may be deteriorated in the use for water-absorbent structures for sanitary materials such as diapers. In addition, such a water-extractable component content is unfavorable also from the viewpoint of safety. As a cause of the performance deterioration, it can be cited that, when the water-absorbent resin  
15 composition absorbs water to swell, a high-molecular component elutes from the inside of the water-absorbent resin to thereby hinder the liquid permeation. The eluted high-molecular component can be considered to resist when a liquid flows across surfaces of water-absorbent resin particles. In addition, similarly, the elution of the high-molecular component has a possibility of increasing the viscosity of an  
20 absorbed solution to thus deteriorate the capillary suction force. The water-extractable component content of the water-absorbent resin composition is measured by the below-mentioned method.

Although not especially limited, the water content of the water-absorbent resin composition (2) according to the present invention is favorably in the range of 0 to  
25 100 mass %, more favorably 0.01 to 40 mass %, still more favorably 0.1 to 10 mass %.

Although not especially limited, the bulk density of the water-absorbent resin composition (2) according to the present invention is favorably in the range of 0.40

to 0.90 g/ml, more favorably 0.50 to 0.80 g/ml (the method for measuring the bulk density is specified in JIS K-3362). In the case of water-absorbent resin compositions which have a bulk density of less than 0.40 g/ml or more than 0.90 g/ml, there is a possibility that they may be damaged easily by the process and may  
5 accordingly be deteriorated in property.

The water-absorbent resin composition (2) according to the present invention may possess such functions as given or enhanced by causing this composition to, besides the water-absorbent resin particles and the metal compound, further contain additives such as: water-insoluble finely-particulate inorganic powders (e.g. silicon  
10 dioxide, titanium dioxide, aluminum oxide, magnesium oxide, zinc oxide, talc, boric acid, silicic acid, clay, diatomite, zeolite, bentonite, kaolin, hydrotalcite, and salts (e.g. activated clay)); deodorants, perfumes, antibacterial agents, cationic polymer compounds (e.g. polyamines), foaming agents, pigments, dyes, manures, oxidizing agents, and reducing agents. The ratio of the additives as used is favorably less than  
15 10 mass %, more favorably less than 5 mass %, still more favorably less than 1 mass %, relative to the mass of the water-absorbent resin composition.

In the case where the water-absorbent resin composition (2) according to the present invention is used for the sanitary materials, if the metal compound is fused to the surface-crosslinked water-absorbent resin particles, then the wettability to  
20 aqueous liquids is good, and further, a liquid-absorbed gel little causes what is called gel-blocking, and spaces between gel particles are not clogged up due to close cohesion of the gel, either. Therefore, even in the case where the water-absorbent resin composition is used in a high concentration in absorbent structures such as diapers, it is possible that, at the second time or thereafter, urine or body fluids  
25 diffuse into the absorbent structures, without losing a place to go on surfaces of the absorbent structures, so that the urine or body fluids can be distributed to the inside water-absorbent resin particles. Furthermore, a mixture of the water-absorbent resin of the water-absorbent resin composition and an agglomerated material of this

water-absorbent resin composition has spaces of the appropriate size between particles and therefore combines a property of sucking a liquid by the capillary force and therefore can diffuse an absorbed liquid into the entire absorbent structure also by the capillary suction force.

5 [Process (2) for production of water-absorbent resin composition]:

The water-absorbent resin particles and the metal compound, which are used for the production of the water-absorbent resin composition (2) according to the present invention, are as previously explained.

10 The process for production of the water-absorbent resin composition (2) according to the present invention is a process comprising the steps of: heating the water-absorbent resin particles and/or the metal compound to a temperature of not lower than the melting point of the metal compound; and thereby fusing at least a part of the metal compound to surfaces of the water-absorbent resin particles; wherein the water-absorbent resin particles are obtained by polymerizing a monomer  
15 including acrylic acid and/or its salt.

In the present invention, the term "fusion" refers to a phenomenon such that at least a part of the metal compound melts by heat or becomes softer by heat than its solid state, thus adhering to another substance, and the term "fusion" is used in the same meaning as heat-fusion. However, there is also included a case where the  
20 metal compound is fused by melting it or making it softer than its solid state, by means other than heat.

That is to say, when it comes to the above process, there is no especial limitation. However, for example, there are the following processes (a) to (d):

(a) A process in which the water-absorbent resin particles, which have been  
25 heated to not lower than the melting point of the metal compound, and the metal compound are mixed together.

(b) A process in which the metal compound, which has been heated to not lower than the melting point of the metal compound, and the water-absorbent resin

particles are mixed together.

(c) A process in which the water-absorbent resin particles and the metal compound are mixed together, and then the resultant mixture is heated to not lower than the melting point of the metal compound.

- 5 (d) A process in which the water-absorbent resin particles, which have been heated to not lower than the melting point of the metal compound, and the metal compound, which has been heated to not lower than the melting point of the metal compound, are mixed together.

These processes are favorable, but other processes may be used.

- 10 In the present invention, examples of means of heating the water-absorbent resin particles and/or the metal compound include heaters, microwaves, ultrasonic waves, and far infrared rays.

In the present invention, favorably, the water-absorbent resin particles are materials having been surface-crosslink-treated with a compound having at least two  
15 functional groups which make a dehydration reaction or transesterification reaction with a carboxyl group.

In the present invention, the metal compound is used favorably in the form of not such as an aqueous dispersion or solution, but a powder. However, the metal compound may be used in a melted state without adding such as water thereto.

- 20 In the present invention, the metal compound is used favorably in an amount of 0.001 to 10 mass %, more favorably 0.01 to 5 mass %, most favorably 0.1 to 3 mass %, relative to the mass of the water-absorbent resin particles. In the case where the amount of the metal compound being added is smaller than 0.001 mass %, it is difficult to obtain the effects of the present invention. Also, in the case of the  
25 addition amount larger than 10 mass %, not only are there economical disadvantages, but also there is a possibility that the performance deterioration of the water-absorbent resin composition may be brought about.

In the present invention, there is no especial limitation on the adding and

mixing method. Publicly known methods for addition and mixing of powders may be used. In a favorable method, a predetermined amount of metal compound is added to the water-absorbent resin particles in a lump or divisionally or continuously.

In the present invention, favorably, the mixing of the water-absorbent resin particles and the metal compound is carried out under stirring. Also favorably under stirring, at least a part of the metal compound is fused to surfaces of the water-absorbent resin particles. Usable as stirring apparatuses are such as paddle blenders, ribbon mixers, rotary blenders, jar tumblers, plowshare mixers, and mortar mixers. These stirring apparatuses may be heatable apparatuses or may be apparatuses which cools the heated mixture.

In the present invention, when the water-absorbent resin particles and/or the metal compound is heated, the water-absorbent resin particles and/or the metal compound needs to be heated to not lower than the melting point of the metal compound. The temperature range to which the heating is carried out is favorably not higher than 250 °C, more favorably the range of 30 to 250 °C, most favorably 50 to 200 °C. In addition, a state where the metal compound does not entirely melt is favorable. In the case where the metal compound has entirely melted, there is a possibility that the water-absorbent resin particles may be coated entirely with the metal compound, thus resulting in deterioration of properties.

In the present invention, the stirring duration of the water-absorbent resin particles and the metal compound is not especially limited. However, it is favorably not more than 60 minutes, more favorably not more than 30 minutes.

In the present invention, when at least a part of the metal compound is fused to surfaces of the water-absorbent resin particles, it is favorable to give pressure to a mixture of the water-absorbent resin particles and the metal compound. This pressurization promotes the fusion.

In the present invention, after at least a part of the metal compound has been fused to surfaces of the water-absorbent resin particles, it is favorable to regulate the

particle diameter distribution of the resultant water-absorbent resin composition. As means for regulating the particle diameter distribution, it is enough to use publicly known means. However, a disintegrator and/or a classifier is favorably used.

5           In the present invention, when the water-absorbent resin particles and the metal compound are mixed together, addition of water and mixing of the metal compound in an aqueous solution state must be avoided. In the case where the addition of water or the mixing of the metal compound in an aqueous solution state is made, the metal compound component unfavorably permeates the inside of the water-absorbent  
10 resin particles, thus resulting in failure to sufficiently display the effects of the present invention. For example, in JP-A-523289/2001 (Kohyo) (WO 98/48857), there is disclosed a process for preparation of a super-water-absorbent polymer which process is characterized by including the steps of mixing a super-water-absorbent polymer with a polyvalent metal salt and then bringing the  
15 resultant mixture into close contact with a binder. Water or a water-soluble liquid is stated therein as the aforementioned binder. On the other hand, the present invention entirely differs from conventional processes in that, when the water-absorbent resin composition (2) according to the present invention is produced, an aqueous liquid (e.g. water, water-soluble liquid) does not need to be added after  
20 the mixing of the water-absorbent resin particles and the metal compound. Therefore, in the present invention, methods without using any binder are favorable. The use of the binder has a possibility of deteriorating the surface tension, the CSF, and other properties. In addition, as an example of methods without using the water or the water-soluble liquid, there can be considered a method in which the  
25 water-absorbent resin particles and the metal compound are mixed together in a dry manner. However, this method has a possibility of causing such as segregation or increase of dust. In addition, there are problems such that: even if the metal compound adheres to the water-absorbent resin particles, its binding force is so weak

that the metal compound is unfavorably easily released merely by application of a little force. In the present invention, it is an important method to fuse at least a part of the metal compound to surfaces of the water-absorbent resin particles. By this method, it is possible to provide the water-absorbent resin composition which solves  
5 the aforementioned prior problems, and involves little segregation of the additive, and is excellent in the liquid permeability and the liquid diffusibility, and also, undergoes little deterioration of the absorption capacity under load, and further, surprisingly, is excellent also in the handling property during the moisture absorption, and further has a dust prevention effect.

10 [Water-absorbent structure]:

The water-absorbent resin composition (1) and/or (2), according to the present invention, can be combined with an appropriate material and thereby formed into the water-absorbent structure which is, for example, favorable as an absorbent layer for sanitary materials. Hereinafter, a description is made about the water-absorbent  
15 structure in the present invention.

The water-absorbent structure in the present invention refers to a molded composition which comprises a water-absorbent resin composition and another material and is used for sanitary materials (e.g. disposable diapers, sanitary napkins, incontinent pads, and medical pads) for absorption of such as blood, body fluids, and  
20 urine. Examples of the above other material include cellulose fibers. Specific examples of the cellulose fibers include: wood pulp fibers from wood, such as mechanical pulp, chemical pulp, semichemical pulp, and dissolving pulp; and synthetic cellulose fibers, such as rayon and acetate. Favorable cellulose fibers are the wood pulp fibers. These cellulose fibers may partially contain synthetic fibers  
25 such as nylon and polyester. When the water-absorbent resin composition (1) and/or (2) according to the present invention is used as a portion of the water-absorbent structure, the mass of the water-absorbent resin composition (1) and/or (2) according to the present invention as contained in the water-absorbent

structure is favorably in the range of not smaller than 20 mass %. In the case where the mass of the water-absorbent resin composition (1) and/or (2) according to the present invention as contained in the water-absorbent structure is smaller than 20 mass %, there is a possibility that no sufficient effects can be obtained.

5        For the purpose of obtaining the water-absorbent structure from the water-absorbent resin composition (1) and/or (2) according to the present invention and the cellulose fibers, for example, publicly known means for obtaining water-absorbent structures can appropriately be selected from among such as: a method in which the water-absorbent resin composition (1) and/or (2) is spread onto paper or  
10    mat made of such as the cellulose fibers and is, if necessary, interposed therebetween; and a method in which the cellulose fibers and the water-absorbent resin composition (1) and/or (2) are uniformly blended together. A favorable method is a method in which the water-absorbent resin composition (1) and/or (2) and the cellulose fibers are mixed together in a dry manner and then compressed.  
15    This method can greatly prevent the water-absorbent resin composition (1) and/or (2) from falling off from the cellulose fibers. The compression is favorably carried out under heating, and its temperature range is, for example, the range of 50 to 200 °C. In addition, for the purpose of obtaining the water-absorbent structure, methods as disclosed in JP-A-509591/1997 (Kohyo) and JP-A-290000/1997 (Kokai) are also  
20    favorably used.

      In the case where used for water-absorbent structures, the water-absorbent resin composition (1) and/or (2) according to the present invention is so excellent in the properties as to give water-absorbent structures which are very excellent in that they quickly take liquids in and further in that the amount of the liquids remaining on  
25    their surface layers is small.

      Because the water-absorbent resin compositions (1) and (2) according to the present invention have excellent water absorption properties, these water-absorbent resin compositions can be used as water-absorbing and water-retaining agents for



various purposes. For example, these water-absorbent resin compositions can be used for such as: water-absorbing and water-retaining agents for absorbent articles (e.g. disposable diapers, sanitary napkins, incontinent pads, and medical pads); agricultural and horticultural water-retaining agents (e.g. substitutes for peat moss, soil-modifying-and-improving agents, water-retaining agents, and agents for duration of effects of agricultural chemicals); water-retaining agents for buildings (e.g. dew-condensation-preventing agents for interior wall materials, cement additives); release control agents; coldness-retaining agents; disposable portable body warmers; sludge-solidifying agents; freshness-retaining agents for foods; ion-exchange column materials; dehydrating agents for sludge or oil; desiccating agents; and humidity-adjusting materials. In addition, the water-absorbent resin compositions (1) and (2) according to the present invention can be used particularly favorably for sanitary materials for absorption of excrement, urine, or blood, such as disposable diapers and sanitary napkins.

15 In the case where the water-absorbent structure in the present invention is used for sanitary materials (e.g. disposable diapers, sanitary napkins, incontinent pads, and medical pads), this water-absorbent structure is used favorably with a constitution including: (a) a liquid-permeable top sheet placed so as to be adjacent to a wearer's body; (b) a liquid-impermeable back sheet placed so as to be adjacent to the wearer's clothes at a distance from the wearer's body; and (c) the water-absorbent structure placed between the top sheet and the back sheet. The water-absorbent structure may be in more than one layer or used along with such as a pulp layer.

#### DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

Hereinafter, the present invention is more specifically illustrated by the following Examples of some preferred embodiments in comparison with Comparative Examples not according to the present invention. However, the present invention is not limited to them. Hereinafter, for convenience, the units "mass part(s)" and "liter(s)" may be referred to simply as "part(s)" and "L"

respectively. In addition, the unit "mass %" may be referred to as "wt %".

The measurement and evaluation methods in the Examples and the Comparative Examples are shown below.

Unless otherwise noted, the following measurement is stated as having been  
5 carried out under conditions of a room temperature (25 °C) and a humidity of 50 RH%.

Incidentally, in cases of water-absorbent resin compositions having been used for end products such as sanitary materials, the water-absorbent resin compositions have already absorbed moisture. Therefore, the measurement may be carried out  
10 after appropriately separating the water-absorbent resin compositions from the end products and then drying the separated water-absorbent resin compositions under a reduced pressure at a low temperature (e.g. under not higher than 1 mmHg at 60 °C for 12 hours). In addition, all the water-absorbent resin compositions as used in the Examples and Comparative Examples of the present invention had water contents of  
15 not higher than 6 mass %.

<Absorption capacity without load (CRC)>:

An amount of 0.20 g of water-absorbent resin or water-absorbent resin composition was weighed out precisely to a level of 0.0001 g, and then uniformly placed and sealed into a bag (85 mm × 60 mm or 60 mm × 60 mm) made of  
20 nonwoven fabric (trade name: Heatron Paper, type: GSP-22, produced by Nangoku Pulp Kogyo Co., Ltd.).

A container of 1 L was charged with 1 L of 0.9 mass % aqueous sodium chloride solution (physiological saline solution), in which one evaluation sample per one container was then immersed for 1 hour. Incidentally, because the present  
25 invention is an invention made by directing attention to effects of ion migration, more than one sample per one container must not be immersed.

After 1 hour, the bag was pulled up and then drained of water by centrifugal force of 250 G with a centrifugal separator (produced by Kokusan Co., Ltd.,

centrifugal separator: model H-122) for 3 minutes, and then the mass W1 (g) of the bag was measured. In addition, the same procedure as the above was carried out without the water-absorbent resin or water-absorbent resin composition, and the resultant mass W0 (g) was measured. Then, the absorption capacity (g/g) without  
5 load was calculated from these W1 and W0 in accordance with the following equation:

$$\text{CRC (g/g)} = [(W1 \text{ (g)} - W0 \text{ (g)}) / \text{mass (g) of water-absorbent resin or water-absorbent resin composition}] - 1$$

<Absorption capacity under load (AAP)>:

10 The AAP was measured by the following method A or B. The AAP may be measured by either of these methods, and the measured value is little influenced by the measurement method.

As to the AAP in the below-mentioned Referential Examples, Examples, and Comparative Examples, the AAP as shown in Table 1 is AAP as measured by the  
15 method A, and the other AAP is AAP as measured by the method B.

(Method A):

The absorption capacity under a load (AAP) was measured with an apparatus of Fig. 1.

There was prepared a load 21 as adjusted so as to give a pressure of 4.83 kPa  
20 (0.7 psi). Onto a metal gauze 18 of 400 meshes (mesh opening size of 38  $\mu\text{m}$ ) of a plastic cylinder 19 of 60 mm in diameter with the metal gauze stuck on its bottom, there was dispersed about 0.90 g (Wp2) of water-absorbent resin composition or water-absorbent resin, and further thereon the above load 21 (in the case of 0.7 psi) was put to prepare a liquid absorption instrument, which was then put on a filter  
25 paper 17 on a glass filter 13 of Fig. 1. After 60 minutes, there was measured a value (Wc) of the physiological saline solution (0.90 mass % aqueous NaCl solution) as absorbed by the water-absorbent resin composition or water-absorbent resin. The absorption capacity under the load was determined using the following equation.

$$AAP \text{ (g/g)} = Wc/Wp2$$

(Method B):

The measurement was carried out with an apparatus as shown in Fig. 2.

A stainless metal gauze 101, which was a screen of 400 meshes (mesh opening size: 38  $\mu\text{m}$ ), was attached by fusion to a bottom of a plastic supporting cylinder 100 having an inner diameter of 60 mm. Then, under conditions of a room temperature (20 to 25  $^{\circ}\text{C}$ ) and a humidity of 50 RH%, onto the above metal gauze, there was uniformly spread 0.90 g of water-absorbent resin or water-absorbent resin composition 102, and further thereon, there were mounted a piston 103 and a load 104 in sequence, wherein the piston had an outer diameter of only a little smaller than 60 mm and made no gap with the inner wall surface of the supporting cylinder, but was not hindered from moving up and down, and wherein the piston and the load were adjusted so that a load of 4.83 kPa (0.7 psi) could uniformly be applied to the water-absorbent resin or water-absorbent resin composition. Then, the mass  $W_a$  (g) of the resultant one set of measurement apparatus was measured.

A glass filter plate 106 having a diameter of 90 mm (produced by Sogo Rikagaku Glass Seisakusho Co., Ltd., pore diameter: 100 to 120  $\mu\text{m}$ ) was mounted inside a Petri dish 105 having a diameter of 150 mm, and then a 0.90 mass % aqueous sodium chloride solution (physiological saline solution) 108 (20 to 25  $^{\circ}\text{C}$ ) was added up to the same level as the top surface of the glass filter plate, on which a filter paper 107 having a diameter of 90 mm (produced by ADVANTEC Toyo Co., Ltd., trade name: (JIS P 3801, No.2), thickness: 0.26 mm, diameter of captured particles: 5  $\mu\text{m}$ ) was then mounted so that its entire surface would be wetted, and further, an excess of liquid was removed.

The one set of measurement apparatus was mounted on the above wet filter paper, thereby getting the liquid absorbed under the load for a predetermined duration. This absorption duration was defined as 1 hour from the start of the measurement. Specifically, 1 hour later, the one set of measurement apparatus was

removed by being lifted to measure its mass  $W_b$  (g). This measurement of the mass must be carried out as quickly as possible and so as not to give any vibration. Then, the absorption capacity under load (AAP) (g/g) was calculated from the  $W_a$  and  $W_b$  in accordance with the following equation:

5        
$$\text{AAP (g/g)} = [W_b \text{ (g)} - W_a \text{ (g)}] / \text{mass (g) of water-absorbent resin or water-absorbent resin composition}$$

<Saline flow conductivity (SFC)>:

The following test was carried out according to the saline flow conductivity (SFC) test as described in JP-A-509591/1997 (Kohyo).

10        An apparatus as shown in Fig. 3 was used, and a water-absorbent resin or water-absorbent resin composition (0.900 g) as uniformly placed in a receptacle 40 was swollen in synthetic urine (1) under a load of 0.3 psi (2.07 kPa) for 60 minutes (which was 120 minutes in the case of measuring the retention ratio of the saline flow conductivity (SFC)), and the gel layer height of the resultant gel 44 was  
15        recorded. Next, under the load of 0.3 psi (2.07 kPa), a 0.69 mass % aqueous sodium chloride solution 33 was passed through the swollen gel layer from a tank 31 under a constant hydrostatic pressure. This SFC test was carried out at room temperature (20 to 25 °C). The amount of the liquid passing through the gel layer was recorded as a function to time with a computer and a balance at twenty seconds'  
20        intervals for 10 minutes. The rate  $F_s(t)$  of the flow passing through the swollen gel 44 (mainly between particles thereof) was determined in a unit of g/s by dividing the incremental mass (g) by the incremental time (s). The time when the constant hydrostatic pressure and the stable flow rate are obtained was represented by  $t_s$ , and only the data as obtained between  $t_s$  and 10 minutes were used for the flow rate  
25        calculation. The  $F_s(t=0)$  value, namely, the initial rate of the flow passing through the gel layer, was calculated from the flow rates as obtained between  $t_s$  and 10 minutes. The  $F_s(t=0)$  was calculated by extrapolating the results of a least-squares fit of  $F_s(t)$  versus time to  $t=0$ .

SFC

$$= (F_s(t=0) \times L_0) / (\rho \times A \times \Delta P)$$

$$= (F_s(t=0) \times L_0) / 139,506$$

where:

5  $F_s(t=0)$ : flow rate denoted by g/s;

$L_0$ : initial thickness of gel layer denoted by cm;

$\rho$ : density of NaCl solution (1.003 g/cm<sup>3</sup>);

$A$ : area of top of gel layer in cell 41 (28.27 cm<sup>2</sup>);

$\Delta P$ : hydrostatic pressure applied to gel layer (4,920 dyne/cm<sup>2</sup>); and

10 the unit of the SFC is: ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ ).

As to the apparatus as shown in Fig. 3, a glass tube 32 was inserted in the tank 31, and the lower end of the glass tube 32 was placed so that the 0.69 mass % aqueous sodium chloride solution 33 could be maintained at a height of 5 cm from the bottom of the swollen gel 44 in the cell 41. The 0.69 mass % aqueous sodium chloride solution 33 in the tank 31 was supplied to the cell 41 through an L-tube 34 having a cock. A receptacle 48 to collect the passed liquid was placed under the cell 41, and this collecting receptacle 48 was set on a balance 49. The inner diameter of the cell 41 was 6 cm, and a No. 400 stainless metal gauze (mesh opening size: 38  $\mu\text{m}$ ) 42 was set at the bottom thereof. Holes 47 sufficient for the liquid to pass through were opened in the lower portion of a piston 46, and its bottom portion was equipped with a well-permeable glass filter 45 so that the water-absorbent resin or water-absorbent resin composition or their swollen gels would not enter the holes 47. The cell 41 was placed on a stand to put the cell thereon. The face, contacting with the cell, of the stand was set on a stainless metal gauze 43 that did not inhibit the liquid permeation.

25 The synthetic urine (1) as used was obtained by mixing together the following: 0.25 g of calcium chloride dihydrate; 2.0 g of potassium chloride; 0.50 g of magnesium chloride hexahydrate; 2.0 g of sodium sulfate; 0.85 g of ammonium

dihydrogenphosphate; 0.15 g of diammonium hydrogenphosphate; and 994.25 g of pure water.

<Saline flow conductivity (SFC) after paint shaker test (SFC after PS)>:

The following measurement was carried out on the basis of the apparatus as disclosed in JP-A-235378/1997 (Kokai).

A glass bottle of 6 cm in diameter and 11 cm in height having a lid was charged with 30 g of water-absorbent resin or water-absorbent resin composition and about 10 g of glass beads (diameter: 6 mm), and then attached to a TOYOSEIKI PAINT SHAKER (for 100 V, 60 HZ), and then shaken for 30 minutes. The glass beads and the water-absorbent resin or water-absorbent resin composition were sieved with a metal gauze having a mesh opening size of about 2 mm, thus obtaining a water-absorbent resin or water-absorbent resin composition after the paint shaker test.

The saline flow conductivity of the obtained water-absorbent resin or water-absorbent resin composition was measured by the aforementioned method.

Incidentally, the CRC and AAP after PS can also be measured by the same method.

<Mass-average particle diameter (D50) and logarithmic standard deviation ( $\sigma\zeta$ ) of particle diameter distribution>:

Water-absorbent resins or water-soluble polyvalent metal salt particles or water-absorbent resin compositions were classified with JIS standard sieves having mesh opening sizes of such as 850  $\mu\text{m}$ , 710  $\mu\text{m}$ , 600  $\mu\text{m}$ , 500  $\mu\text{m}$ , 425  $\mu\text{m}$ , 300  $\mu\text{m}$ , 212  $\mu\text{m}$ , 150  $\mu\text{m}$ , and 45  $\mu\text{m}$ . Then, the percentages R of the residues on these sieves were plotted on a logarithmic probability paper. Therefrom, a particle diameter corresponding to R = 50 mass % was read as the mass-average particle diameter (D50). In addition, the logarithmic standard deviation ( $\sigma\zeta$ ) of the particle diameter distribution is shown by the following equation. The smaller  $\sigma\zeta$  value shows the narrower particle diameter distribution.

$$\sigma\zeta = 0.5 \times \ln(X2/X1)$$

(wherein: X1 is a particle diameter when R = 84.1 mass %, and X2 is a particle diameter when R = 15.9 mass %)

As to the classification method for measuring the mass-average particle diameter (D50) and the logarithmic standard deviation ( $\sigma\zeta$ ) of the particle diameter distribution, 10.0 g of water-absorbent resin or water-soluble polyvalent metal salt particles or water-absorbent resin composition was placed onto JIS standard sieves (having mesh opening sizes of 850  $\mu\text{m}$ , 710  $\mu\text{m}$ , 600  $\mu\text{m}$ , 500  $\mu\text{m}$ , 425  $\mu\text{m}$ , 300  $\mu\text{m}$ , 212  $\mu\text{m}$ , 150  $\mu\text{m}$ , and 45  $\mu\text{m}$ ) (THE IIDA TESTING SIEVE: diameter = 8 cm) under conditions of a room temperature (20 to 25 °C) and a humidity of 50 RH%, and then classified with a shaking classifier (IIDA SIEVE SHAKER, TYPE: ES-65 type, SER. No. 0501) for 5 minutes.

<Dust generation degree>:

The water-absorbent resin or water-absorbent resin composition was placed into a PE bag (No. 13) as beforehand coated with an antistatic agent. This bag was shaken 30 times and then opened to carry out the measurement with DIGITAL DUST INDICATOR P-5L (produced by SHIBATA) for 1 minute. This measurement was carried out 10 times, and its average value was determined.

<Capillary absorption capacity (CSF)>:

The CSF is an index showing the capillary suction force of the water-absorbent resin or water-absorbent resin composition.

The capillary absorption capacity is determined by measuring the ability of the absorbent structure to absorb a liquid against a negative pressure gradient of the water column of 20 cm under a load of 0.06 psi within a predetermined time.

While referring to Fig. 4, an apparatus and method for measuring the capillary absorption capacity are described.

A conduit 3 was connected to a lower portion of a glass filter 2 of 60 mm in diameter having a liquid-absorbing surface of a porous glass plate 1 (glass filter



particle No. #3: Buchner type filter TOP 17G-3 (code no. 1175-03) produced by Sogo Rikagaku Glass Seisakusho Co., Ltd.), and this conduit 3 was connected to an opening as provided to a lower portion of a liquid storage container 4 of 10 cm in diameter. The porous glass plate of the aforementioned glass filter has an average pore diameter of 20 to 30  $\mu\text{m}$ , and can retain water in the porous glass plate by its capillary force against the negative pressure of the water column even in a state where a difference of 60 cm between heights of liquid surfaces is made, so that a state of no introduction of air can be kept. A supporting ring 5 was fitted to the glass filter 2 in order to let up and down its height, and the system was filled with a 0.90 mass % physiological saline solution (0.90 mass % aqueous NaCl solution) 6, and the liquid storage container was put on a balance 7. After it had been confirmed that there was no air in the conduit and under the porous glass plate of the glass filter, the difference in height between a liquid surface level of the top of the physiological saline solution (0.90 mass % aqueous NaCl solution) 6 in the liquid storage container 4 and a level of the upside of the porous glass plate 1 was adjusted to 20 cm, and then the glass filter was fixed to a stand 8.

An amount of 0.44 g of specimen to be measured 9 (water-absorbent resin composition) was quickly dispersed uniformly onto the glass filter (porous glass plate 1) in the funnel, and further thereon a load 10 (0.06 psi) having a diameter of 59 mm was put, and then, 30 minutes later, there was measured a value (W20) of the 0.90 mass % physiological saline solution (0.90 mass % aqueous NaCl solution) as absorbed by the specimen to be measured 9.

The capillary absorption capacity is determined from the following equation.

$$\text{Capillary absorption capacity (CSF) (g/g)} = \text{absorption amount (W20) (g)} / 0.44 \text{ (g)}$$

<Retention ratio of saline flow conductivity (SFC) (retention ratio of SFC)>:

In the aforementioned method for measurement of the saline flow conductivity (SFC), the swelling time under the load of 0.3 psi (2.07 kPa) was changed from 60

minutes to 120 minutes, and thereafter the measurement was carried out in the same way. The saline flow conductivity (SFC) as measured after the swelling time of 60 minutes is herein referred to as SFC (1 hr), and the saline flow conductivity (SFC) as measured after the swelling time of 120 minutes is herein referred to as SFC (2 hr).

- 5 The retention ratio of the SFC is represented by the following equation:

$$\text{Retention ratio (\% of SFC)} = [\text{SFC (2 hr)}/\text{SFC (1 hr)}] \times 100$$

<Retention ratio of saline flow conductivity (SFC) after paint shaker test (retention ratio after PS)>:

- 10 The following measurement was carried out on the basis of the apparatus as disclosed in JP-A-235378/1997 (Kokai).

A glass bottle of 6 cm in diameter and about 11 cm in height having a lid was charged with 30 g of water-absorbent resin composition and about 10 g of glass beads (diameter: 6 mm), and then attached to a TOYOSEIKI PAINT SHAKER (for 100 V, 60 HZ), and then shaken for 30 minutes. The glass beads and the  
15 water-absorbent resin composition were sieved with a metal gauze having a mesh opening size of about 2 mm, thus obtaining a water-absorbent resin composition after the paint shaker test. The liquid permeation rate under load of the obtained water-absorbent resin composition was measured by the aforementioned method. When the liquid permeation rate under load of the water-absorbent resin composition  
20 after the paint shaker test is represented by SFC (after PS) and when the liquid permeation rate under load of the water-absorbent resin composition before the paint shaker test is represented by SFC (before PS), then the retention ratio after PS is represented by the following equation:

$$\text{Retention ratio (\% after PS)} = [\text{SFC (after PS)}/\text{SFC (before PS)}] \times 100$$

- 25 <Contact angle>:

A double-coated pressure-sensitive adhesive tape was stuck onto an SUS sheet, and then the water-absorbent resin or water-absorbent resin composition was closely and uniformly spread onto this double-coated tape, and then the water-absorbent

resin or water-absorbent resin composition which had not adhered to the double-coated tape was scraped off to prepare a specimen sheet of which the surface was covered with the water-absorbent resin or water-absorbent resin composition. When a physiological saline solution (0.90 mass %) was made to contact with the  
5 above specimen sheet, the contact angle was measured by the sessile drop method with a contact angle meter (FACE CA-X model, produced by Kyowa Kaimen Kagaku K.K.) under conditions of 20 °C and 60 % RH. The contact angle at 1 second later than dropping a liquid drop of the physiological saline solution onto the specimen sheet was measured 5 times per one specimen. Its average value was  
10 determined and taken as the contact angle of the water-absorbent resin or water-absorbent resin composition.

<Extractable (water-extractable) component content>:

Into a plastic receptacle of 250 ml in capacity having a lid, 184.3 g of 0.90 mass % physiological saline solution was weighed out. Then, 1.00 g of  
15 water-absorbent resin particles or water-absorbent resin composition was added to this aqueous solution, and they were stirred for 16 hours, thereby the extractable component content in the resin was extracted. This extract liquid was filtrated with a filter paper (produced by ADVANTEC Toyo Co., Ltd., trade name: (JIS P 3801, No.2), thickness: 0.26 mm, diameter of captured particles: 5 µm), and then 50.0 g of  
20 the resultant filtrate was weighed out and used as a measuring solution.

To begin with, only the 0.90 mass % physiological saline solution was firstly titrated with an aqueous 0.1N NaOH solution until the pH reached 10, and then the resultant solution was titrated with an aqueous 0.1N HCl solution until the pH reached 2.7, thus obtaining blank titration amounts ([bNaOH] ml and [bHCl] ml).

25 The same titration procedure was carried out for the measuring solution, thus obtaining titration amounts ([NaOH] ml and [HCl] ml).

For example, if the water-absorbent resin or water-absorbent resin particles or water-absorbent resin composition comprises acrylic acid and its sodium salt in

known amounts, the extractable component content of the water-absorbent resin can be calculated from the average molecular weight of the monomers and the titration amounts, as obtained from the above procedures, in accordance with the following equation. In the case of unknown amounts, the average molecular weight of the monomers is calculated from the neutralization degree as determined by the titration.

$$\text{Extractable component content (mass \%)} = 0.1 \times (\text{average molecular weight}) \times 184.3 \times 100 \times ([\text{HCl}] - [\text{bHCl}]) / 1,000 / 1.0 / 50.0$$

$$\text{Neutralization degree (mol \%)} = (1 - ([\text{NaOH}] - [\text{bNaOH}]) / ([\text{HCl}] - [\text{bHCl}])) \times 100$$

10 <Quantification of metal compound>:

Appropriate publicly known methods are used for the quantification of the metal compound. For example, the quantification is carried out by such as atomic absorption photometry, plasma emission spectrometry, and absorption photometry which uses color reaction reagents. When it comes to the water-soluble metal compounds, the plasma emission spectrometry is favorably used. Hereinafter an example of the measurement methods is cited.

An amount of 1.0 g of water-absorbent resin composition was weighed out into a polypropylene-made beaker of 260 ml in capacity, and then thereto 190.0 g of 0.90 mass % physiological saline solution and 10.0 g of 2N hydrochloric acid were added, and then they were stirred at room temperature for 30 minutes. After the stirring, the resultant supernatant was filtered with a chromatodisk (GL Chromatodisk 25A of GL Science). The filtrate was analyzed by plasma emission spectrometry (with ULTIMA, produced by Horiba Seisakusho) to determine the metal salt concentration. Incidentally, the calibration curve was prepared from a 0.90 mass % physiological saline solution containing a known amount of polyvalent metal. Based on the determined polyvalent metal concentration, the polyvalent metal concentration in the water-absorbent resin composition is shown by the following equation:

Polyvalent metal concentration (mass %) in water-absorbent resin composition  
= (polyvalent metal concentration (mass %) in solution)  $\times$  200

<Metal compound segregation index>:

This is a numerical value indicating the tendency for the segregation of the  
5 metal compound to occur and is measured by the following method.

The metal compound as contained in the water-absorbent resin composition is  
quantified by the aforementioned method and calculated from the following  
equation:

$$\text{Metal compound segregation index} = W71 \text{ (ppm)} / W72 \text{ (ppm)}$$

10 The W71 represents a concentration (ppm) of a metal component which is, for  
example, contained in a water-absorbent resin composition having passed through a  
JIS standard sieve of 300  $\mu\text{m}$  in mesh opening size (a portion having particle  
diameters of not larger than 300  $\mu\text{m}$ ) by carrying out the classification operation in  
the same way as of the aforementioned measurement of the mass-average particle  
15 diameter. In addition, the W72 represents a concentration (ppm) of the metal  
component which is contained in a water-absorbent resin composition before the  
classification operation. The mesh opening size of the sieve in the present  
measurement method is changed appropriately for the mass-average particle diameter  
of the water-absorbent resin composition. Concretely, it is specified as follows:  
20 when the mass-average particle diameter is not larger than 150  $\mu\text{m}$ , a JIS standard  
sieve of 45  $\mu\text{m}$  in mesh opening size should be used; when the mass-average particle  
diameter is in the range of 150 to 300  $\mu\text{m}$ , a JIS standard sieve of 150  $\mu\text{m}$  in mesh  
opening size should be used; when the mass-average particle diameter is in the range  
of 300 to 500  $\mu\text{m}$ , a JIS standard sieve of 300  $\mu\text{m}$  in mesh opening size should be  
25 used; and, when the mass-average particle diameter is not smaller than 500  $\mu\text{m}$ , a JIS  
standard sieve of 500  $\mu\text{m}$  in mesh opening size should be used.

<Blocking ratio (BR)>:

This refers to a blocking ratio when having been put at 25  $^{\circ}\text{C}$  and a relative

humidity of 70 % for 1 hour.

An amount of 2 g of water-absorbent resin particles (or water-absorbent resin composition) is uniformly spread onto a bottom of a polypropylene-made cup of 50 mm in inner diameter of the bottom and 10 mm in height and then quickly placed  
5 into a thermohumidistatic incubator (PLATIOOUS LUCIFER PL-2G, produced by Tabai Espec Co., Ltd.) (which had beforehand been adjusted to 25 °C and the relative humidity of 70 %) and then left alone for 60 minutes. Thereafter, the water-absorbent resin particles having absorbed the moisture were transferred onto a JIS standard sieve of 7.5 cm in diameter and 2,000 µm in mesh opening size and then  
10 sieved with a shaking classifier (JDA SIEVE SHAKER, TYPE: ES-65 type, SER. No. 0501) for 5 minutes. Then, mass W4 (g) of water-absorbent resin particles remaining on the sieve and mass W5 (g) of water-absorbent resin particles having passed through the sieve were measured.

The moisture absorption blocking ratio (%) was calculated from the following  
15 equation:

$$\text{Moisture absorption blocking ratio (\%)} = \frac{\text{mass W4 (g)}}{\text{mass W4 (g)} + \text{mass W5 (g)}} \times 100$$

The lower the moisture absorption blocking ratio is, the more excellent the moisture absorption flowability is.

20 As to the below-mentioned experiments for demonstrating the effects of adding the metal compound, it is favorable that the same precursor is used to compare the above effects. For example, if the particle diameter distribution of the precursor varies, there is a possibility that the parameters depending on the particle diameter distribution, such as SFC, cannot precisely be evaluated. For example, when the  
25 performance, as indicated by the SFC, of water-absorbent resin particles is compared, it is favorable to compare the SFC using water-absorbent resin particles having almost the same CRC and particle diameter distributions.

[Referential Example 1]:

In a reactor as prepared by lidding a jacketed stainless twin-arm kneader of 10 liters in capacity having two sigma-type blades, there was prepared a reaction liquid by dissolving 11.7 g (0.10 mol %) of polyethylene glycol diacrylate into 5,438 g of aqueous solution of sodium acrylate having a neutralization degree of 71.3 mol %  
5 (monomer concentration: 39 mass %). Next, this reaction liquid was deaerated under an atmosphere of nitrogen gas for 30 minutes. Subsequently, 29.34 g of 10 mass % aqueous sodium persulfate solution and 24.45 g of 0.1 mass % aqueous L-ascorbic acid solution were added thereto under stirred conditions. As a result, polymerization started after about 1 minute. Then, the polymerization was carried  
10 out in the range of 20 to 95 °C while the forming gel was pulverized. Then, the resultant crosslinked hydrogel polymer was taken out after 30 minutes from the start of the polymerization. The crosslinked hydrogel polymer as obtained above was in the form of finely divided pieces having diameters of not larger than about 5 mm. This finely divided crosslinked hydrogel polymer was spread onto a metal gauze of  
15 50 meshes (mesh opening size: 300 µm) and then hot-air-dried at 180 °C for 50 minutes, thus obtaining a water-absorbent resin (1) which was of the irregular shape and easy to pulverize, such as in the form of particles, a powder, or a particulate dried material agglomerate.

The resultant water-absorbent resin (1) was pulverized with a roll mill and then  
20 further classified with a JIS standard sieve having a mesh opening size of 850 µm. Next, particles having passed through the 850 µm in the aforementioned operation were classified with a JIS standard sieve having a mesh opening size of 150 µm, whereby a water-absorbent resin (1aF) passing through the JIS standard sieve having the mesh opening size of 150 µm was removed, thus obtaining a particulate  
25 water-absorbent resin (1a).

The resultant water-absorbent resin (1aF) was agglomerated according to the method of Granulation Example 1 as disclosed in USP 6,228,930. The resultant agglomerated material was pulverized and classified by the same procedure as the

aforementioned, thus obtaining an agglomerated water-absorbent resin (1aA).

An amount of 90 mass parts of the water-absorbent resin (1a) and 10 mass parts of the water-absorbent resin (1aA), as obtained in the above way, were uniformly mixed together to obtain a water-absorbent resin (A).

5 In addition, similarly, the resultant water-absorbent resin (1) was pulverized with a roll mill and then further classified with a JIS standard sieve having a mesh opening size of 710  $\mu\text{m}$ . Next, particles having passed through the 710  $\mu\text{m}$  in the aforementioned operation were classified with a JIS standard sieve having a mesh opening size of 150  $\mu\text{m}$ , whereby water-absorbent resin particles (1bF) passing  
10 through the JIS standard sieve having the mesh opening size of 150  $\mu\text{m}$  were removed, thus obtaining a particulate water-absorbent resin (1b).

The resultant water-absorbent resin (1bF) was agglomerated according to the method of Granulation Example 1 as disclosed in USP 6,228,930. The resultant agglomerated material was pulverized and classified by the same procedure as the  
15 aforementioned, thus obtaining an agglomerated water-absorbent resin (1bA).

An amount of 85 mass parts of the water-absorbent resin (1b) and 15 mass parts of the water-absorbent resin (1bA), as obtained in the above way, were uniformly mixed together to obtain a water-absorbent resin (B).

In addition, similarly, the resultant water-absorbent resin (1) was pulverized  
20 with a roll mill and then further classified with a JIS standard sieve having a mesh opening size of 600  $\mu\text{m}$ . Next, particles having passed through the 600  $\mu\text{m}$  in the aforementioned operation were classified with a JIS standard sieve having a mesh opening size of 150  $\mu\text{m}$ , whereby water-absorbent resin particles (1cF) passing through the JIS standard sieve having the mesh opening size of 150  $\mu\text{m}$  were  
25 removed, thus obtaining a particulate water-absorbent resin (1c).

The resultant water-absorbent resin (1cF) was agglomerated according to the method of Granulation Example 1 as disclosed in USP 6,228,930. The resultant agglomerated material was pulverized and classified by the same procedure as the



aforementioned, thus obtaining an agglomerated water-absorbent resin (1cA).

An amount of 80 mass parts of the water-absorbent resin (1c) and 20 mass parts of the water-absorbent resin (1cA), as obtained in the above way, were uniformly mixed together to obtain a water-absorbent resin (C).

5 [Referential Example 2]:

An amount of 100 g of the water-absorbent resin (C) as obtained from the aforementioned Referential Example 1 was mixed with a surface-treating agent comprising a mixed liquid of 0.5 g of 1,4-butanediol, 1.0 g of propylene glycol, and 3.0 g of pure water, and then the resultant mixture was heat-treated at 210 °C for 30  
10 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they could pass through a JIS standard sieve having a mesh opening size of 600 µm. As a result, a surface-crosslink-treated water-absorbent resin (C1) was obtained.

The results of having measured the properties of the water-absorbent resin (C1)  
15 are shown in Table 1.

[Referential Example 3]:

An amount of 100 g of the water-absorbent resin (A) as obtained from the aforementioned Referential Example 1 was mixed with a surface-treating agent comprising a mixed liquid of 0.1 g of 2-ethyloxetane, 3.0 g of pure water, and 0.3 g  
20 of 24 mass % aqueous sodium hydroxide solution, and then the resultant mixture was heat-treated at 200 °C for 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they could pass through a JIS standard sieve having a mesh opening size of 850 µm. As a result, a surface-crosslink-treated water-absorbent resin (A1) was obtained.

25 The results of having measured the properties of the water-absorbent resin (A1) are shown in Table 1.

[Referential Example 4]:

An amount of 100 g of the water-absorbent resin (B) as obtained from the

aforementioned Referential Example 1 was mixed with a surface-treating agent comprising a mixed liquid of 0.5 g of 2-oxazolidinone, 1.0 g of propylene glycol, and 4.0 g of pure water, and then the resultant mixture was heat-treated at 190 °C for 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree  
5 that they could pass through a JIS standard sieve having a mesh opening size of 710  $\mu\text{m}$ . As a result, a surface-crosslink-treated water-absorbent resin (B1) was obtained.

The results of having measured the properties of the water-absorbent resin (B1) are shown in Table 1.

10 [Referential Example 5]:

An amount of 100 g of the water-absorbent resin (C) as obtained from the aforementioned Referential Example 1 was mixed with a surface-treating agent comprising a mixed liquid of 0.5 g of ethylene carbonate, 1.0 g of propylene glycol, and 4.0 g of pure water, and then the resultant mixture was heat-treated at 195 °C for  
15 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they could pass through a JIS standard sieve having a mesh opening size of 600  $\mu\text{m}$ . As a result, a surface-crosslink-treated water-absorbent resin (C2) was obtained.

The results of having measured the properties of the water-absorbent resin (C2)  
20 are shown in Table 1.

[Referential Example 6]:

An amount of 100 g of the water-absorbent resin (A) as obtained from the aforementioned Referential Example 1 was mixed with a surface-treating agent comprising a mixed liquid of 0.5 g of ethylene carbonate, 6.0 g of pure water, and 0.5  
25 g of aluminum sulfate tetradeca- to octadecahydrates (obtained from Kanto Chemical Co., Inc.), and then the resultant mixture was heat-treated at 195 °C for 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they could pass through a JIS standard sieve having a mesh opening size of 850  $\mu\text{m}$ . As

a result, a surface-crosslink-treated water-absorbent resin (A2) was obtained.

The results of having measured the properties of the water-absorbent resin (A2) are shown in Table 1.

[Referential Example 7]:

5 Aluminum sulfate tetradeca- to octadecahydrates (obtained from Kanto Chemical Co., Inc.) were classified with JIS standard sieves having mesh opening sizes of 600  $\mu\text{m}$ , 300  $\mu\text{m}$ , and 150  $\mu\text{m}$ , thus obtaining aluminum sulfate tetradeca- to octadecahydrates (1) having particle diameters of substantially not larger than 150  $\mu\text{m}$ , aluminum sulfate tetradeca- to octadecahydrates (2) having particle diameters of  
10 substantially 300 to 150  $\mu\text{m}$ , and aluminum sulfate tetradeca- to octadecahydrates (3) having particle diameters of substantially 600 to 300  $\mu\text{m}$ . The mass-average particle diameter of the aluminum sulfate tetradeca- to octadecahydrates (1) was 95  $\mu\text{m}$ , the mass-average particle diameter of the aluminum sulfate tetradeca- to octadecahydrates (2) was 203  $\mu\text{m}$ , and the mass-average particle diameter of the  
15 aluminum sulfate tetradeca- to octadecahydrates (3) was 401  $\mu\text{m}$ .

[Example 1]:

An amount of 100 mass parts of the water-absorbent resin (C1), as obtained from Referential Example 2, was uniformly mixed with 0.5 mass part of aluminum sulfate hydrates (trideca- to tetradecahydrates, obtained from Sumitomo Chemical  
20 Co., Ltd., mass-average particle diameter: 165  $\mu\text{m}$ , bulk density: 0.86  $\text{g}/\text{cm}^3$ , solubility into pure water of 0 °C: 46.4 mass %), thus obtaining a water-absorbent resin composition (1).

The results of having measured the properties of the resultant water-absorbent resin composition (1) are shown in Table 1. In addition, the result of having further  
25 measured the CSF is shown in Table 4.

[Example 2]:

An amount of 100 mass parts of the water-absorbent resin (C1), as obtained from Referential Example 2, was uniformly mixed with 1.0 mass part of aluminum

sulfate tetradeca- to octadecahydrates (obtained from Kanto Chemical Co., Inc., mass-average particle diameter: 182  $\mu\text{m}$ , bulk density: 0.60  $\text{g}/\text{cm}^3$ ), thus obtaining a water-absorbent resin composition (2).

The results of having measured the properties of the resultant water-absorbent resin composition (2) are shown in Table 1. In addition, the result of having further measured the retention ratio of the saline flow conductivity (SFC) is shown in Table 2, and the result of having further measured the CSF is shown in Table 4.

[Example 3]:

An amount of 100 mass parts of the water-absorbent resin (A1), as obtained from Referential Example 3, was uniformly mixed with 1.0 mass part of aluminum sulfate hydrates (trideca- to tetradecahydrates, obtained from Sumitomo Chemical Co., Ltd., mass-average particle diameter: 165  $\mu\text{m}$ , bulk density: 0.86  $\text{g}/\text{cm}^3$ ), thus obtaining a water-absorbent resin composition (3).

The results of having measured the properties of the resultant water-absorbent resin composition (3) are shown in Table 1.

[Example 4]:

An amount of 100 mass parts of the water-absorbent resin (B1), as obtained from Referential Example 4, was uniformly mixed with 0.1 mass part of aluminum sulfate hydrates (trideca- to tetradecahydrates, obtained from Sumitomo Chemical Co., Ltd., mass-average particle diameter: 165  $\mu\text{m}$ , bulk density: 0.86  $\text{g}/\text{cm}^3$ ), thus obtaining a water-absorbent resin composition (4).

The results of having measured the properties of the resultant water-absorbent resin composition (4) are shown in Table 1.

[Example 5]:

An amount of 100 mass parts of the water-absorbent resin (C2), as obtained from Referential Example 5, was uniformly mixed with 0.5 mass part of aluminum sulfate hydrates (trideca- to tetradecahydrates, obtained from Sumitomo Chemical Co., Ltd., mass-average particle diameter: 165  $\mu\text{m}$ , bulk density: 0.86  $\text{g}/\text{cm}^3$ ), thus

obtaining a water-absorbent resin composition (5).

The results of having measured the properties of the resultant water-absorbent resin composition (5) are shown in Table 1.

[Example 6]:

5       An amount of 100 mass parts of the water-absorbent resin (C1), as obtained from Referential Example 2, was uniformly mixed with 0.5 mass part of the aluminum sulfate tetradeca- to octadecahydrates (1) as obtained from Referential Example 7, thus obtaining a water-absorbent resin composition (6).

10       The results of having measured the properties of the resultant water-absorbent resin composition (6) are shown in Table 1.

[Example 7]:

15       An amount of 100 mass parts of the water-absorbent resin (C1), as obtained from Referential Example 2, was uniformly mixed with 0.5 mass part of the aluminum sulfate tetradeca- to octadecahydrates (2) as obtained from Referential Example 7, thus obtaining a water-absorbent resin composition (7).

The results of having measured the properties of the resultant water-absorbent resin composition (7) are shown in Table 1.

[Example 8]:

20       An amount of 100 mass parts of the water-absorbent resin (C1), as obtained from Referential Example 2, was uniformly mixed with 0.5 mass part of the aluminum sulfate tetradeca- to octadecahydrates (3) as obtained from Referential Example 7, thus obtaining a water-absorbent resin composition (8).

The results of having measured the properties of the resultant water-absorbent resin composition (8) are shown in Table 1.

25       [Example 9]:

The wettability to a physiological saline solution was evaluated by the contact angle as to water-absorbent resin compositions.

The contact angle with the physiological saline solution was measured by the

aforementioned method as to the water-absorbent resin composition (1) as obtained from Example 1 (aluminum sulfate-added material) (before the PS) and as to a water-absorbent resin composition as obtained in the same way as of Example 1 except the aluminum sulfate was replaced with Aerosil R-972 (produced by Nippon  
5 Aerosil Co., Ltd.).

The results of having measured the contact angle are shown in Table 3.

[Comparative Example 1]:

The water-absorbent resin (C1), as obtained from Referential Example 2, was taken as a comparative water-absorbent resin (1).

10 The results of having measured the properties of the resultant comparative water-absorbent resin (1) are shown in Table 1.

[Comparative Example 2]:

An amount of 100 mass parts of the water-absorbent resin (A), as obtained from Referential Example 1, was uniformly mixed with 0.5 mass part of aluminum  
15 sulfate tetradeca- to octadecahydrates (obtained from Kanto Chemical Co., Inc., mass-average particle diameter: 182  $\mu\text{m}$ , bulk density: 0.60  $\text{g}/\text{cm}^3$ ), thus obtaining a comparative water-absorbent resin composition (2).

The results of having measured the properties of the resultant comparative water-absorbent resin composition (2) are shown in Table 1. In addition, the result  
20 of having further measured the CSF is shown in Table 4.

[Comparative Example 3]:

An amount of 100 mass parts of the water-absorbent resin (A), as obtained from Referential Example 1, was uniformly mixed with 5 mass parts of 10 mass % aqueous solution of aluminum sulfate tetradeca- to octadecahydrates (obtained from  
25 Kanto Chemical Co., Inc., mass-average particle diameter: 182  $\mu\text{m}$ , bulk density: 0.60  $\text{g}/\text{cm}^3$ , solubility into pure water of 23  $^{\circ}\text{C}$ : 37.5 mass %), thus obtaining a comparative water-absorbent resin composition (3).

The results of having measured the properties of the resultant comparative

water-absorbent resin composition (3) are shown in Table 1. In addition, the result of having further measured the CSF is shown in Table 4.

[Comparative Example 4]:

An amount of 100 mass parts of the water-absorbent resin (C1), as obtained  
5 from Referential Example 2, was uniformly mixed with 5 mass parts of 10 mass % aqueous solution of aluminum sulfate tetradeca- to octadecahydrates (obtained from Kanto Chemical Co., Inc., mass-average particle diameter: 182  $\mu\text{m}$ , bulk density: 0.60  $\text{g}/\text{cm}^3$ ), thus obtaining a comparative water-absorbent resin composition (4).

The results of having measured the properties of the resultant comparative  
10 water-absorbent resin composition (4) are shown in Table 1. In addition, the result of having further measured the retention ratio of the saline flow conductivity (SFC) is shown in Table 2.

[Comparative Example 5]:

The water-absorbent resin (A2), as obtained from Referential Example 6, was  
15 taken as a comparative water-absorbent resin (5).

The results of having measured the properties of the resultant comparative water-absorbent resin (5) are shown in Table 1. In addition, the result of having further measured the CSF is shown in Table 4.

[Comparative Example 6]:

20 An amount of 100 mass parts of the water-absorbent resin (C2), as obtained from Referential Example 5, was uniformly mixed with 0.5 mass part of aluminum sulfate hydrates (trideca- to tetradecahydrates, obtained from Sumitomo Chemical Co., Ltd., mass-average particle diameter: 165  $\mu\text{m}$ , bulk density: 0.86  $\text{g}/\text{cm}^3$ ), and then adding 5 mass parts of pure water to the resultant mixture, thus obtaining a  
25 comparative water-absorbent resin composition (6).

The results of having measured the properties of the resultant comparative water-absorbent resin composition (6) are shown in Table 1.

[Example 10]:

Examinations with electron photomicrographs were made about the states of the aluminum sulfates in the water-absorbent resin composition (2) as obtained from Example 2, the comparative water-absorbent resin composition (4) as obtained from Comparative Example 4, and the comparative water-absorbent resin composition (6) as obtained from Comparative Example 6.

In the water-absorbent resin composition (2), the aluminum sulfate was present in the form of particles. However, in the comparative water-absorbent resin compositions (4) and (6), the aluminum sulfate was partially or entirely dissolved and was therefore not present in the form of particles. Incidentally, the presence of the aluminum sulfate was confirmed by EPMA analysis into aluminum.



Table 1 (part 1 of 2)

	Water-absorbent resin as precursor		Mass-average particle diameter $\mu\text{m}$	CRC g/g	AAP g/g	SFC (before PS) $(\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$	SFC (after PS) $(\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$	Retention ratio after PS %
Referential Example 2	C (600-150 $\mu\text{m}$ )	Water-absorbent resin (C1)	302	26	25	72	—	—
Referential Example 3	A (850-150 $\mu\text{m}$ )	Water-absorbent resin (A1)	435	30	26	42	—	—
Referential Example 4	B (710-150 $\mu\text{m}$ )	Water-absorbent resin (B1)	362	24	23	121	—	—
Referential Example 5	C (600-150 $\mu\text{m}$ )	Water-absorbent resin (C2)	301	28	25	54	—	—
Referential Example 6	A (850-150 $\mu\text{m}$ )	Water-absorbent resin (A2)	437	29	23	50	—	—
Example 1	C1	Water-absorbent resin composition (1)	300	26	24	150	121	81
Example 2	C1	Water-absorbent resin composition (2)	304	26	24	168	132	79
Example 3	A1	Water-absorbent resin composition (3)	432	30	25	63	45	71
Example 4	B1	Water-absorbent resin composition (4)	359	24	22	186	135	73
Example 5	C2	Water-absorbent resin composition (5)	297	28	24	102	73	72
Example 6	C1	Water-absorbent resin composition (6)	298	26	24	156	128	82
Example 7	C1	Water-absorbent resin composition (7)	301	26	24	140	134	96
Example 8	C1	Water-absorbent resin composition (8)	305	26	24	128	131	102

Table 1 (part 2 of 2)

	Water-absorbent resin as precursor		Mass-average particle diameter $\mu\text{m}$	CRC g/g	AAP g/g	SFC (before PS) ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )	SFC (after PS) ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )	Retention ratio after PS %
Comparative Example 1	C1	Comparative water-absorbent resin (1)	302	26	25	72	42	58
Comparative Example 2	A	Comparative water-absorbent resin composition (2)	431	33	11	3	—	—
Comparative Example 3	A	Comparative water-absorbent resin composition (3)	430	33	10	2	—	—
Comparative Example 4	C1	Comparative water-absorbent resin composition (4)	301	25	21	112	46	41
Comparative Example 5	A2	Comparative water-absorbent resin (5)	437	29	22	50	23	46
Comparative Example 6	C2	Comparative water-absorbent resin composition (6)	314	27	23	98	43	44

5

Table 2

	Water-absorbent resin		SFC (1 hr) ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )	SFC (2 hr) ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )	Retention ratio of SFC %
Example 2	C1	Water-absorbent resin composition (2)	168	112	67
Comparative Example 4	C1	Comparative water-absorbent resin composition (4)	112	43	38

Table 3

Inorganic compound added	Contact angle (degrees)
Aluminum sulfate tetradeca- to octadecahydrates	16.9
Aerosil R-972	87.7

Note) Aerosil R-972: hydrophobic silica, produced by Nippon Aerosil Co., Ltd. in dry manner

5

Table 4

		CSF g/g
Example 1	Water- absorbent resin composition (1)	23
Example 2	Water- absorbent resin composition (2)	22
Comparative Example 2	Comparative water- absorbent resin composition (2)	7
Comparative Example 3	Comparative water- absorbent resin composition (3)	5
Comparative Example 5	Comparative water- absorbent resin (5)	18

10 [Referential Example 8]:

In a reactor as prepared by lidding a jacketed stainless twin-arm kneader of 10  
liters in capacity having two sigma-type blades, there was prepared a reaction liquid  
by dissolving 11.7 g (0.10 mol %) of polyethylene glycol diacrylate into 5,438 g of  
aqueous solution of sodium acrylate having a neutralization degree of 71.3 mol %  
15 (monomer concentration: 39 mass %). Next, this reaction liquid was deaerated  
under an atmosphere of nitrogen gas for 30 minutes. Subsequently, 29.34 g of 10

mass % aqueous sodium persulfate solution and 24.45 g of 0.1 mass % aqueous L-ascorbic acid solution were added thereto under stirred conditions. As a result, polymerization started after about 1 minute. Then, the polymerization was carried out in the range of 20 to 95 °C while the forming gel was pulverized. Then, the resultant crosslinked hydrogel polymer was taken out after 30 minutes from the start of the polymerization. The crosslinked hydrogel polymer as obtained above was in the form of finely divided pieces having diameters of not larger than about 5 mm. This finely divided crosslinked hydrogel polymer was spread onto a metal gauze of 50 meshes (mesh opening size: 300 µm) and then hot-air-dried at 180 °C for 50 minutes, thus obtaining a water-absorbent resin (1) which was of the irregular shape and easy to pulverize, such as in the form of a particulate dried material agglomerate.

The resultant water-absorbent resin (1) was pulverized with a roll mill and then further classified with a JIS standard sieve having a mesh opening size of 850 µm. Next, particles having passed through the 850 µm in the aforementioned operation were classified with a JIS standard sieve having a mesh opening size of 150 µm, whereby a water-absorbent resin (1aF) passing through the JIS standard sieve having the mesh opening size of 150 µm was removed, thus obtaining a particulate water-absorbent resin (1a).

The removed water-absorbent resin (1aF) was agglomerated according to the method of Granulation Example 1 as disclosed in USP 6,228,930. The resultant agglomerated material was pulverized and classified by the same procedure as the aforementioned, thus obtaining an agglomerated water-absorbent resin (1aA).

An amount of 80 mass parts of the water-absorbent resin (1a) and 20 mass parts of the water-absorbent resin (1aA), as obtained in the above way, were uniformly mixed together to obtain a water-absorbent resin (1C).

Next, 500 g of the water-absorbent resin (1C) and a surface-treating agent comprising a mixed liquid of 2.5 g of 1,4-butanediol, 5.0 g of propylene glycol, and 15.0 g of pure water were mixed together, and then the resultant mixture was

heat-treated at 210 °C for 30 minutes. The heat-treated water-absorbent resin was disintegrated to such a degree that it could pass through a JIS standard sieve having a mesh opening size of 600  $\mu\text{m}$ . As a result, a surface-crosslinked water-absorbent resin (1D) was obtained.

5        Various properties of the surface-crosslinked water-absorbent resin (1D) are shown in Table 5.

The dust generation degree of the surface-crosslinked water-absorbent resin (1D) is shown in Table 6.

[Example 11]:

10        An amount of 300 g of the surface-crosslinked water-absorbent resin (1D), as obtained from Referential Example 8, was preheated to 60 °C, and then spraywise mixed with 1.5 g of water under stirring by a Lödige mixer. Subsequently, 3 g of aluminum sulfate tetradeca- to octadecahydrates was added thereto to mix them together under stirring by the Lödige mixer, and then the resultant mixture was left  
15        alone at room temperature for 30 minutes. The resultant mixture was disintegrated to such a degree that it could pass through a JIS standard sieve having a mesh opening size of 600  $\mu\text{m}$ . As a result, a water-absorbent resin composition (11) was obtained.

20        Various properties of the water-absorbent resin composition (11) are shown in Table 5.

The dust generation degree of the water-absorbent resin composition (11) is shown in Table 6.

[Examples 12 to 14]:

25        Water-absorbent resin compositions (12) to (14) were obtained in the same way as of Example 11 except that the amount of water being used was changed to 3 g, 4.5 g, and 6 g respectively.

Various properties of the water-absorbent resin compositions (12) to (14) are shown in Table 5.

The dust generation degrees of the water-absorbent resin compositions (12) to (14) are shown in Table 6.

[Example 15]:

A water-absorbent resin composition (15) was obtained in the same way as of Example 11 except that the 1.5 g of water was replaced with 6 g of aqueous solution of water/glycerol = 50/50 (wt/wt).

Various properties of the water-absorbent resin composition (15) are shown in Table 5.

[Example 16]:

A water-absorbent resin composition (16) was obtained in the same way as of Example 11 except that the 1.5 g of water was replaced with 6 g of aqueous solution of water/propylene glycol = 50/50 (wt/wt).

Various properties of the water-absorbent resin composition (16) are shown in Table 5.

[Example 17]:

A water-absorbent resin composition (17) was obtained in the same way as of Example 11 except that the 1.5 g of water was replaced with 6 g of aqueous solution of water/polyethylene glycol (average molecular weight: 600) = 50/50 (wt/wt).

Various properties of the water-absorbent resin composition (17) are shown in Table 5.

[Comparative Example 7]:

An amount of 300 g of the surface-crosslinked water-absorbent resin (1D), as obtained from Referential Example 8, and 3 g of aluminum sulfate tetradeca- to octadecahydrates were mixed (dry-blended) together under stirring by a Lödige mixer, thus obtaining a comparative water-absorbent resin composition (7).

Various properties of the comparative water-absorbent resin composition (7) are shown in Table 5.

[Example 18]:

An amount of 500 g of the not yet surface-crosslinked water-absorbent resin (1C), as obtained from Referential Example 8, was preheated to 50 °C, and then mixed with a surface-treating agent comprising a mixed liquid of 2.5 g of 1,4-butanediol, 5.0 g of propylene glycol, and 15.0 g of pure water by a Lödige mixer. Subsequently, 3 g of aluminum sulfate tetradeca- to octadecahydrates was added thereto to mix them together under stirring by the Lödige mixer. The resultant mixture was heat-treated at 210 °C for 30 minutes and then disintegrated to such a degree that it could pass through a JIS standard sieve having a mesh opening size of 600 µm. As a result, a water-absorbent resin composition (18) was obtained.

Various properties of the water-absorbent resin composition (18) are shown in Table 5.

[Comparative Example 8]:

An amount of 500 g of the not yet surface-crosslinked water-absorbent resin (1C), as obtained from Referential Example 8, and a surface-treating agent comprising a mixed liquid of 2.5 g of 1,4-butanediol, 5.0 g of propylene glycol, 3 g of aluminum sulfate tetradeca- to octadecahydrates, and 15.0 g of pure water were mixed together by a Lödige mixer. The resultant mixture was heat-treated at 210 °C for 30 minutes and then disintegrated to such a degree that it could pass through a JIS standard sieve having a mesh opening size of 600 µm. As a result, a comparative water-absorbent resin composition (8) was obtained.

Various properties of the comparative water-absorbent resin composition (8) are shown in Table 5.

[Comparative Example 9]:

An amount of 500 g of the surface-crosslinked water-absorbent resin (1D), as obtained from Referential Example 8, was spraywise mixed with 50 g of 12 mass % aqueous solution of aluminum sulfate tetradeca- to octadecahydrates under stirring by a Lödige mixer. Next, the resultant mixture was dried at 80 °C for 30 minutes. The resultant dried material was disintegrated to such a degree that it could pass

through a JIS standard sieve having a mesh opening size of 600  $\mu\text{m}$ . As a result, a comparative water-absorbent resin composition (9) was obtained.

Various properties of the comparative water-absorbent resin composition (9) are shown in Table 5.

5 [Comparative Example 10]:

An amount of 500 g of the comparative water-absorbent resin composition (7), as obtained from Comparative Example 7, was spraywise mixed with 10 g of water by a Lödige mixer and then left alone at room temperature for 30 minutes. The resultant mixture was disintegrated to such a degree that it could pass through a JIS  
10 standard sieve having a mesh opening size of 600  $\mu\text{m}$ . As a result, a comparative water-absorbent resin composition (10) was obtained.

Various properties of the comparative water-absorbent resin composition (10) are shown in Table 5.

[Example 19]:

15 A water-absorbent resin composition (19) was obtained in the same way as of Example 13 except that aluminum sulfate tetradeca- to octadecahydrates having passed through a JIS standard sieve having a mesh opening size of 105  $\mu\text{m}$  was used.

Various properties of the water-absorbent resin composition (19) are shown in Table 5.

20 [Referential Example 9]:

First of all, a water-absorbent resin (2), which was of the irregular shape and easy to pulverize, such as in the form of a particulate dried material agglomerate, was obtained in the same way as of Referential Example 8 except that the polyethylene glycol diacrylate was used in an amount of 7.6 g.

25 Specifically, in a reactor as prepared by lidding a jacketed stainless twin-arm kneader of 10 liters in capacity having two sigma-type blades, there was prepared a reaction liquid by dissolving 7.6 g (0.065 mol %) of polyethylene glycol diacrylate into 5,438 g of aqueous solution of sodium acrylate having a neutralization degree of



71.3 mol % (monomer concentration: 39 mass %). Next, this reaction liquid was deaerated under an atmosphere of nitrogen gas for 30 minutes. Subsequently, 29.34 g of 10 mass % aqueous sodium persulfate solution and 24.45 g of 0.1 mass % aqueous L-ascorbic acid solution were added thereto under stirred conditions. As a  
5 result, polymerization started after about 1 minute. Then, the polymerization was carried out in the range of 20 to 95 °C while the forming gel was pulverized. Then, the resultant crosslinked hydrogel polymer was taken out after 30 minutes from the start of the polymerization. The crosslinked hydrogel polymer as obtained above was in the form of finely divided pieces having diameters of not larger than about 5  
10 mm. This finely divided crosslinked hydrogel polymer was spread onto a metal gauze of 50 meshes (mesh opening size: 300 µm) and then hot-air-dried at 180 °C for 50 minutes, thus obtaining the water-absorbent resin (2) which was of the irregular shape and easy to pulverize, such as in the form of a particulate dried material agglomerate.

15 The resultant water-absorbent resin (2) was pulverized with a roll mill and then further classified with a JIS standard sieve having a mesh opening size of 850 µm. Next, particles having passed through the 850 µm in the aforementioned operation were classified with a JIS standard sieve having a mesh opening size of 150 µm, whereby a water-absorbent resin (2aF) passing through the JIS standard sieve having  
20 the mesh opening size of 150 µm was removed, thus obtaining a particulate water-absorbent resin (2a).

The removed water-absorbent resin (2aF) was agglomerated according to the method of Granulation Example 1 as disclosed in USP 6,228,930. The resultant agglomerated material was pulverized and classified by the same procedure as the  
25 aforementioned, thus obtaining an agglomerated water-absorbent resin (2aA).

An amount of 90 mass parts of the water-absorbent resin (2a) and 10 mass parts of the water-absorbent resin (2aA), as obtained in the above way, were uniformly mixed together to obtain a water-absorbent resin (2C).

Next, 500 g of the water-absorbent resin (2C) and a surface-treating agent comprising a mixed liquid of 2.5 g of 1,4-butanediol, 5.0 g of propylene glycol, and 15.0 g of pure water were mixed together, and then the resultant mixture was heat-treated at 210 °C for 30 minutes. The heat-treated water-absorbent resin was  
5 disintegrated to such a degree that it could pass through a JIS standard sieve having a mesh opening size of 850 µm. As a result, a surface-crosslinked water-absorbent resin (2D) was obtained.

Various properties of the surface-crosslinked water-absorbent resin (2D) are shown in Table 5.

10 [Example 20]:

A water-absorbent resin composition (20) was obtained in the same way as of Example 13 except that the surface-crosslinked water-absorbent resin (1D) was replaced with the surface-crosslinked water-absorbent resin (2D) as obtained from Referential Example 9.

15 Various properties of the water-absorbent resin composition (20) are shown in Table 5.

[Example 21]:

The distribution of the aluminum sulfate, which was contained every particle diameter range in the water-absorbent resin compositions (11) to (17) as obtained  
20 from Examples 11 to 17 and in the comparative water-absorbent resin composition (7) as obtained from Comparative Example 7, was determined in the following way. The results are shown in Table 7.

(i) The water-absorbent resin composition was sieved with JIS standard sieves having mesh opening sizes of 600 µm, 425 µm, and 300 µm to determine the  
25 distribution into the particle diameter ranges of 600/425 µm, 425/300 µm, and 300 µm-pass.

(ii) An amount of 1 g of the water-absorbent resin composition, as classified into each particle diameter range in the above step (i), was precisely weighed out.

(iii) A Teflon (registered trademark) rotator of 35 mm was placed into a polypropylene-made beaker of 260 ml., and then thereto 1 g of the water-absorbent resin composition, as weighed out in the above step (ii), 190 g of 0.9 mass % aqueous sodium chloride solution, and 10 g of 2N hydrochloric acid were added, and then they were stirred with a magnetic stirrer for 5 minutes.

(iv) After the stirring, the resultant supernatant was sucked up with a polypropylene-made syringe and then filtered with a chromatodisk (GL Chromatodisk 25A, produced by GL Science).

(v) The filtrate was analyzed by ICP (plasma emission spectrometry) to quantify the amount (%) of the aluminum sulfate which was contained every particle diameter range in the water-absorbent resin composition as classified into each particle diameter range.

(vi) The distribution of the aluminum sulfate every particle diameter range was determined in accordance with the following equation:

Distribution (%) of aluminum sulfate every particle diameter range  

$$= [\text{amount (\%)} \text{ of aluminum sulfate every particle diameter range} \times \text{particle diameter distribution (\%)}] \times 100 / \Sigma [\text{amount (\%)} \text{ of aluminum sulfate every particle diameter range} \times \text{particle diameter distribution (\%)}]$$

For example, when the particle diameter distribution in the range of 600/425  $\mu\text{m}$  is 13 % and the amount of the aluminum sulfate in the range of 600/425  $\mu\text{m}$  is 0.20 %, and when the particle diameter distribution in the range of 425/300  $\mu\text{m}$  is 45 % and the amount of the aluminum sulfate in the range of 425/300  $\mu\text{m}$  is 0.17 %, and when the particle diameter distribution in the range of 300  $\mu\text{m}$ -pass is 42 % and the amount of the aluminum sulfate in the range of 300  $\mu\text{m}$ -pass is 0.89 %, then the distribution (%) of the aluminum sulfate in the range of 600/425  $\mu\text{m}$  is determined as follows:

Distribution (%) of aluminum sulfate in the range of 600/425  $\mu\text{m}$   

$$= [0.20 \times 0.13] \times 100 / (0.20 \times 0.13 + 0.17 \times 0.45 + 0.89 \times 0.42) = 5 (\%)$$

Table 5

	Mass-average particle diameter ( $\mu\text{m}$ )	CRC (g/g)	AAP (g/g)	SFC after PS ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )
Water-absorbent resin (1D)	320	26	25	45
Water-absorbent resin (2D)	430	29	24	35
Water-absorbent resin composition (11)	325	26	24	130
Water-absorbent resin composition (12)	323	26	24	120
Water-absorbent resin composition (13)	328	26	24	128
Water-absorbent resin composition (14)	321	26	24	125
Water-absorbent resin composition (15)	335	26	24	123
Water-absorbent resin composition (16)	336	26	24	121
Water-absorbent resin composition (17)	340	26	24	122
Water-absorbent resin composition (18)	320	25	23	118
Water-absorbent resin composition (19)	326	26	24	160
Water-absorbent resin composition (20)	435	29	24	85
Comparative water-absorbent resin composition (7)	320	26	25	120
Comparative water-absorbent resin composition (8)	316	26	22	100
Comparative water-absorbent resin composition (9)	335	26	21	45
Comparative water-absorbent resin composition (10)	340	27	21	40

5

Table 6

	Dust generation degree ( $\text{mg}/\text{m}^3$ )
Water-absorbent resin composition (11)	0.19
Water-absorbent resin composition (12)	0.14
Water-absorbent resin composition (13)	0.14
Water-absorbent resin composition (14)	0.12
Water-absorbent resin (1D)	0.29

Table 7

	Distribution of aluminum sulfate every particle diameter range		
	600/425 $\mu\text{m}$ (%)	425/300 $\mu\text{m}$ (%)	300 $\mu\text{m}$ -pass (%)
Water-absorbent resin composition (11)	4	9	87
Water-absorbent resin composition (12)	5	15	80
Water-absorbent resin composition (13)	5	15	80
Water-absorbent resin composition (14)	5	16	79
Water-absorbent resin composition (15)	6	23	71
Water-absorbent resin composition (16)	7	20	73
Water-absorbent resin composition (17)	11	23	66
Comparative water-absorbent resin composition (7)	2	10	88

## 5 [Referential Example 10]:

In a reactor as prepared by lidding a jacketed stainless twin-arm kneader of 10 liters in capacity having two sigma-type blades, there was prepared a reaction liquid by dissolving 7.14 g (0.06 mol %) of polyethylene glycol diacrylate into 5,438 g of aqueous solution of sodium acrylate having a neutralization degree of 71.3 mol % (monomer concentration: 39 mass %). Next, this reaction liquid was deaerated 10 under an atmosphere of nitrogen gas for 30 minutes. Subsequently, 29.34 g of 10 mass % aqueous sodium persulfate solution and 24.45 g of 0.1 mass % aqueous L-ascorbic acid solution were added thereto under stirred conditions. As a result, polymerization started after about 1 minute. Then, the polymerization was carried 15 out in the range of 20 to 95 °C while the forming gel was pulverized. Then, the resultant crosslinked hydrogel polymer was taken out after 30 minutes from the start of the polymerization.

The crosslinked hydrogel polymer as obtained above was in the form of finely divided pieces having diameters of not larger than about 5 mm. This finely divided 20 crosslinked hydrogel polymer was spread onto a metal gauze of 50 meshes (mesh opening size: 300  $\mu\text{m}$ ) and then hot-air-dried at 180 °C for 50 minutes. The resultant water-absorbent resin was pulverized with a roll mill and then further classified with JIS standard sieves having mesh opening sizes of 850  $\mu\text{m}$  and 150  $\mu\text{m}$ , thus obtaining water-absorbent resin particles (a).

## [Referential Example 11]:

The same polymerization operation as of Referential Example 10 was carried out except that the reaction liquid was replaced with a reaction liquid as prepared by dissolving 4.02 g (0.035 mol %) of polyethylene glycol diacrylate into 5,444 g of aqueous solution of sodium acrylate having a neutralization degree of 75 mol %  
5 (monomer concentration: 38 mass %).

The resultant water-absorbent resin was pulverized with a roll mill and then further classified with JIS standard sieves having mesh opening sizes of 710  $\mu\text{m}$  and 150  $\mu\text{m}$ , thus obtaining water-absorbent resin particles (b).

## 10 [Referential Example 12]:

The same polymerization operation as of Referential Example 10 was carried out except that the reaction liquid was replaced with a reaction liquid as prepared by dissolving 11.7 g (0.1 mol %) of polyethylene glycol diacrylate into 5,438 g of aqueous solution of sodium acrylate having a neutralization degree of 71 mol %  
15 (monomer concentration: 38 mass %).

The resultant water-absorbent resin was pulverized with a roll mill and then further classified with JIS standard sieves having mesh opening sizes of 600  $\mu\text{m}$  and 150  $\mu\text{m}$ , thus obtaining water-absorbent resin particles (c).

## [Referential Example 13]:

20 An amount of 100 g of the water-absorbent resin (a) as obtained from the aforementioned Referential Example 10 was uniformly mixed with a surface-treating agent comprising a mixed liquid of 1.0 g of ethylene carbonate and 3.0 g of pure water, and then the resultant mixture was heat-treated at 180  $^{\circ}\text{C}$  for 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they  
25 could pass through a JIS standard sieve having a mesh opening size of 850  $\mu\text{m}$ . As a result, water-absorbent resin particles (a1) were obtained.

## [Referential Example 14]:

An amount of 100 g of the water-absorbent resin (b) as obtained from the

aforementioned Referential Example 11 was uniformly mixed with a surface-treating agent comprising a mixed liquid of 1.0 g of 2-oxazolidone and 3.0 g of pure water, and then the resultant mixture was heat-treated at 185 °C for 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they could pass through a JIS standard sieve having a mesh opening size of 710 µm. As a result, water-absorbent resin particles (b1) were obtained.

[Referential Example 15]:

An amount of 100 g of the water-absorbent resin particles (c) as obtained from the aforementioned Referential Example 12 were uniformly mixed with a surface-treating agent comprising a mixed liquid of 1.0 g of 1,4-butanediol and 3.0 g of pure water, and then the resultant mixture was heat-treated at 190 °C for 30 minutes. Furthermore, the resultant particles were disintegrated to such a degree that they could pass through a JIS standard sieve having a mesh opening size of 600 µm. As a result, water-absorbent resin particles (c1) were obtained.

[Example 22]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were heated to 140 °C, and then mixed with 2 g of potassium alum (potassium aluminum sulfate dodecahydrate) under stirring, and then the stirring was continued for 10 minutes, thus obtaining a water-absorbent resin composition (22).

[Comparative Example 11]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were mixed with an aqueous solution comprising 2 g of potassium alum (potassium aluminum sulfate dodecahydrate) and 8 g of water under stirring, thus obtaining a comparative water-absorbent resin composition (11).

[Comparative Example 12]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained

from the aforementioned Referential Example 15 were mixed with 2 g of potassium alum (potassium aluminum sulfate dodecahydrate), and then thereto 3 g of water was added under stirring, thus obtaining a comparative water-absorbent resin composition (12).

5 [Example 23]:

An amount of 100 g of the water-absorbent resin particles (b1) as obtained from the aforementioned Referential Example 14 were mixed with 1.5 g of ammonium alum (ammonium aluminum sulfate dodecahydrate) under stirring, and then the resultant mixture was heated to 130 °C, and then the stirring was continued  
10 for 15 minutes, thus obtaining a water-absorbent resin composition (23).

[Comparative Example 13]:

An amount of 100 g of the water-absorbent resin particles (b1) as obtained from the aforementioned Referential Example 14 were mixed with an aqueous solution comprising 1.5 g of ammonium alum (ammonium aluminum sulfate  
15 dodecahydrate) and 8.5 g of water under stirring, thus obtaining a comparative water-absorbent resin composition (13).

[Comparative Example 14]:

An amount of 100 g of the water-absorbent resin particles (b1) as obtained from the aforementioned Referential Example 14 were mixed with 1.5 g of  
20 ammonium alum (ammonium aluminum sulfate dodecahydrate), and then thereto 3 g of water was added under stirring, thus obtaining a comparative water-absorbent resin composition (14).

[As to water-absorbent resin compositions (22) to (23) and comparative water-absorbent resin compositions (11) to (14) as obtained from Examples 22 to 23  
25 and Comparative Examples 11 to 14]:

Shown in Table 8 are the CRC, AAP, and SFC of the water-absorbent resin compositions (22) to (23) and comparative water-absorbent resin compositions (11) to (14) as obtained from Examples 22 to 23 and Comparative Examples 11 to 14.



In addition, shown in Figs. 5 and 6 are: views (Fig. 5-(a) and Fig. 6-(a)) obtained by taking electron photomicrographs of the water-absorbent resin composition (22) as obtained from Example 22; and, as to these photomicrographs, views (Fig. 5-(b) and Fig. 6-(b)) obtained by taking X-ray image photomicrographs of the sulfur element by an SEM-EDS (Energy Dispersive X-ray Spectrometer), wherein the sulfur element is originated from the sulfate ion as contained in the potassium alum. The views obtained by taking the X-ray image photomicrographs of the sulfur element are more sensitive and easier to see than X-ray images of the aluminum element, and thus have been used (the X-ray images of the aluminum element also display the same distributions as of the sulfur element).

Table 8

Example No.	Water-absorbent resin particles or water-absorbent resin composition	Metal compound	Addition method	CRC (g/g)	AAP (g/g)	SFC ( $\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )
Referential Example 15	Water-absorbent resin particles (c1)	—	—	26.3	23.7	73
Example 22	Water-absorbent resin composition (22)	Potassium alum	Heat-fusion	26.4	23.6	164
Comparative Example 11	Comparative water-absorbent resin composition (11)	Potassium alum	Aqueous solution addition	25.3	20.0	101
Comparative Example 12	Comparative water-absorbent resin composition (12)	Potassium alum	Dry mixing + addition of water	25.8	22.9	120
Referential Example 14	Water-absorbent resin particles (b1)	—	—	31.2	25.1	25
Example 23	Water-absorbent resin composition (23)	Ammonium alum	Heat-fusion	31.5	24.7	56
Comparative Example 13	Comparative water-absorbent resin composition (13)	Ammonium alum	Aqueous solution addition	30.1	21.1	32
Comparative Example 14	Comparative water-absorbent resin composition (14)	Ammonium alum	Dry mixing + addition of water	30.5	24.1	36

From Table 8, it can be understood that: even in the case where the same amount of metal compound is added to the same water-absorbent resin particles, the water-absorbent resin compositions according to the present invention have excellent liquid permeability and liquid diffusibility, and also are water-absorbent resin compositions excellent in point of little deterioration of the CRC and AAP.

From Figs. 5 and 6, there can be well seen a state where the metal compound is fused to surfaces of the water-absorbent resin particles. As is illustrated by Fig. 5, in the water-absorbent resin composition according to the present invention, favorably, at least a part of the metal compound is fused in the form of coating at least a part of surfaces of the water-absorbent resin particles in a layered state. In addition, as is illustrated by Fig. 6, there is also preferred a form such that inter-particle binding between the water-absorbent resin particles and particles of the metal compound is formed by the fusion.

[Example 24]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were heated to 170 °C, and then mixed with 1 g of aluminum sulfate tetradeca- to octadecahydrates (mass-average particle diameter: 150 µm) under stirring for 5 minutes, thus obtaining a water-absorbent resin composition (24).

[Comparative Example 15]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were heated to 70 °C, and then mixed with 1 g of aluminum sulfate tetradeca- to octadecahydrates (mass-average particle diameter: 150 µm) under stirring for 5 minutes, thus obtaining a comparative water-absorbent resin composition (15).

[Comparative Example 16]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were mixed with 1 g of aluminum

sulfate tetradeca- to octadecahydrates (mass-average particle diameter: 150  $\mu\text{m}$ ) under stirring for 5 minutes, thus obtaining a comparative water-absorbent resin composition (16).

[Comparative Example 17]:

5       An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were mixed with 1 g of aluminum sulfate tetradeca- to octadecahydrates (mass-average particle diameter: 150  $\mu\text{m}$ ), and then thereto 3 g of water was added under stirring, thus obtaining a comparative water-absorbent resin composition (17).

10       [Comparative Example 18]:

      An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were mixed with an aqueous solution comprising 1 g of aluminum sulfate tetradeca- to octadecahydrates (mass-average particle diameter: 150  $\mu\text{m}$ ) and 4 g of water under stirring, thus  
15       obtaining a comparative water-absorbent resin composition (18).

      [As to water-absorbent resin composition (24) and comparative water-absorbent resin compositions (15) to (18) as obtained from Example 24 and Comparative Examples 15 to 18]:

      Shown in Table 9 are the results of having measured the content of the  
20       aluminum sulfate tetradeca- to octadecahydrates (hereinafter abbreviated to ASH) every particle diameter range in the water-absorbent resin compositions (22) and (24) and comparative water-absorbent resin compositions (15) to (18) as obtained from Examples 22 and 24 and Comparative Examples 15 to 18. Incidentally, the results were calculated from the aluminum contents as measured by the aforementioned  
25       plasma emission spectrometry.

Table 9

Example No.	Water-absorbent resin particles or water-absorbent resin composition	Addition method	AL ppm	AL (<300 $\mu$ m) ppm	Metal compound segregation index
Referential Example 15	Water-absorbent resin particles (c1)	—	—	—	—
Example 22	Water-absorbent resin composition (22)	Heat-fusion	114	104	0.91
Example 24	Water-absorbent resin composition (24)	Heat-fusion	91	137	1.51
Comparative Example 15	Comparative water-absorbent resin composition (15)	Heating to lower than melting point	90	218	2.42
Comparative Example 16	Comparative water-absorbent resin composition (16)	Dry mixing	91	228	2.51
Comparative Example 17	Comparative water-absorbent resin composition (17)	Dry mixing + addition of water	92	204	2.22
Comparative Example 18	Comparative water-absorbent resin composition (18)	Aqueous solution addition	90	143	1.59

AL: aluminum content (ppm) of the water-absorbent resin composition or comparative water-absorbent resin composition  
 AL (<300  $\mu$ m): aluminum content (ppm) of particles passing through a mesh having a mesh opening size of 300  $\mu$ m in the water-absorbent resin composition or comparative water-absorbent resin composition

5

As can be understood from Table 9, because the ASH which is contained in the water-absorbent resin composition (24) as obtained from Example 24 is fused and entirely fixed to surfaces of water-absorbent resin particles, the ASH content of the particles passing through the mesh having a mesh opening size of 300  $\mu$ m is low and almost the same as that in the case where the ASH is added in the form of an aqueous solution (comparative water-absorbent resin composition (18)). In comparison, in the cases of other methods such as dry mixing (comparative water-absorbent resin compositions (15) to (17)), because the ASH having fine particle diameters is not sufficiently fixed, the ASH content of the particles passing through the mesh having a mesh opening size of 300  $\mu$ m is unfavorably high. From these results, it can be understood that the water-absorbent resin compositions according to the present invention are water-absorbent resin compositions excellent in that the segregation of

the contained metal compound occurs very little. In addition, the process in which the aqueous solution of the metal compound is added may be effective for the prevention of the segregation of the metal compound similarly to the process according to the present invention, but, as shown in Table 8, has demerits in that the SFC-enhancing effect is low and in that the deterioration of the AAP is large.

[Example 25]:

An amount of 100 g of the water-absorbent resin particles (a1) as obtained from the aforementioned Referential Example 13 were mixed with 0.8 g of aluminum nitrate nonahydrate under stirring, and then the resultant mixture was heated to 110 °C, and then the stirring was continued for 10 minutes, thus obtaining a water-absorbent resin composition (25).

[Comparative Example 19]:

An amount of 100 g of the water-absorbent resin particles (a1) as obtained from the aforementioned Referential Example 13 were mixed with an aqueous solution comprising 0.8 g of aluminum nitrate nonahydrate and 7.2 g of water under stirring, thus obtaining a comparative water-absorbent resin composition (19).

[Comparative Example 20]:

An amount of 100 g of the water-absorbent resin particles (a1) as obtained from the aforementioned Referential Example 13 were mixed with 0.8 g of aluminum nitrate nonahydrate, and then thereto an aqueous solution comprising 3 g of water and 1 g of polyethylene glycol (average molecular weight: 300) was added under stirring, thus obtaining a comparative water-absorbent resin composition (20).

[As to water-absorbent resin composition (25) and comparative water-absorbent resin compositions (19) to (20) as obtained from Example 25 and Comparative Examples 19 to 20]:

Shown in Table 10 are the results of having measured the CRC, AAP, and BR of the water-absorbent resin composition (25) and comparative water-absorbent resin compositions (19) to (20) as obtained from Example 25 and Comparative Examples

19 to 20.

Table 10

Example No.	Water-absorbent resin particles or water-absorbent resin composition	Metal compound	Addition method	CRC (g/g)	AAP (g/g)	BR %
Referential Example 13	Water-absorbent resin particles (a1)	—	—	36.1	25.9	35.5
Example 25	Water-absorbent resin composition (25)	ANH	Heat-fusion	36.3	25.5	13.2
Comparative Example 19	Comparative water-absorbent resin composition (19)	ANH	Aqueous solution addition	35.2	22.0	33.1
Comparative Example 20	Comparative water-absorbent resin composition (20)	ANH	Dry mixing + addition of water	35.6	24.1	21.7

ANH: aluminum nitrate nonahydrate

5

From Table 10, it can be understood that: even in the case where the same amount of metal compound is added to the same water-absorbent resin particles, the water-absorbent resin composition according to the present invention is a water-absorbent resin composition which undergoes little deterioration of the CRC and AAP and is excellent in the BR.

[Example 26]:

An amount of 100 g of the water-absorbent resin particles (b1) as obtained from the aforementioned Referential Example 14 were heated to 120 °C, and then mixed with 1 g of aluminum chloride hexahydrate under stirring for 5 minutes, thus obtaining a water-absorbent resin composition (26).

[Example 27]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were heated to 100 °C, and then mixed with 2 g of sodium aluminum sulfate dodecahydrate under stirring for 5 minutes, thus obtaining a water-absorbent resin composition (27).

[Comparative Example 21]:

An amount of 100 g of the water-absorbent resin particles (b1) as obtained from the aforementioned Referential Example 14 were heated to 120 °C, and then mixed with 1 g of paraffin wax (melting point: 83 °C) under stirring for 5 minutes, thus obtaining a comparative water-absorbent resin composition (21).

[Comparative Example 22]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were heated to 160 °C, and then mixed with 1 g of zinc caprylate under stirring for 5 minutes, thus obtaining a comparative water-absorbent resin composition (22).

[Comparative Example 23]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were uniformly mixed with 5 g of TRICLOSAN, and then the resultant mixture was heated to 80 °C, and then the mixing was continued under stirring for 1 hour, thus obtaining a comparative water-absorbent resin composition (23).

[As to water-absorbent resin compositions (26) to (27) and comparative water-absorbent resin compositions (21) to (23) as obtained from Examples 26 to 27 and Comparative Examples 21 to 23]:

Shown in Table 11 are the CRC, SFC, and CSF of the water-absorbent resin compositions (26) to (27) and comparative water-absorbent resin compositions (21) to (23) as obtained from Examples 26 to 27 and Comparative Examples 21 to 23.

Table 11

Example No.	Particulate water-absorbent resin or water-absorbent resin composition	Metal compound	Addition method	CRC (g/g)	SFC ( $\times 10^{-7} \text{ cm}^3 \cdot \text{s} \cdot \text{g}^{-1}$ )	CSF (g/g)
Referential Example 14	Water-absorbent resin particles (b1)	—	—	31.2	25	24.1
Example 23	Water-absorbent resin composition (23)	Ammonium alum	Heat-fusion	31.5	56	23.6
Example 26	Water-absorbent resin composition (26)	ACH	Heat-fusion	31.2	51	23.5
Comparative Example 21	Comparative water-absorbent resin composition (21)	Paraffin wax	Heat-fusion	31.0	24	12.1
Referential Example 15	Water-absorbent resin particles (c1)	—	—	26.3	73	25.1
Example 22	Water-absorbent resin composition (22)	Potassium alum	Heat-fusion	26.4	164	22.4
Example 27	Water-absorbent resin composition (27)	ASSH	Heat-fusion	26.3	157	22.1
Comparative Example 22	Comparative water-absorbent resin composition (22)	Zinc caprylate	Heat-fusion	26.2	81	9.8
Comparative Example 23	Comparative water-absorbent resin composition (23)	TRICLOSAN	Heat-fusion	26.1	71	13.1

ACH: aluminum chloride hexahydrate

ASSH: sodium aluminum sulfate dodecahydrate

Zinc caprylate:  $(\text{CH}_3(\text{CH}_2)_6\text{COO})_2\text{Zn}$ 

TRICLOSAN: 2',4',4'-trichloro-2-hydroxydiphenyl ether

5

From Table 11, it can be understood that the water-absorbent resin compositions according to the present invention have excellent SFC and CSF and are excellent in the liquid permeability and liquid diffusibility and in the capillary suction force. As to such as heat-fusible resins like the comparative water-absorbent resin composition (21), the effects of the present invention are not obtained, and, on the contrary, such as deterioration of the CSF is brought about. In



the case where the polyvalent metal salt of the organic acid having not fewer than 7 carbon atoms per molecule is used like the case of the comparative water-absorbent resin composition (22), great deterioration of the CSF is caused. Also in the case of using the organic substance like the case of the comparative water-absorbent resin composition (23), the effects of the present invention are not obtained, and, on the contrary, such as deterioration of the CSF is brought about.

[Comparative Example 24]:

An amount of 100 g of the water-absorbent resin particles (c1) as obtained from the aforementioned Referential Example 15 were mixed with 2 g of potassium alum (potassium aluminum sulfate dodecahydrate) under stirring, and then the stirring was continued for 10 minutes, thus obtaining a comparative water-absorbent resin composition (24).

[As to comparative water-absorbent resin composition (24) as obtained from Comparative Example 24]:

Shown in Table 12 are the BR of the water-absorbent resin composition (22) and comparative water-absorbent resin compositions (11), (12) and (24) as obtained from Example 22 and Comparative Examples 11, 12 and 24.

Table 12

Example No.	Water-absorbent resin particles or water-absorbent resin composition	Metal compound	Addition method	BR %
Referential Example 15	Water-absorbent resin particles (c1)	—	—	28.9
Example 22	Water-absorbent resin composition (22)	Potassium alum	Heat-fusion	4.7
Comparative Example 11	Comparative water-absorbent resin composition (11)	Potassium alum	Aqueous solution addition	25.1
Comparative Example 12	Comparative water-absorbent resin composition (12)	Potassium alum	Dry mixing + addition of water	15.6
Comparative Example 24	Comparative water-absorbent resin composition (24)	Potassium alum	Dry mixing	10.3

From Table 12, it can be understood that the water-absorbent resin composition

according to the present invention has a more excellent BR and is more excellent in the handling property during the moisture absorption when compared with other addition methods.

Shown in Table 13 are the particle diameter distributions of the water-absorbent resin particles (a1), (b1) and (c1) as obtained from Referential Examples 13 to 15 and the water-absorbent resin compositions (22) to (27) as obtained from Examples 22 to 27.

Table 13

Referential Example No. and Example No.	Referential Example 13	Referential Example 14	Referential Example 15	Example 22	Example 23	Example 24	Example 25	Example 26	Example 27
Water- absorbent resin particles	(a1)	(b1)	(c1)	—	—	—	—	—	—
Water- absorbent resin composition	—	—	—	(22)	(23)	(24)	(25)	(26)	(27)
≥ 850μm (mass %)	0.0	0.0	0.0	0.0	0.0	0.0	0.0	0.0	0.0
850-710μm (mass %)	2.9	0.0	0.0	0.0	0.1	0.0	3.1	0.1	0.0
710-600μm (mass %)	28.0	6.5	0.0	0.1	6.8	0.1	28.4	6.6	0.0
600-500μm (mass %)	17.1	8.9	3.1	3.4	9.3	3.3	17.5	9.2	3.3
500-425μm (mass %)	13.7	18.3	17.1	18.4	19.2	17.6	14.1	18.7	17.5
425-300μm (mass %)	20.7	37.3	35.4	35.8	37.1	34.4	21.5	37.4	35.3
300-212μm (mass %)	10.9	16.8	27.4	26.3	16.5	27.2	9.8	16.6	27.8
212-150μm (mass %)	4.2	7.3	13.1	12.5	6.9	13.2	3.5	7.1	12.9
150-45μm (mass %)	2.3	4.8	3.8	3.5	4.1	4.1	2.1	4.3	3.2
≤ 45μm (mass %)	0.2	0.1	0.1	0.0	0.0	0.1	0.0	0.0	0.0
Total (mass %)	100.0	100.0	100.0	100.0	100.0	100.0	100.0	100.0	100.0
D50 (μm)	488	366	315	322	372	315	494	369	317
σ <sub>g</sub>	0.40	0.38	0.37	0.37	0.37	0.38	0.37	0.37	0.36

- 10 (≥ A μm) represents water-absorbent resin particles or a water-absorbent resin composition remaining on a sieve of A in mesh opening size as a result of the classification operation.
- (≤ B μm) represents water-absorbent resin particles or a water-absorbent resin composition having passed through a sieve of B in mesh opening size as a result of the classification operation.
- 15 (A-B μm) represents water-absorbent resin particles or a water-absorbent resin composition having passed through the sieve of A in mesh opening size and remaining on the sieve of B in mesh opening size as a result of the classification operation.

## INDUSTRIAL APPLICATION

Because the water-absorbent resin compositions (1) and (2) according to the present invention have excellent water absorption properties, these water-absorbent resin compositions can be used as water-absorbing and water-retaining agents for various purposes. For example, these water-absorbent resin compositions can be  
5 used for such as: water-absorbing and water-retaining agents for absorbent articles (e.g. disposable diapers, sanitary napkins, incontinent pads, and medical pads); agricultural and horticultural water-retaining agents (e.g. substitutes for peat moss, soil-modifying-and-improving agents, water-retaining agents, and agents for duration of effects of agricultural chemicals); water-retaining agents for buildings (e.g.  
10 dew-condensation-preventing agents for interior wall materials, cement additives); release control agents; coldness-retaining agents; disposable portable body warmers; sludge-solidifying agents; freshness-retaining agents for foods; ion-exchange column materials; dehydrating agents for sludge or oil; desiccating agents; and humidity-adjusting materials. In addition, the water-absorbent resin compositions  
15 (1) and (2) according to the present invention can be used particularly favorably for sanitary materials for absorption of excrement, urine, or blood, such as disposable diapers and sanitary napkins.

## CLAIMS

1. A water-absorbent resin composition, which is a water-absorbent resin composition comprising water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt,  
5 with the composition having a mass-average particle diameter of 100 to 600  $\mu\text{m}$  and comprising water-soluble polyvalent metal salt particles and the water-absorbent resin particles that have been surface-crosslinked.
- 10 2. A water-absorbent resin composition according to claim 1, wherein at least a part of the water-absorbent resin particles are agglomerates.
3. A water-absorbent resin composition according to claim 1 or 2, wherein the water-soluble polyvalent metal salt particles are particles of an aluminum salt  
15 having water of crystallization.
4. A water-absorbent resin composition according to any one of claims 1 to 3, wherein the water-absorbent resin particles are those which have been surface-crosslinked with a polyhydric alcohol.  
20
5. A water-absorbent resin composition, which is a water-absorbent resin composition comprising water-absorbent resin particles and water-soluble polyvalent metal salt particles, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt,  
25 with the composition of which the saline flow conductivity is at least  $50 (\times 10^{-7} \text{ cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$  and of which the retention ratio of the saline flow conductivity is not less than 40 %.

6. A water-absorbent resin composition according to claim 5, of which the retention ratio of the saline flow conductivity after a paint shaker test is not less than 70 %.

5 7. A process for production of a water-absorbent resin composition, which is characterized by comprising the steps of:

adding a binder to water-absorbent resin particles obtained by polymerizing a monomer including acrylic acid and/or its salt; and then

10 mixing the binder and the water-absorbent resin particles with water-soluble polyvalent metal salt particles.

8. A process for production of a water-absorbent resin composition according to claim 7, wherein the water-absorbent resin particles are surface-crosslinked ones.

15

9. A process for production of a water-absorbent resin composition according to claim 7, wherein the binder contains a surface-crosslinking agent.

20 10. A process for production of a water-absorbent resin composition according to any one of claims 7 to 9, wherein the binder includes water and/or a polyhydric alcohol.

25 11. A process for production of a water-absorbent resin composition according to any one of claims 7 to 10, wherein, when the binder is added to the water-absorbent resin particles, the temperature of the water-absorbent resin particles is in the range of 40 to 100 °C.

12. A water-absorbent resin composition, which comprises water-absorbent

resin particles and a metal compound, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt, and wherein:

the metal compound is one or not fewer than two members selected from  
5 among alkaline metal salts and polyvalent metal salts (except polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule); and

at least a part of the metal compound is fused to surfaces of the water-absorbent resin particles.

10 13. A water-absorbent resin composition according to claim 12, wherein the water-absorbent resin particles are materials having been surface-crosslinked with a compound having at least two functional groups which make a dehydration reaction or transesterification reaction with a carboxyl group.

15 14. A water-absorbent resin composition according to claim 12 or 13, wherein at least a part of the metal compound is fused in the form of coating at least a part of surfaces of the water-absorbent resin particles in a layered state

20 15. A water-absorbent resin composition according to any one of claims 12 to 14, wherein the metal compound has a melting point of not higher than 250 °C.

16. A water-absorbent resin composition according to any one of claims 12 to 15, wherein the metal compound is a water-soluble polyvalent metal salt.

25 17. A water-absorbent resin composition according to claim 16, wherein the metal compound is a water-soluble polyvalent metal salt having water of hydration and containing aluminum.

18. A water-absorbent resin composition according to any one of claims 12 to 17, which displays an absorption capacity of not less than 20 g/g under load.

19. A water-absorbent resin composition according to any one of claims 12 to 18, which displays a saline flow conductivity of not less than  $30 (\times 10^{-7} \cdot \text{cm}^3 \cdot \text{s} \cdot \text{g}^{-1})$  for a 0.69 mass % physiological saline solution.

20. A process for production of a water-absorbent resin composition, wherein the water-absorbent resin composition includes water-absorbent resin particles and a metal compound, wherein the water-absorbent resin particles are obtained by polymerizing a monomer including acrylic acid and/or its salt, and wherein:

the metal compound is one or not fewer than two members selected from among alkaline metal salts and polyvalent metal salts (except polyvalent metal salts of organic acids having not fewer than 7 carbon atoms per molecule);

with the process comprising the steps of:

heating the water-absorbent resin particles and/or the metal compound to a temperature of not lower than the melting point of the metal compound; and

thereby fusing at least a part of the metal compound to surfaces of the water-absorbent resin particles.

21. A process for production of a water-absorbent resin composition according to claim 20, wherein the fusion is carried out under stirring of the water-absorbent resin particles and/or the metal compound.

22. A process for production of a water-absorbent resin composition according to claim 20 or 21, wherein the fusion is carried out after a surface-crosslinking treatment of the water-absorbent resin particles.

23. A process for production of a water-absorbent resin composition according to any one of claims 20 to 22, wherein the metal compound has a melting point of not higher than 250 °C.

5

24. A process for production of a water-absorbent resin composition according to any one of claims 20 to 23, wherein the metal compound is a water-soluble polyvalent metal salt having water of hydration and containing aluminum.

10



Fig. 1

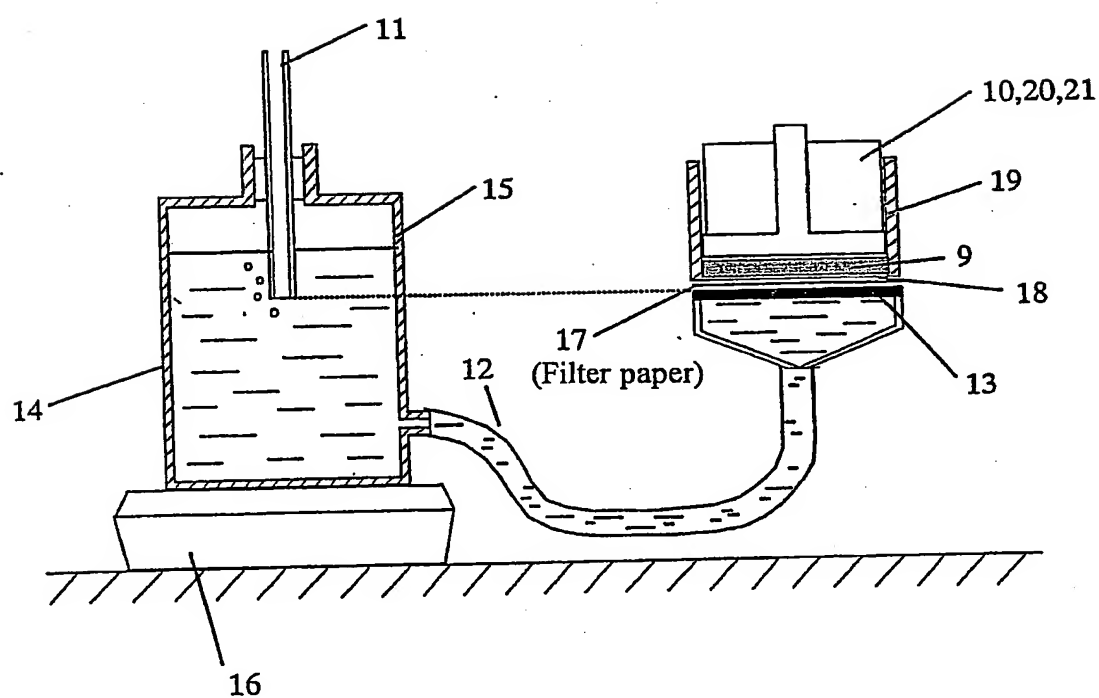


Fig. 2

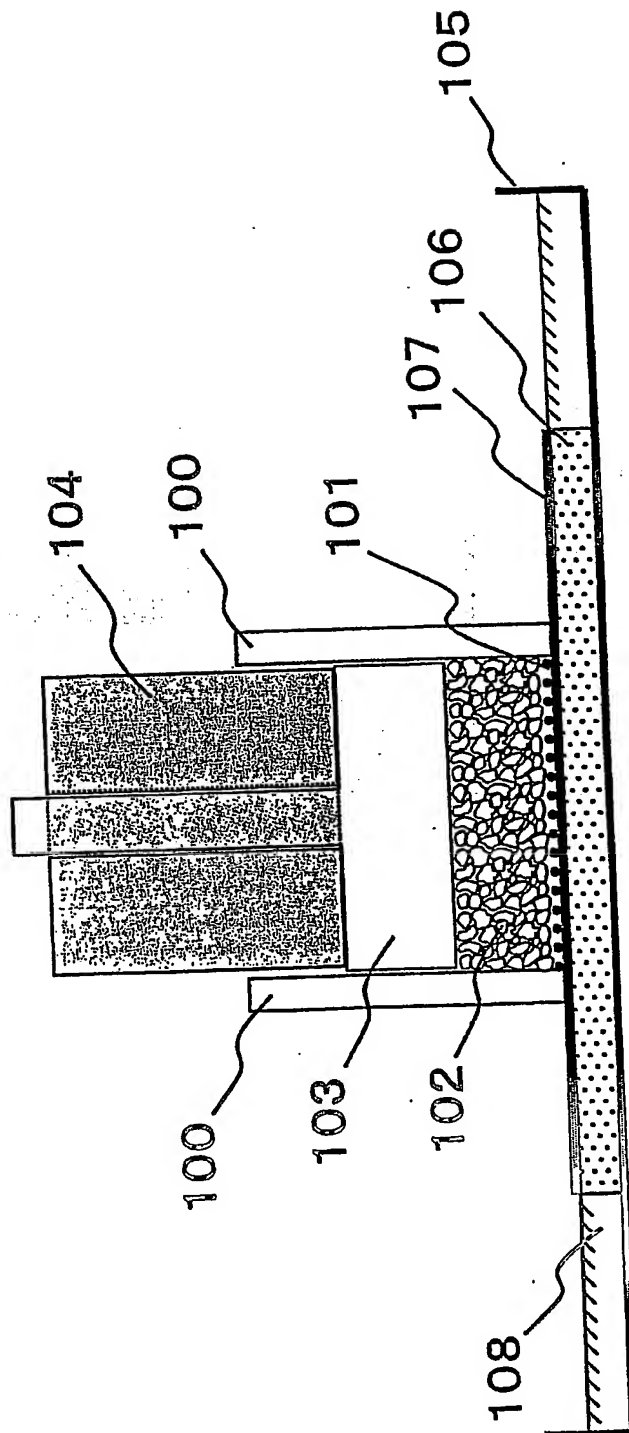


Fig. 3

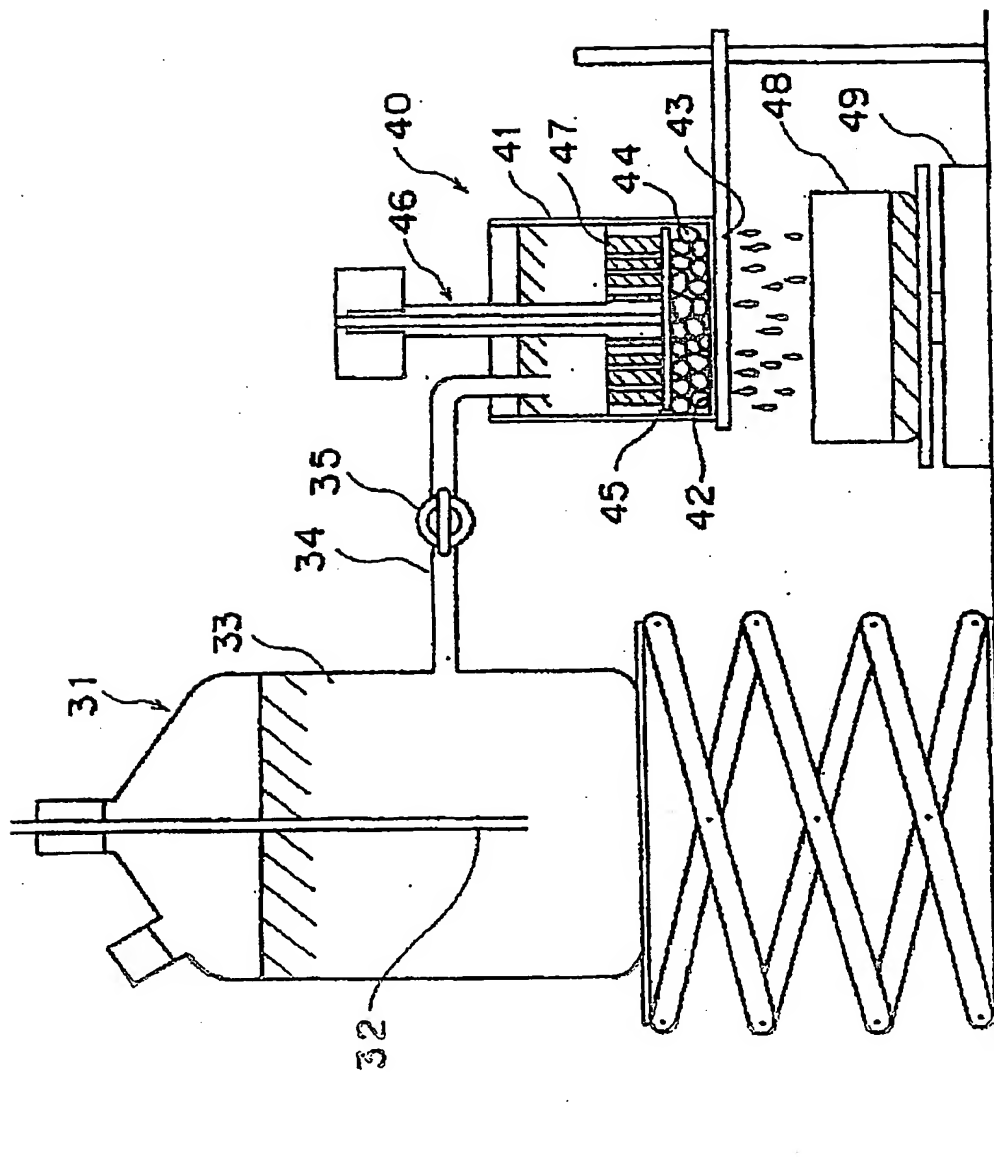


Fig. 4

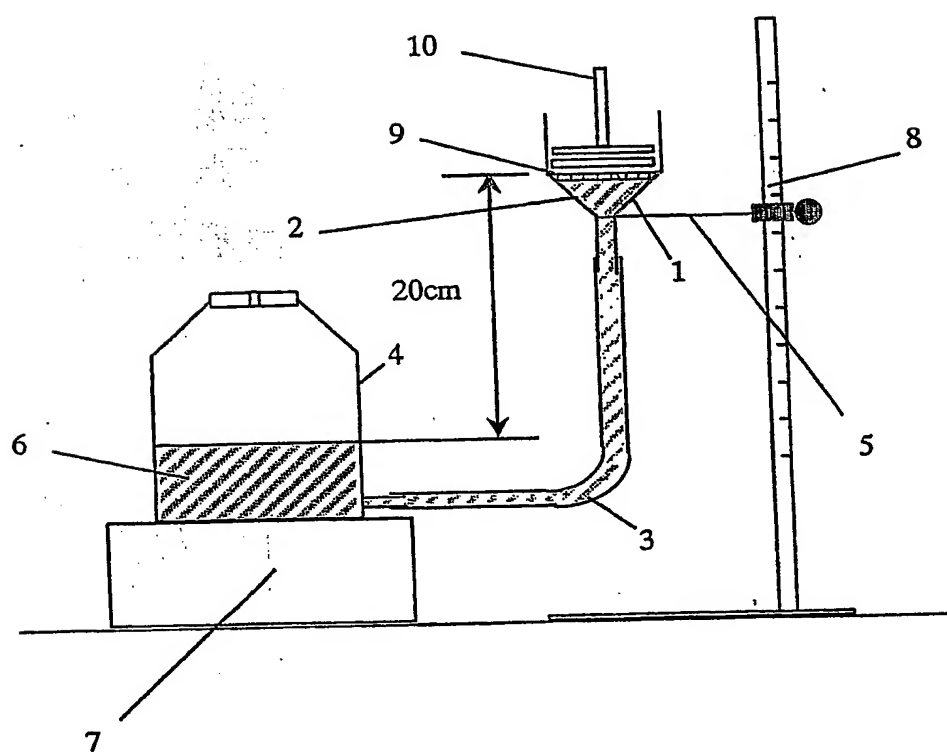


Fig. 5

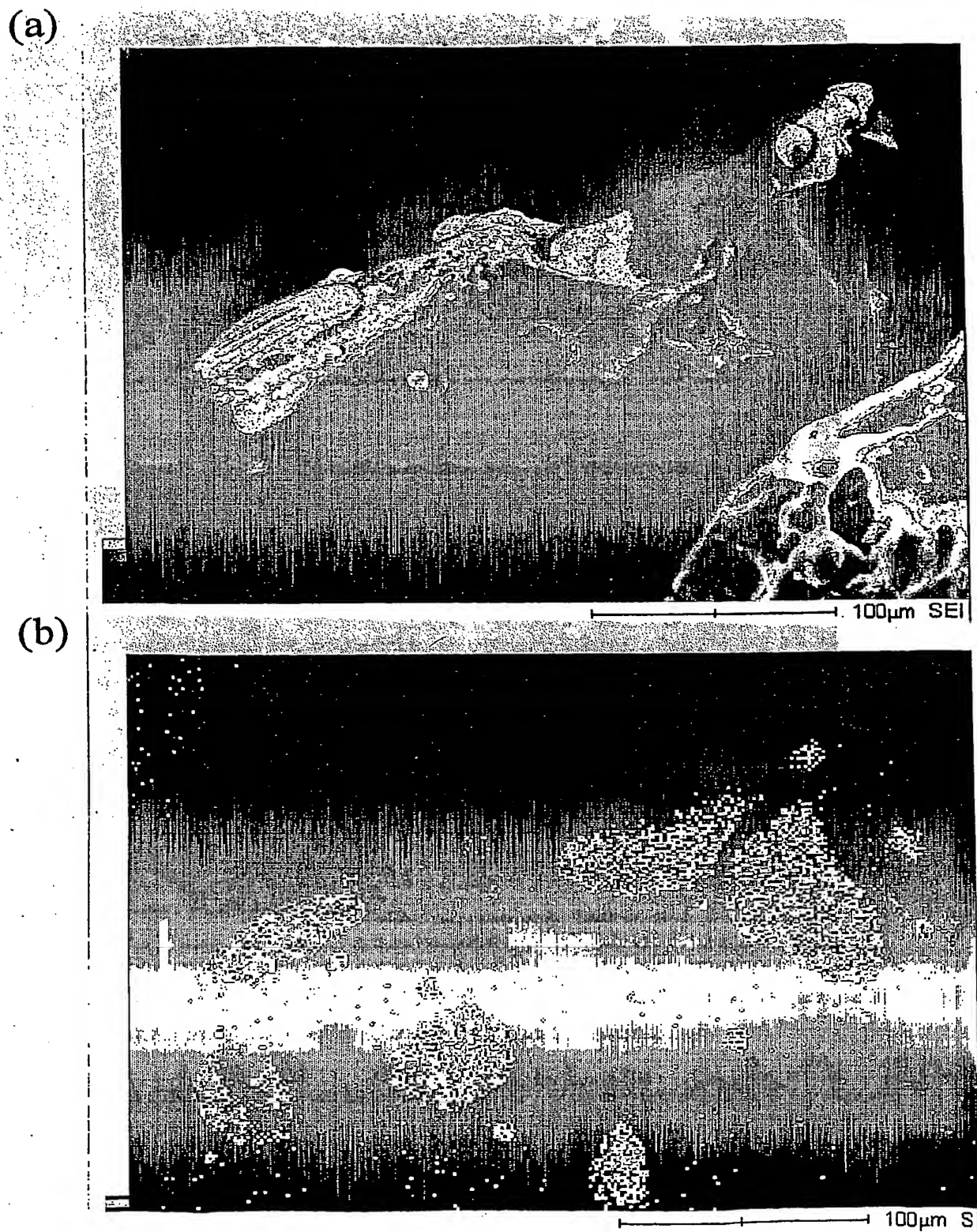
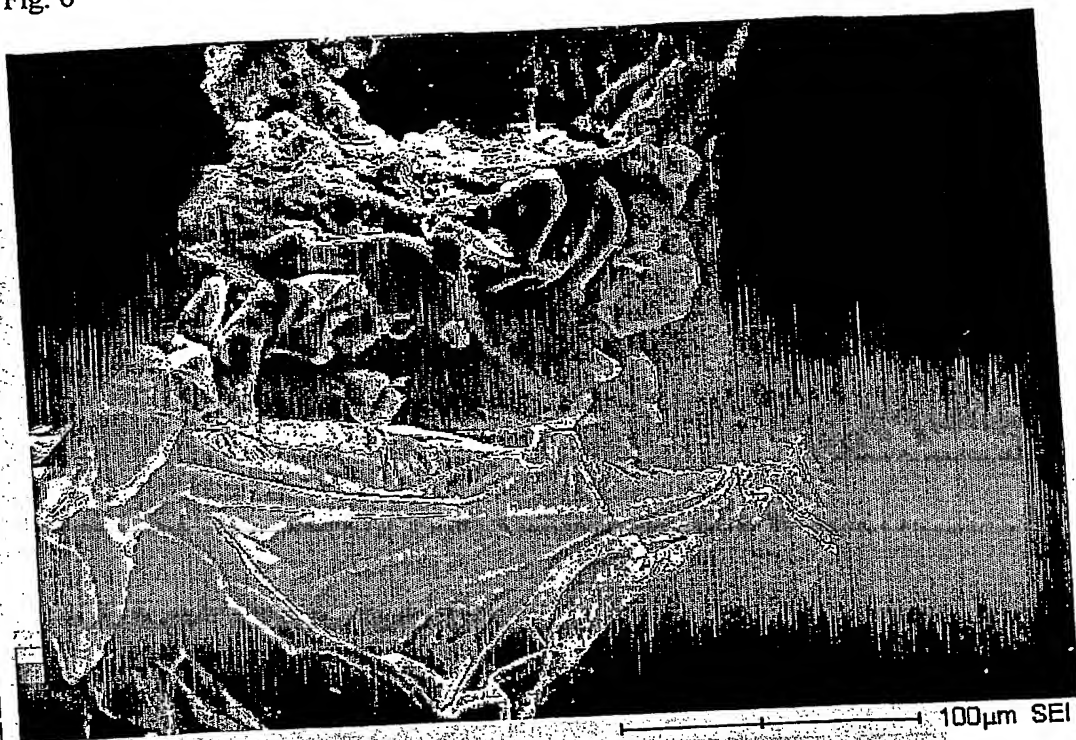
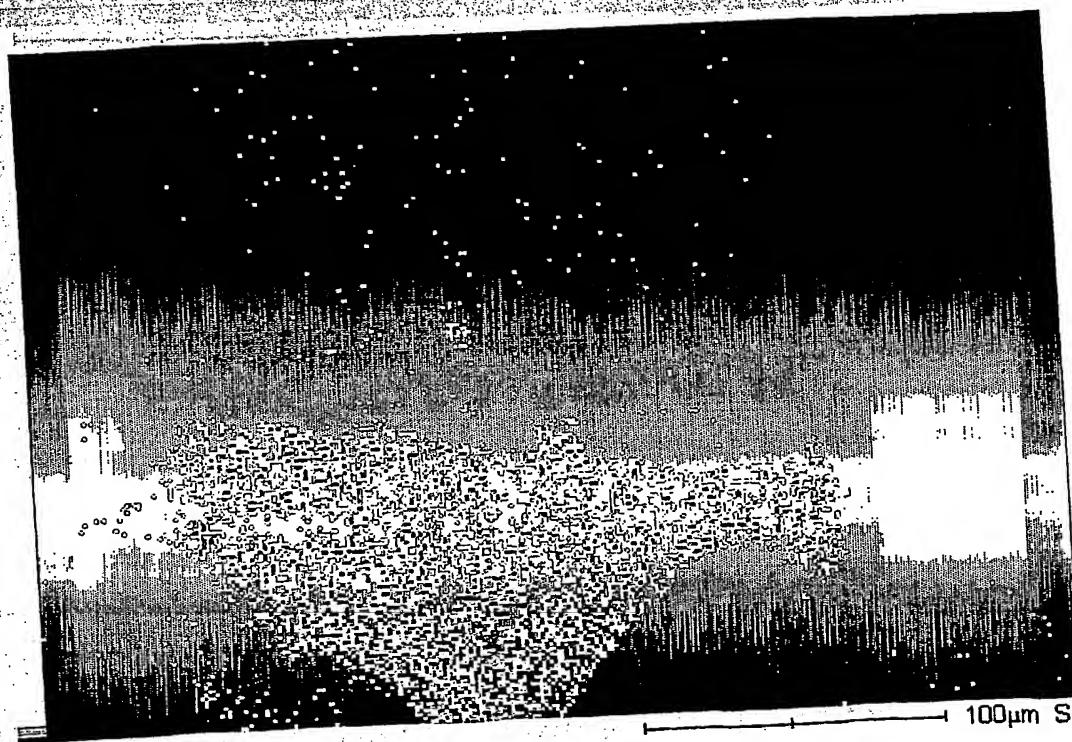


Fig. 6

(a)



(b)



# BEST AVAILABLE COPY

## INTERNATIONAL SEARCH REPORT

International Application No  
PCT/JP2004/001294

**A. CLASSIFICATION OF SUBJECT MATTER**  
IPC 7 A61L15/60

According to International Patent Classification (IPC) or to both national classification and IPC

**B. FIELDS SEARCHED**

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 A61L A61F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data, PAJ

**C. DOCUMENTS CONSIDERED TO BE RELEVANT**

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 4 587 308 A (TANIOKU SHOZO ET AL) 6 May 1986 (1986-05-06) column 2, lines 23-63 column 3 column 4, lines 15-26 column 5 claims; examples	1-24
X	US 5 382 610 A (HARADA NOBUYUKI ET AL) 17 January 1995 (1995-01-17) column 5 column 7 column 9 claims; examples	1-24

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

\* Special categories of cited documents:

- \*A\* document defining the general state of the art which is not considered to be of particular relevance
- \*E\* earlier document but published on or after the international filing date
- \*L\* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- \*O\* document referring to an oral disclosure, use, exhibition or other means
- \*P\* document published prior to the international filing date but later than the priority date claimed

- \*T\* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- \*X\* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- \*Y\* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
- \*Z\* document member of the same patent family

Date of the actual completion of the international search

30 June 2004

Date of mailing of the international search report

09/07/2004

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2  
NL - 2280 HV Rijswijk  
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,  
Fax: (+31-70) 340-3016

Authorized officer

Böhm, I

## INTERNATIONAL SEARCH REPORT

In -  
nal Application No  
PCT/JP2004/001294

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 0 509 708 A (DOW CHEMICAL CO) 21 October 1992 (1992-10-21) page 3 page 4 page 6 claims; examples	1-24
A	EP 0 612 533 A (SANYO CHEMICAL IND LTD) 31 August 1994 (1994-08-31) page 3 page 4 claims; examples	1-24
A	US 5 002 986 A (FUJIURA YOJI ET AL) 26 March 1991 (1991-03-26) the whole document	1-24
A	US 5 112 902 A (MORIYA TETSUO ET AL) 12 May 1992 (1992-05-12) the whole document	1-24
A	WO 91/18042 A (DOW CHEMICAL CO) 28 November 1991 (1991-11-28) the whole document	1



## INTERNATIONAL SEARCH REPORT

Information on patent family members

In International Application No

PCT/JP2004/001294

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
US 4587308	A	06-05-1986	JP 1053974 B	16-11-1989
			JP 1779799 C	13-08-1993
			JP 60163956 A	26-08-1985
			DE 3503458 A1	08-08-1985
			FR 2559158 A1	09-08-1985
US 5382610	A	17-01-1995	JP 3202767 B2	27-08-2001
			JP 5031362 A	09-02-1993
			EP 0493011 A2	01-07-1992
			JP 3017584 B2	13-03-2000
			JP 5096159 A	20-04-1993
EP 0509708	A	21-10-1992	AT 161549 T	15-01-1998
			AU 1487992 A	22-10-1992
			CA 2066010 A1	16-10-1992
			DE 69223674 D1	05-02-1998
			DE 69223674 T2	09-07-1998
			DK 509708 T3	31-08-1998
			EP 0509708 A1	21-10-1992
			ES 2110470 T3	16-02-1998
			FI 921668 A	16-10-1992
			JP 3283570 B2	20-05-2002
			JP 5156034 A	22-06-1993
			KR 195778 B1	15-06-1999
			MX 9201747 A1	01-08-1993
			NO 921492 A	16-10-1992
			US 5633316 A	27-05-1997
EP 0612533	A	31-08-1994	JP 2909692 B2	23-06-1999
			JP 6248187 A	06-09-1994
			JP 2652316 B2	10-09-1997
			JP 6245958 A	06-09-1994
			AT 186649 T	15-12-1999
			AU 663240 B2	28-09-1995
			AU 5484994 A	01-09-1994
			CA 2116160 A1	25-08-1994
			CN 1093932 A , B	26-10-1994
			DE 69421634 D1	23-12-1999
			DE 69421634 T2	02-03-2000
			EP 0612533 A1	31-08-1994
			ES 2138003 T3	01-01-2000
			FI 940845 A	25-08-1994
			KR 138239 B1	27-04-1998
			NO 940618 A	25-08-1994
			TR 28964 A	21-07-1997
			US 5672419 A	30-09-1997
			US 5716707 A	10-02-1998
US 5002986	A	26-03-1991	JP 2111623 C	21-11-1996
			JP 2227435 A	10-09-1990
			JP 6008353 B	02-02-1994
US 5112902	A	12-05-1992	JP 1982928 C	25-10-1995
			JP 2308820 A	21-12-1990
			JP 7010923 B	08-02-1995
			DE 4016338 A1	29-11-1990
			FR 2647364 A1	30-11-1990
			GB 2233656 A , B	16-01-1991

INTERNATIONAL SEARCH REPORT  
information on patent family membersInternational Application No  
PCT/JP2004/001294

Patent document cited in search report		Publication date		Patent family member(s)	Publication date
WO 9118042	A	28-11-1991	AU	7863091 A	10-12-1991
			CA	2082623 A1	20-11-1991
			EP	0530231 A1	10-03-1993
			WO	9118042 A1	28-11-1991
			JP	5507511 T	28-10-1993